

Centre Number	Candidate Number	Name
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UNIVERSITY OF CAMBRIDGE INTERNATIONAL EXAMINATIONS
International General Certificate of Secondary Education

CHEMISTRY

0620/02

Paper 2

October/November 2006

1 hour 15 minutes

Candidates answer on the Question Paper.
No Additional Materials required.

READ THESE INSTRUCTIONS FIRST

Write your Centre number, candidate number and name on all the work you hand in.
Write in dark blue or black pen in the spaces provided on the Question Paper.
You may use a pencil for any diagrams, graphs or rough working.
Do not use staples, paper clips, highlighters, glue or correction fluid.

Answer **all** questions.

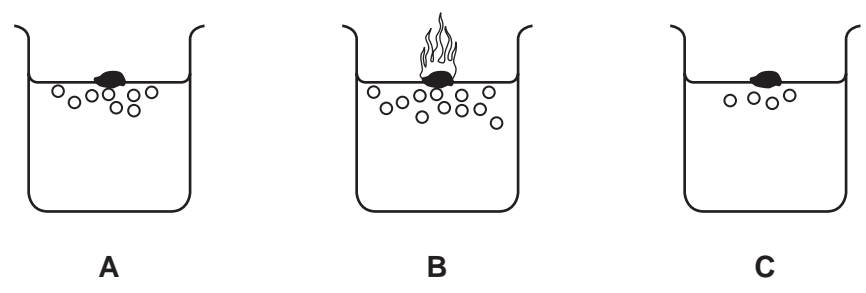
The number of marks is given in brackets [] at the end of each question or part questions.
A copy of the Periodic Table is printed on page 20.

For Examiner's Use	
1	
2	
3	
4	
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6	
7	
Total	

This document consists of **18** printed pages and **2** blank pages.



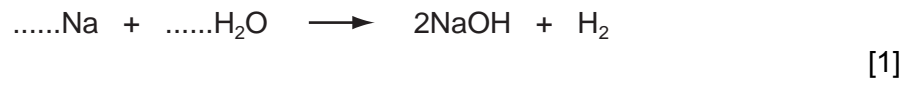
1 When Group I elements react with water, hydrogen gas is given off. The diagram shows the reaction of lithium, potassium and sodium with water.



(a) Which **one** of these elements **A**, **B** or **C** is lithium?

..... [1]

(b) (i) Balance the equation for the reaction of sodium with water by completing the left-hand side.



(ii) Apart from fizzing, describe **two** things that you would **see** when sodium reacts with water.

.....

.....

..... [2]

(iii) After the sodium had reacted with the water, the solution was tested with red litmus paper. What colour did the litmus paper turn? Give a reason for your answer.

colour

reason [2]

(iv) Which of the following statements about sodium are true?
Tick **two** boxes.

It is made by reducing sodium oxide with carbon.

It reacts with chlorine to form sodium chloride.

It reacts readily with oxygen.

It only conducts electricity when molten.

[2]

(c) Rubidium also reacts with water. How does the speed of reaction of rubidium with water compare with that of potassium with water?

..... [1]

(d) Sodium has only one stable isotope whereas potassium has several isotopes.

(i) What do you understand by the term *isotopes*?

.....
..... [1]

(ii) How many protons does sodium have in its nucleus?
Use the Periodic Table to help you.

..... [1]

(iii) How many electrons are there in an atom of potassium?

..... [1]

(iv) Uranium has many isotopes. One of these is uranium-235 (²³⁵U).
What is the main use of this isotope of uranium?

..... [1]

2 Copper can be extracted by heating copper carbonate with carbon.

(a) The copper carbonate breaks down into copper oxide and releases a gas. Complete the equation for this reaction.



(b) The copper oxide then reacts with the carbon.



(i) Complete the following sentences using words from the list.

- endothermic**
- exothermic**
- halogen**
- metal**
- neutralised**
- oxidised**
- reduced**

In this reaction copper oxide is to copper.

The copper obtained is a pinkish-brown

The reaction is because heat is absorbed. [3]

(ii) State the name of the substance which is oxidised during this reaction.

..... [1]

(iii) How would you test for the carbon dioxide given off in this reaction?

test

result [2]

(c) Describe a test for aqueous copper ions and state the result.

.....
.....
..... [3]

(d) Carbon is in Group IV of the Periodic Table.

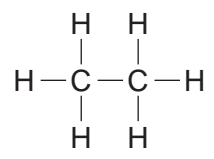
(i) Draw a diagram to show how the electrons are arranged in an atom of carbon.

(ii) To which Period in the Periodic Table does carbon belong? [1]

..... [1]

(e) Organic compounds contain carbon and hydrogen.

(i) To which homologous series does the organic compound **A** belong?



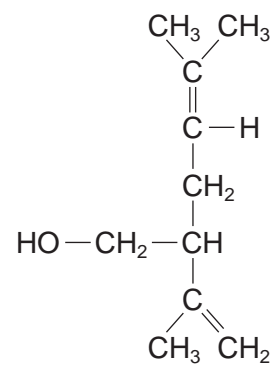
compound **A**

..... [1]

(ii) State the name of compound **A**.

..... [1]

3 Lavandulol is found in lavender plants. The formula of lavandulol is shown below.

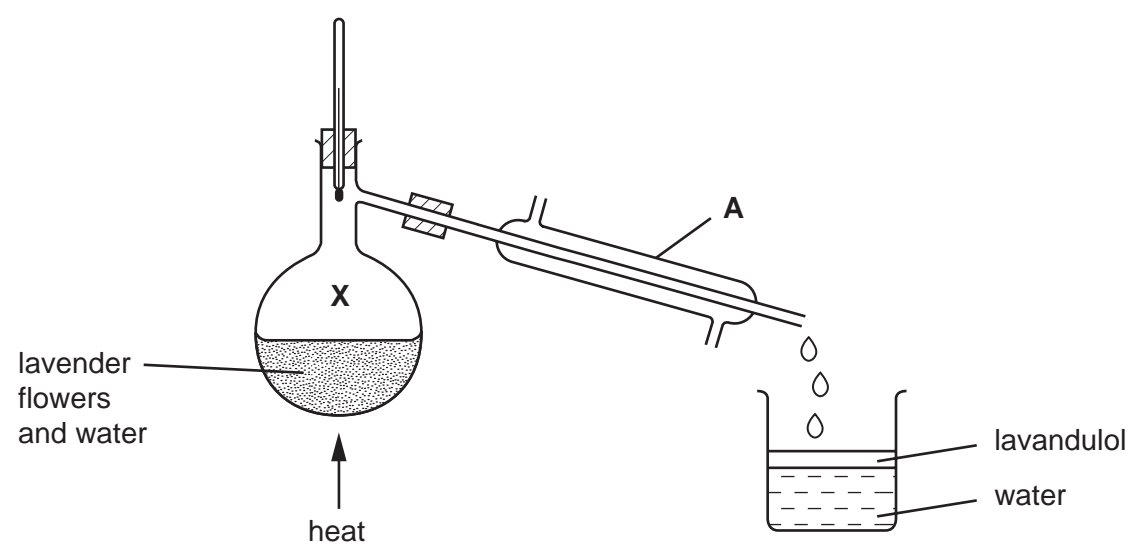


(a) Put a ring around the alcohol functional group in this formula. [1]

(b) Is lavandulol a saturated or unsaturated compound?
Give a reason for your answer.
.....
..... [1]

(c) State the names of the **two** products formed when lavandulol is burnt in excess oxygen.
..... and [2]

(d) Lavandulol can be extracted from lavender flowers by distillation using the apparatus shown below. The lavandulol is carried off in small droplets with the steam.



(i) State the name of the piece of apparatus labelled A.

.....

(ii) What is the temperature of the water at point X in the diagram?

.....

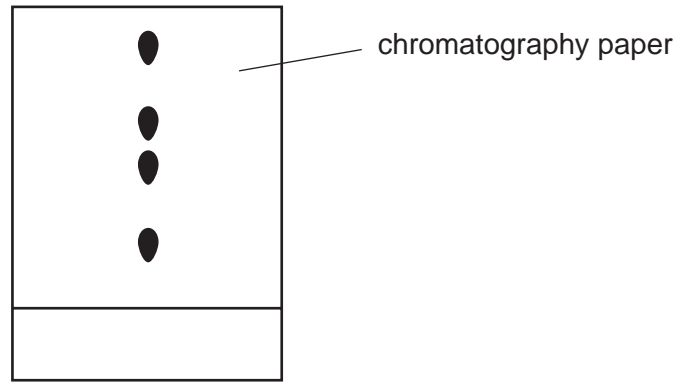
[1]

(iii) The lavender oil and water are collected in the beaker.
What information in the diagram shows that lavender oil is less dense than water?

.....

[1]

(e) Lavender flowers contain a variety of different pigments (colourings).
A student separated these pigments using paper chromatography.
The results are shown in the diagram below.



(i) Put an X on this diagram to show where the mixture of pigments was placed at the start of the experiment. [1]

(ii) How many different pigments have been separated?

.....

[1]

(iii) Draw a diagram to show how the chromatography apparatus was set up.
On your diagram label
• the solvent
• the origin line

[1]

(iv) During chromatography, the solvent evaporates and then diffuses through the chromatography jar.
What do you understand by the term *diffusion*?

.....
..... [1]

(v) Ethanol can be used as a solvent in chromatography.
Draw the formula for ethanol showing all atoms and bonds.

[1]

(vi) Which of the following statements about ethanol are true?
Tick **two** boxes.

It is a carboxylic acid.

It is a product of the fermentation of glucose.

It is an unsaturated compound.

It is formed by the catalytic addition of steam to ethene.

[1]

4 This question is about compounds.

(a) What do you understand by the term *compound*?

.....
..... [1]

(b) Complete the table below to show the formulae and uses of some compounds.

compound	relative number of atoms present	formula	use
calcium oxide	Ca = 1 O = 1	CaO	
sodium chloride	Na = 1 Cl = 1		table salt
calcium carbonate	Ca = 1 C = 1 O = 3		
		NH ₄ NO ₃	in fertilizers

[6]

(c) Calculate the relative formula mass of NH₄NO₃.

[1]

5 The list shows part of the reactivity series.

- strontium
 - calcium
 - magnesium
 - iron
 - copper
- more reactive
↑
less reactive

(a) Calcium is manufactured by the electrolysis of molten calcium chloride. Suggest why calcium is extracted by electrolysis.

..... [1]

(b) Equal sized pieces of magnesium, strontium and calcium are placed in water. Some observations about these reactions are shown in the table. Complete the box for strontium.

metal	observations
magnesium	Gives off a few bubbles of gas with hot water. Dissolves very slowly.
calcium	Gives off bubbles steadily with cold water. Dissolves slowly.
strontium	

[2]

(c) When water is added to calcium carbide, acetylene and calcium hydroxide are formed. State a use for acetylene.

..... [1]

(d) A solution of calcium hydroxide is alkaline.

(i) Complete and balance the equation for the reaction of calcium hydroxide with hydrochloric acid.

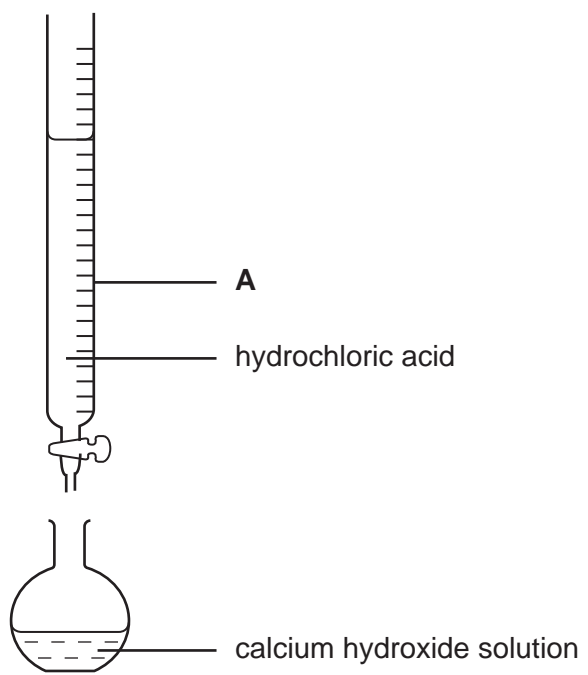


[1]

(ii) What type of chemical reaction is this?

..... [1]

(e) A student used the apparatus shown below to calculate the concentration of a solution of calcium hydroxide.



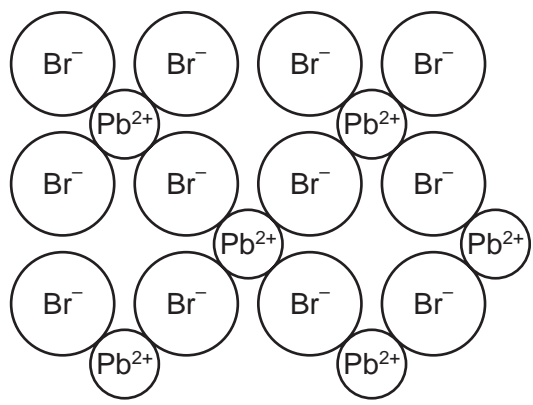
(i) State the name of the piece of apparatus labelled **A**.

..... [1]

(ii) Describe how the pH of the solution in the flask changes as the hydrochloric acid is added.

.....
..... [2]

6 The diagram shows the structure of lead bromide.



(a) What is the simplest formula for lead bromide?

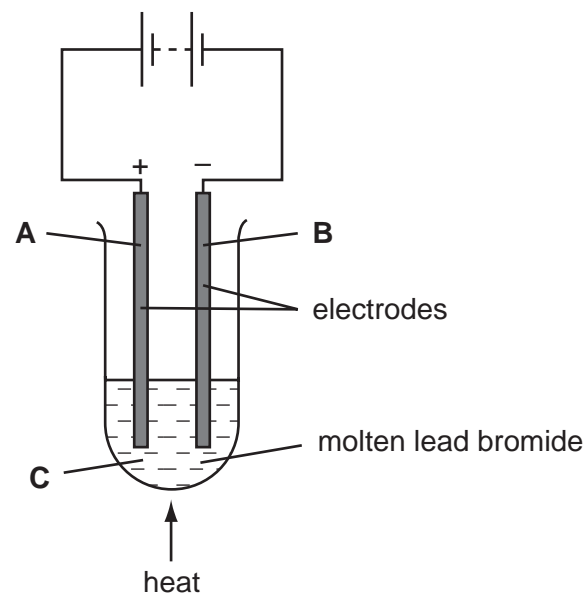
..... [1]

(b) What type of structure and bonding is present in lead bromide?
Choose **two** words from the following:

- atomic covalent giant ionic metallic molecular

..... [2]

(c) Lead bromide is electrolysed using the apparatus shown below.



(i) Which letter, **A**, **B** or **C** represents the cathode?

..... [1]

- (ii) State the name of a metal which can be used for the electrodes.

..... [1]

- (iii) Why does lead bromide have to be molten for electrolysis to occur?

.....
..... [1]

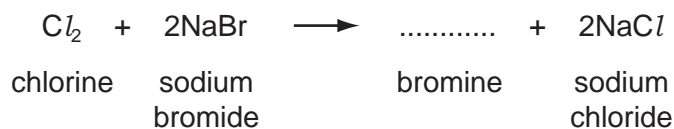
- (iv) State the name of the products formed during this electrolysis;

at the anode,

at the cathode. [2]

- (d) A student bubbled chlorine gas through an aqueous solution of sodium bromide.

- (i) Complete the equation for this reaction.



[1]

- (ii) What colour is the solution at the end of the reaction?

..... [1]

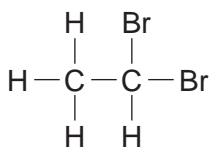
- (iii) An aqueous solution of iodine does not react with a solution of sodium bromide. Explain why there is no reaction.

.....
..... [1]

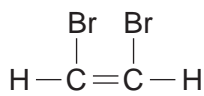
- (e) Bromine becomes decolourised when it reacts with ethene.
(i) Draw the structure of ethene showing all atoms and bonds.

[1]

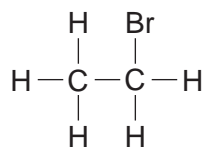
- (ii) Which **one** of the following, **A**, **B**, **C** or **D**, shows the correct structure of the product formed when bromine reacts with ethene?



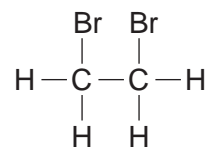
A



B



C



D

answer

[1]

7 The table gives some information about the properties of some metals.

metal	melting point /°C	colour of chloride
A	1890	pink
B	98	white
C	63	white
D	1535	brownish-black

(a) Which **two** of the metals **A** to **D** are transition metals?
Give a reason for your answer.

metals

reason [2]

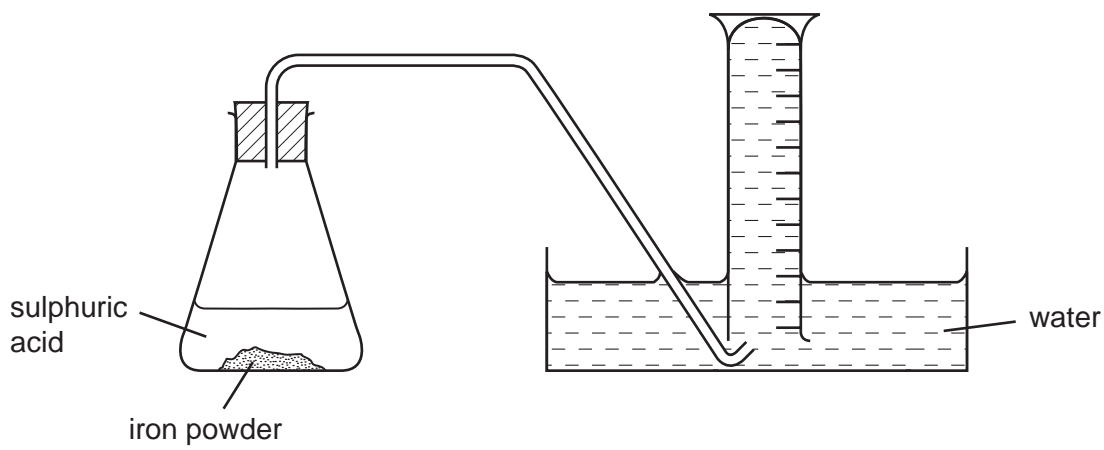
(b) When iron powder reacts with warm sulphuric acid, hydrogen is given off.



State the name of the salt made in this reaction.

..... [1]

(c) A student used the apparatus shown below for investigating the speed of the reaction between iron and sulphuric acid.



Describe how this apparatus can be used to investigate the speed of this reaction.

.....

.....

.....

.....

[3]

(d) The student repeated the experiment with different concentrations of sulphuric acid. In each experiment the mass of iron powder was the same and the temperature was kept at 30°C.

The results are shown in the table.

concentration of sulphuric acid / moles per dm ³	speed of reaction /cm ³ hydrogen per second
0.4	4.2
0.8	8.5
1.6	17.0

(i) Use the information in the table to help you work out how the speed of the reaction is affected by the concentration of sulphuric acid.

.....

.....

[2]

(ii) What will happen to the speed of the reaction if lumps of iron are used instead of iron powder?

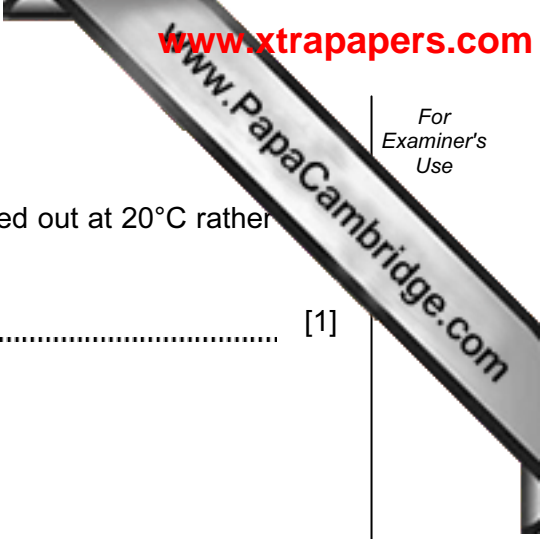
.....

[1]

(iii) What will happen to the speed of the reaction if it is carried out at 20°C rather than at 30°C?

.....

[1]



DATA SHEET
The Periodic Table of the Elements

		Group															
I	II	III	IV	V	VI	VII	O										
1 H Hydrogen																	
3 Li Lithium	4 Be Beryllium	5 B Boron	6 C Carbon	7 N Nitrogen	8 O Oxygen	9 F Fluorine	10 Ne Neon	11 B Boron	12 C Carbon	13 Al Aluminium	14 Si Silicon	15 P Phosphorus	16 S Sulphur	17 Cl Chlorine	18 Ar Argon		
19 K Potassium	20 Ca Calcium	21 Sc Scandium	22 Ti Titanium	23 V Vanadium	24 Cr Chromium	25 Mn Manganese	26 Fe Iron	27 Co Cobalt	28 Ni Nickel	29 Cu Copper	30 Zn Zinc	31 Ga Gallium	32 Ge Germanium	33 As Arsenic	34 Se Selenium	35 Br Bromine	36 Kr Krypton
37 Rb Rubidium	38 Sr Strontium	39 Y Yttrium	40 Zr Zirconium	41 Nb Niobium	42 Mo Molybdenum	43 Tc Technetium	44 Ru Ruthenium	45 Rh Rhodium	46 Pd Palladium	47 Ag Silver	48 Cd Cadmium	49 In Indium	50 Sn Tin	51 Sb Antimony	52 Te Tellurium	53 I Iodine	54 Xe Xenon
55 Cs Caesium	56 Ba Barium	57 La Lanthanum	72 Hf Hafnium	73 Ta Tantalum	74 W Tungsten	75 Re Rhenium	76 Os Osmium	77 Ir Iridium	78 Pt Platinum	79 Au Gold	80 Hg Mercury	81 Tl Thallium	82 Pb Lead	83 Bi Bismuth	84 Po Polonium	85 At Astatine	86 Rn Radon
87 Fr Francium	88 Ra Radium	89 Ac Actinium															

140 Ce Cerium	141 Pr Praseodymium	144 Nd Neodymium	150 Sm Samarium	152 Eu Europium	157 Gd Gadolinium	162 Dy Dysprosium	165 Ho Holmium	167 Er Erbium	169 Tm Thulium	173 Yb Ytterbium	175 Lu Lutetium
58 Th Thorium	91 Pa Protactinium	92 U Uranium	94 Pu Plutonium	95 Am Americium	96 Cm Curium	98 Cf Californium	99 Es Einsteinium	100 Fm Fermium	101 Md Mendelevium	102 No Nobelium	103 Lr Lawrencium

*58-71 Lanthanoid series
90-103 Actinoid series

Key

a	X	a = relative atomic mass
	X	X = atomic symbol
b		b = proton (atomic) number

The volume of one mole of any gas is 24 dm³ at room temperature and pressure (r.t.p.).