

UNIVERSITY OF CAMBRIDGE INTERNATIONAL EXAMINATIONS International General Certificate of Secondary Education

MARK SCHEME for the May/June 2007 question paper

0620 CHEMISTRY

0620/06

Paper 6 (Alternative to Practical), maximum raw mark 60

This mark scheme is published as an aid to teachers and candidates, to indicate the requirements of the examination. It shows the basis on which Examiners were instructed to award marks. It does not indicate the details of the discussions that took place at an Examiners' meeting before marking began.

All Examiners are instructed that alternative correct answers and unexpected approaches in candidates' scripts must be given marks that fairly reflect the relevant knowledge and skills demonstrated.

Mark schemes must be read in conjunction with the question papers and the report on the examination.

• CIE will not enter into discussions or correspondence in connection with these mark schemes.

CIE is publishing the mark schemes for the May/June 2007 question papers for most IGCSE, GCE Advanced Level and Advanced Subsidiary Level syllabuses and some Ordinary Level syllabuses.

Pa	ge 2	Mark Scheme	Syllabus er
		IGCSE – May/June 2007	0620
(a)		mometer (1) ker (1) d (1)	Syllabus 0620 (1) (1)
(b)	to cool / c	condense the vapour (1)	[1]
(c)	measure	the boiling point (1)	[1]
			[Total: 5]
(a)	Correct ir	ndication of electrodes (1)	[1]
(b)	bubbles / bulb light	fizz / effervescence (1) / green gas / level of liquid falls (1) s up (1) max 2	[2]
(c)	(i) chloi	rine / Cl_2 (1)	[1]
	(ii) litmu	is paper / indicator (1) bleaches (1)	[2]
			[Total: 6]
(a)	does not	dissolve in solvent / interfere with results owtte (1)	[1]
(b)	1 and 3 (⁻	1)	[1]
(c)	sample 4 two spots	(1) s present (1)	[2]
(d)	to show p	position of the acids / spots (1)	[1]
			[Total: 5]

Page 3	Mark Scheme	Syllabus er
	IGCSE – May/June 2007	0620 73
Table of result	3	ant
Experiment 1 final reading be	ox correctly completed, 39.2 (1)	Syllabus 0620 Compared of the compared of the
	ox correctly completed (1) npleted correctly, 39.2 (1) and 20.6 (1)	[4]
	cator owtte	[1]
(b) (i) Expe	riment 1 (1)	[1]
(ii) more	in Experiment 1 / greater volume (1)	[1]
(iii) solut	on A more concentrated / stronger than B (1) approx ×2 (1) [2]
(c) 10.3 (1)	cm ³ / ml / cc (1)	[2]
	g. repeat titrations (1) n e.g. average reading more accurate (1)	[2]
		[Total: 13]
(c) bubbles /	ïzz (1) limewater (1) milky (1)	[3]
(d) yellow (1)	precipitate (1)	[2]
(f) carbon die	xide (1)	[1]
(g) ammonia	(1)	[1]
(h) iron (1)	(II) (1) ammonium (1) sulphate (1)	[4]
		[Total: 11]

ble correctly completed	Page 4	Mark Sch		Syllabus 2	er
37 37 37 37 all correct (3) -1 each incorrect [3] all correct (3) -1 each incorrect [4] [6] all points correctly plotted (3) smooth curves (1) labelled (1) [6] solid X (1) faster reaction / more gas given off at 20/40 s (1) [7] same volume of hydrogen peroxide used in both experiments (1) [1] ine sketched on grid with steeper slope than for catalyst X at 25°C (1) levelling out at same level (1) [2] (Total: 14] initial temperature of cold water or cement (1) add cement (1) using thermometer / in beaker etc. (1) measure temperature (1) temperature rise (2) max 4 NB no water = 0 use of heat = 0 would not work = 0) sodium hydrox		IGCSE – May/	June 2007	0620 70	
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