



UNIVERSITY OF CAMBRIDGE INTERNATIONAL EXAMINATIONS
International General Certificate of Secondary Education

CHEMISTRY

0620/12

Paper 1 Multiple Choice

October/November 2012

45 Minutes

Additional Materials: Multiple Choice Answer Sheet
 Soft clean eraser
 Soft pencil (type B or HB is recommended)

* 5 9 9 2 6 7 6 1 5 9 *

READ THESE INSTRUCTIONS FIRST

Write in soft pencil.
Do not use staples, paper clips, highlighters, glue or correction fluid.
Write your name, Centre number and candidate number on the Answer Sheet in the spaces provided unless this has been done for you.

There are **forty** questions on this paper. Answer **all** questions. For each question there are four possible answers **A, B, C** and **D**.

Choose the **one** you consider correct and record your choice in **soft pencil** on the separate Answer Sheet.

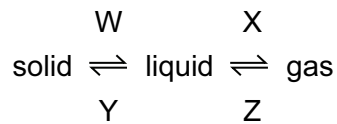
Read the instructions on the Answer Sheet very carefully.

Each correct answer will score one mark. A mark will not be deducted for a wrong answer.
Any rough working should be done in this booklet.
A copy of the Periodic Table is printed on page 16.
You may use a calculator.

This document consists of **15** printed pages and **1** blank page.



- 1 What are the processes W, X, Y and Z in the following diagram?



	W	X	Y	Z
A	condensing	boiling	freezing	melting
B	condensing	freezing	melting	boiling
C	melting	boiling	freezing	condensing
D	melting	freezing	condensing	boiling

- 2 Part of the instructions in an experiment reads as follows.

Quickly add 50 cm³ of acid.

What is the best piece of apparatus to use?

- A** a burette
B a conical flask
C a measuring cylinder
D a pipette
- 3 A mixture of sulfur and iron filings needs to be separated. The solubilities of sulfur and iron filings in water and carbon disulfide are shown in the table below.

	solubility in water	solubility in carbon disulfide
sulfur	x	✓
iron filings	x	x

What are possible methods of separating the sulfur and iron filings?

	using water	using carbon disulfide	using a magnet
A	✓	✓	x
B	x	✓	✓
C	✓	x	✓
D	x	✓	x

- 4 Which row gives the number of electrons in the outer electron shell of fluorine and of neon?

	${}^{19}_{9}\text{F}$	${}^{20}_{10}\text{Ne}$
A	7	8
B	7	10
C	9	8
D	9	10

- 5 Which statements comparing the properties of electrons, neutrons and protons are correct?

	neutrons and protons are both heavier than electrons	only electrons and neutrons are charged
A	✓	✓
B	✓	x
C	x	✓
D	x	x

- 6 The table shows the electronic structures of four atoms.

atom	electronic structure
W	2,1
X	2,7
Y	2,8,4
Z	2,8,8

Which two atoms combine to form an ionic compound?

- A** W and X **B** W and Y **C** X and Y **D** X and Z
- 7 In the molecules CH_4 , HCl and H_2O , which atoms use **all** of their outer shell electrons in bonding?
- A** C and Cl **B** C and H **C** Cl and H **D** H and O

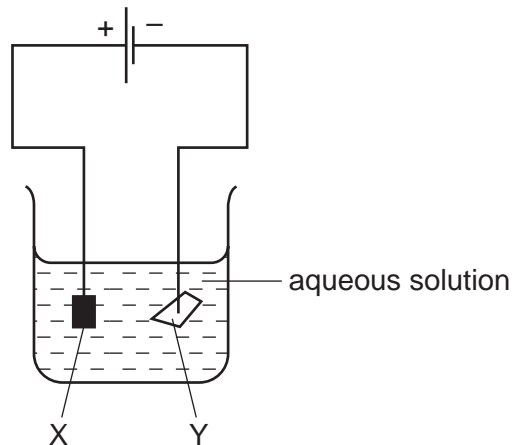
4

8 A compound has the formula $\text{CH}_3\text{CO}_2\text{H}$.

How should the relative molecular mass, M_r , of this compound be calculated?

- A $12 + 1 + 16$
- B $3(12 + 1) + 2(12 + 16) + 1$
- C $(4 \times 12) + (2 \times 1) + 16$
- D $(2 \times 12) + (4 \times 1) + (2 \times 16)$

9 The diagram shows an electrolysis experiment using metals X and Y as electrodes.

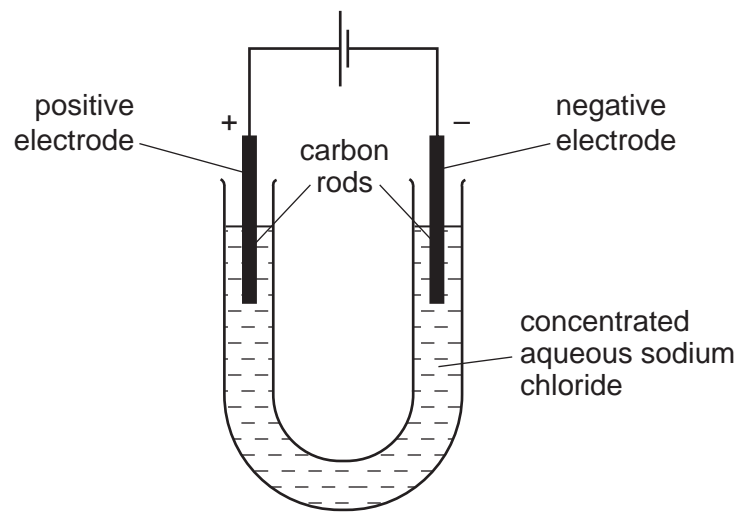


One of the metals becomes coated with copper.

Which metal becomes coated and which aqueous solution is used?

	metal	aqueous solution
A	X	CrCl_3
B	X	CuCl_2
C	Y	CrCl_3
D	Y	CuCl_2

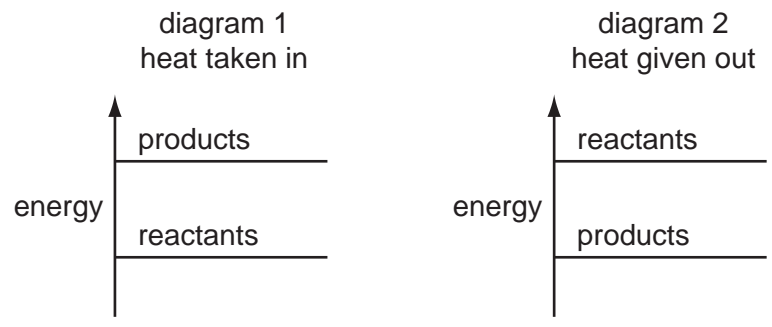
10 The diagram shows the electrolysis of concentrated aqueous sodium chloride.



What is produced at each of the electrodes?

	product at cathode	product at anode
A	hydrogen	chlorine
B	hydrogen	oxygen
C	sodium	chlorine
D	sodium	oxygen

11 The diagrams show the difference in energies of the reactants and products in two types of reaction.



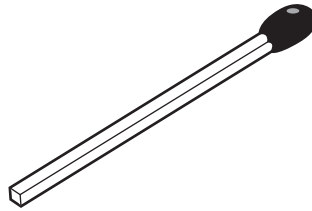
Which diagram and which type of energy change apply to a fuel burning in air?

	diagram	type of energy change
A	1	endothermic
B	1	exothermic
C	2	endothermic
D	2	exothermic

12 Which change is an oxidation?

- A FeO to Fe₂O₃
- B Fe₂O₃ to FeO
- C H₂O₂ to H₂O
- D H₂O to H₂

13 The diagram shows a match.

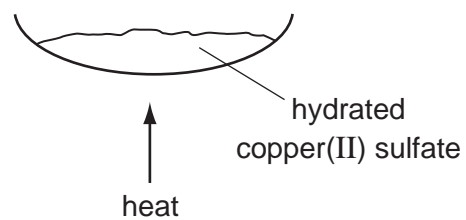
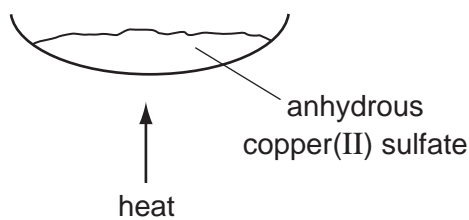


By striking the match, a chemical reaction takes place.

Which statements about the chemical reaction are correct?

	type of reaction	reason
A	endothermic	because energy is used to strike the match
B	endothermic	because energy is given out as the match burns
C	exothermic	because energy is used to strike the match
D	exothermic	because energy is given out as the match burns

14 Separate samples of anhydrous and hydrated copper(II) sulfate are heated.



Which shows the correct colour changes?

	anhydrous copper(II) sulfate	hydrated copper(II) sulfate
A	blue to white	white to blue
B	no change	blue to white
C	white to blue	blue to white
D	white to blue	no change

- 15 Element X forms an acidic, covalent oxide.

Which row shows how many electrons there could be in the outer shell of an atom of X?

	1	2	6	7
A	✓	✓	x	x
B	✓	x	✓	x
C	x	x	✓	✓
D	x	✓	x	✓

- 16 Which change does **not** increase the speed of reaction between zinc and hydrochloric acid?

- A adding a catalyst
- B decreasing the particle size of the zinc
- C decreasing the temperature
- D using more concentrated acid

- 17 Barium hydroxide is an alkali. It reacts with hydrochloric acid.

How does the pH of the hydrochloric acid change as an excess of aqueous barium hydroxide is added?

- A The pH decreases from 14 and becomes constant at 7.
- B The pH decreases from 14 to about 1.
- C The pH increases from 1 and becomes constant at 7.
- D The pH increases from 1 to about 14.

- 18 Which of these pairs of aqueous ions **both** react with dilute sulfuric acid to give a visible result?

- A Ba^{2+} and Cl^-
- B Ba^{2+} and CO_3^{2-}
- C NH_4^+ and Cl^-
- D NH_4^+ and CO_3^{2-}

- 19 A compound is a salt if it

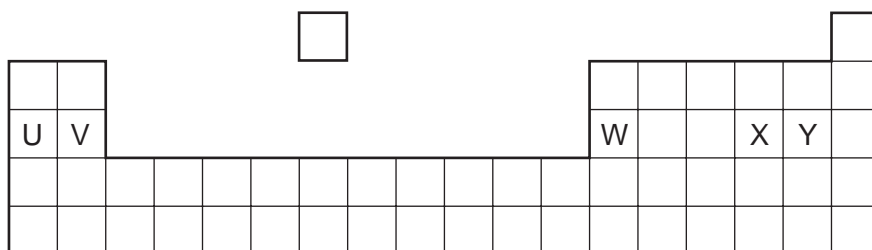
- A can neutralise an acid.
- B contains more than one element.
- C dissolves in water.
- D is formed when an acid reacts with a base.

20 The table gives information about four elements.

Which element is a transition metal?

	colour of element	electrical conductivity of element	colour of oxide
A	black	high	colourless
B	colourless	low	white
C	grey	high	red
D	yellow	low	colourless

21 The diagram shows an outline of the Periodic Table.



Which of the elements U, V, W, X and Y would react together in the ratio of 1 : 1?

- A** U and X **B** U and Y **C** V and Y **D** W and X

22 The element rubidium, Rb, is immediately below potassium in the Periodic Table.

It reacts with bromine to form the compound rubidium bromide.

Which descriptions of this compound are correct?

	type of bond	formula	colour
A	covalent	RbBr	brown
B	covalent	RbBr ₂	white
C	ionic	RbBr	white
D	ionic	RbBr ₂	brown

23 Brass is used in electrical equipment.

It contains two1..... elements. Together they form2..... .

Which words correctly complete gaps 1 and 2?

	1	2
A	metallic	a covalent compound
B	metallic	an alloy
C	non-metallic	a covalent compound
D	non-metallic	an alloy

24 Why are weather balloons filled with helium rather than hydrogen?

- A** Helium is found in air.
- B** Helium is less dense than hydrogen.
- C** Helium is more dense than hydrogen.
- D** Helium is unreactive.

25 Some properties of aluminium are listed.

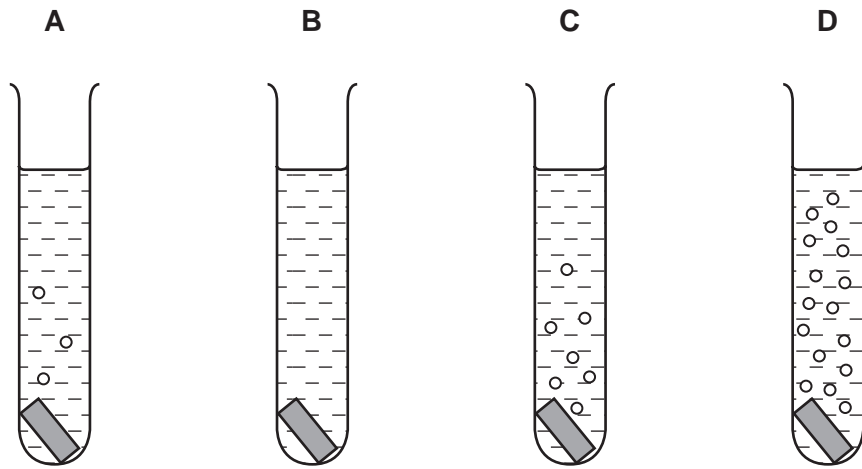
- 1 It has mechanical strength.
- 2 It conducts heat.
- 3 It is resistant to corrosion.
- 4 It has a low density.

Which properties make aluminium useful for making the bodies of aircraft?

- A** 1, 2 and 3 **B** 1, 2 and 4 **C** 1, 3 and 4 **D** 2, 3 and 4

26 Pieces of copper, iron, magnesium and zinc are added to separate test-tubes containing dilute hydrochloric acid.

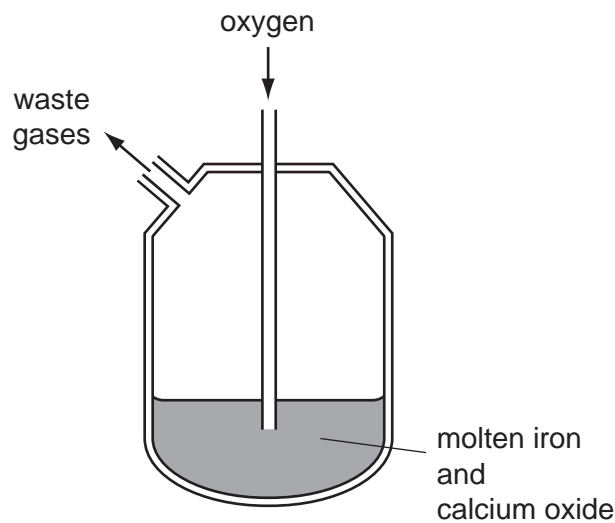
Which test-tube contains iron and dilute hydrochloric acid?



27 The Basic Oxygen Process converts iron into steel.

In step 1, oxygen is blown into impure molten iron.

In step 2, oxides are removed by reaction with calcium oxide.



Which chemical reaction takes place in step 1 and which type of oxides are removed in step 2?

	chemical reaction in step 1	type of oxides removed in step 2
A	carbon is converted to carbon dioxide	acidic
B	carbon is converted to carbon dioxide	basic
C	iron is converted to iron(III) oxide	acidic
D	iron is converted to iron(III) oxide	basic

28 What is the correct order of abundance of the gases in the air?

- A nitrogen → oxygen → argon → carbon dioxide
- B nitrogen → oxygen → carbon dioxide → argon
- C oxygen → nitrogen → argon → carbon dioxide
- D oxygen → nitrogen → carbon dioxide → argon

29 Which processes are used in the treatment of water?

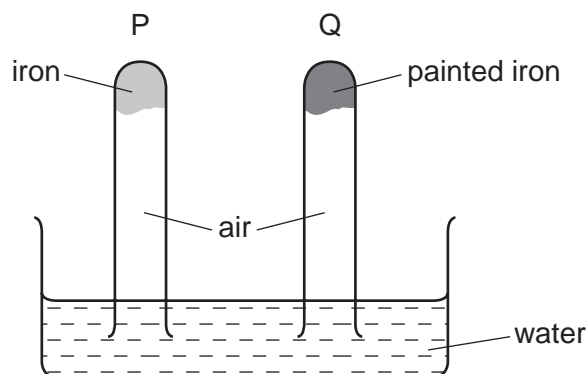
- A filtration and chlorination
- B filtration and reduction
- C neutralisation and chlorination
- D neutralisation and reduction

30 A factory burns coal with a high sulfur content.

Which pollutant is **most** likely to lead to the death of trees?

- A carbon dioxide
- B carbon monoxide
- C lead compounds
- D sulfur dioxide

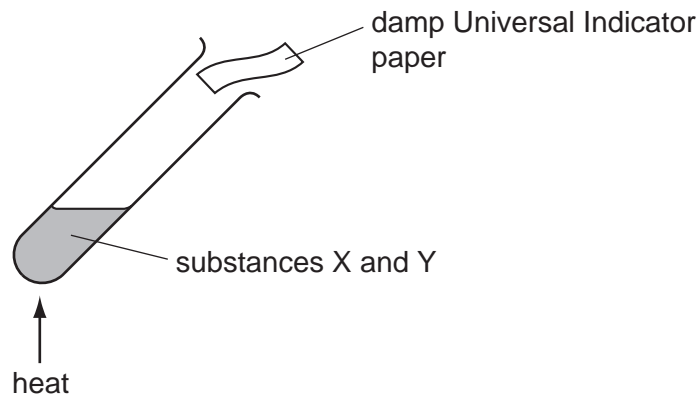
31 The diagram shows an experiment to investigate how paint affects the rusting of iron.



What happens to the water level in tubes P and Q?

	tube P	tube Q
A	falls	rises
B	no change	rises
C	rises	falls
D	rises	no change

- 32 Carbon dioxide is produced when dilute hydrochloric acid reacts with
- A calcium sulfate.
 - B carbon.
 - C copper(II) carbonate.
 - D limewater.
- 33 The diagram shows two substances, X and Y, being heated together.

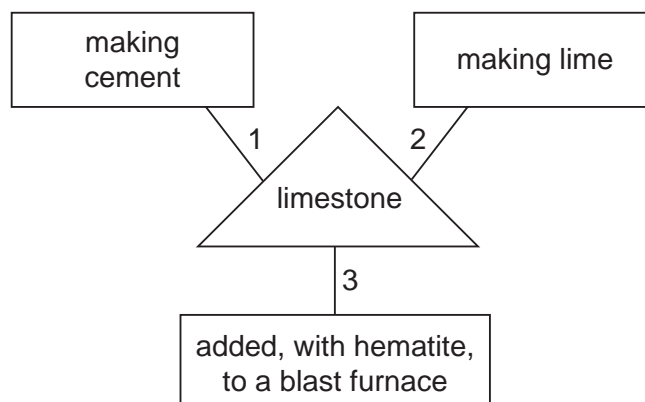


The Universal Indicator paper turns blue during the experiment.

What are substances X and Y?

- A ammonium nitrate and hydrochloric acid
- B ammonium nitrate and sodium hydroxide
- C sodium carbonate and hydrochloric acid
- D sodium carbonate and sodium hydroxide

34 A student is asked to draw a diagram showing the uses of limestone.



Which numbered lines show a correct use of limestone?

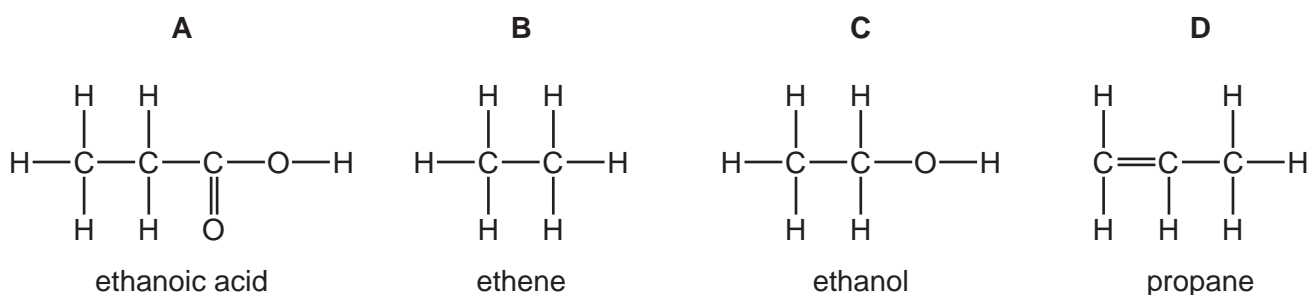
- A 1 and 2 only
- B 1 and 3 only
- C 2 and 3 only
- D 1, 2 and 3

35 Which properties of the different compounds in petroleum enable its separation into fractions?

- 1 boiling point
- 2 chain length
- 3 chemical reactivity
- 4 solubility in water

- A 1 and 2
- B 1 and 3
- C 2 and 4
- D 3 and 4

36 Which structure is correctly named?



37 Alkenes have the general formula C_nH_{2n} .

Which of the following is an alkene?

- A CH_2
- B CH_4
- C C_3H_6
- D C_6H_6

38 A hydrocarbon X is cracked to make Y and hydrogen.

Compound Z is formed by the addition polymerisation of Y.

To which homologous series do X, Y and Z belong?

	alkane	alkene
A	X, Y and Z	–
B	X and Y	Z
C	X and Z	Y
D	Y and Z	X

39 Bitumen is a substance obtained from the fractional distillation of petroleum.

Which row describes its boiling point and the size of its molecules?

	boiling point	size of molecules
A	high	large
B	high	small
C	low	large
D	low	small

40 Which row is correct for ethanol?

	burns	made by fermentation
A	✓	✓
B	✓	x
C	x	✓
D	x	x

DATA SHEET
The Periodic Table of the Elements

		Group																																																																																												
I	II	III	IV	V	VI	VII	0					0																																																																																		
7 Li Lithium 3	9 Be Beryllium 4	1 H Hydrogen 1	11 B Boron 5	12 C Carbon 6	14 N Nitrogen 7	16 O Oxygen 8	19 F Fluorine 9	20 Ne Neon 10	23 Na Sodium 11	24 Mg Magnesium 12	27 Al Aluminium 13	28 Si Silicon 14	31 P Phosphorus 15	32 S Sulfur 16	35.5 Cl Chlorine 17	40 Ar Argon 18	39 K Potassium 19	40 Ca Calcium 20	45 Sc Scandium 21	48 Ti Titanium 22	51 V Vanadium 23	52 Cr Chromium 24	55 Mn Manganese 25	56 Fe Iron 26	59 Co Cobalt 27	59 Ni Nickel 28	64 Cu Copper 29	65 Zn Zinc 30	70 Ga Gallium 31	73 Ge Germanium 32	75 As Arsenic 33	79 Se Selenium 34	80 Br Bromine 35	84 Kr Krypton 36	85 Rb Rubidium 37	88 Sr Strontium 38	89 Y Yttrium 39	91 Zr Zirconium 40	93 Nb Niobium 41	96 Mo Molybdenum 42	101 Ru Ruthenium 44	106 Pd Palladium 46	108 Ag Silver 47	112 Cd Cadmium 48	115 In Indium 49	119 Sn Tin 50	122 Sb Antimony 51	127 I Iodine 53	131 Xe Xenon 54	133 Cs Caesium 55	137 Ba Barium 56	139 La Lanthanum 57	178 Hf Hafnium 72	181 Ta Tantalum 73	184 W Tungsten 74	186 Re Rhenium 75	190 Os Osmium 76	195 Pt Platinum 78	197 Au Gold 79	201 Hg Mercury 80	204 Tl Thallium 81	207 Pb Lead 82	209 Bi Bismuth 83	210 Po Polonium 84	210 At Astatine 85	210 Rn Radon 86	226 Ra Radium 88	227 Ac Actinium 89	232 Th Thorium 90	238 U Uranium 92	238 Np Neptunium 93	238 Pu Plutonium 94	238 Am Americium 95	238 Cm Curium 96	238 Bk Berkelium 97	238 Cf Californium 98	238 Es Einsteinium 99	238 Fm Fermium 100	238 Md Mendelevium 101	238 No Nobelium 102	238 Lr Lawrencium 103	140 Ce Cerium 58	141 Pr Praseodymium 59	144 Nd Neodymium 60	147 Pm Promethium 61	150 Sm Samarium 62	152 Eu Europium 63	157 Gd Gadolinium 64	162 Dy Dysprosium 66	165 Ho Holmium 67	167 Er Erbium 68	169 Tm Thulium 69	173 Yb Ytterbium 70	175 Lu Lutetium 71

*58-71 Lanthanoid series
†90-103 Actinoid series

a	X	a = relative atomic mass
b	X	X = atomic symbol
		b = proton (atomic) number

The volume of one mole of any gas is 24 dm³ at room temperature and pressure (r.t.p.).

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