



UNIVERSITY OF CAMBRIDGE INTERNATIONAL EXAMINATIONS  
International General Certificate of Secondary Education

CANDIDATE  
NAME

CENTRE  
NUMBER

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CANDIDATE  
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**CHEMISTRY**

**0620/23**

Paper 2

**October/November 2012**

**1 hour 15 minutes**

Candidates answer on the Question Paper.

No Additional Materials are required.

**READ THESE INSTRUCTIONS FIRST**

Write your Centre number, candidate number and name in the spaces at the top of this page.

Write in dark blue or black pen.

You may need to use a pencil for any diagrams, graphs or rough working.

Do not use staples, paper clips, highlighters, glue or correction fluid.

**DO NOT WRITE IN ANY BARCODES.**

Answer **all** questions.

A copy of the Periodic Table is printed on page 16.

At the end of the examination, fasten all your work securely together.

The number of marks is given in brackets [ ] at the end of each question or part question.

For Examiner's Use	
1	
2	
3	
4	
5	
6	
7	
8	
<b>Total</b>	

This document consists of **15** printed pages and **1** blank page.



1 Part of the Periodic Table of elements is shown below.

H				He
N	O	F	Ne	
P	S	Cl	Ar	
		Br		
		I		

(a) Answer the following questions using **only** the elements shown in the table above.

Write the symbol for an element which

- (i) is used to fill light bulbs, ..... [1]
- (ii) is in Group VI and Period 3 of the Periodic Table, ..... [1]
- (iii) is a greyish-black solid, ..... [1]
- (iv) forms about 79% of the air, ..... [1]
- (v) consists of single atoms with a full outer shell of electrons, ..... [1]
- (vi) is liberated at the cathode when concentrated hydrochloric acid is  
electrolysed. .... [1]

(b) Hydrogen reacts with chlorine to form hydrogen chloride.

(i) Complete the equation for this reaction.

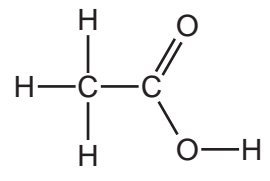


(ii) Draw the electronic structure of a chlorine molecule.  
Show only the outer shell electrons.

[2]

[Total: 10]

2 Vinegar contains ethanoic acid. The formula of ethanoic acid is shown below.

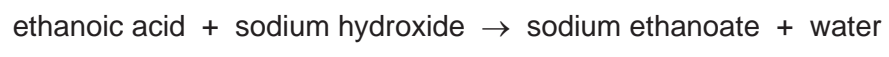


(a) (i) On the formula above, put a ring around the carboxylic acid functional group. [1]

(ii) Write the simplest formula for a molecule of ethanoic acid.

[1]

(b) Ethanoic acid reacts with sodium hydroxide to form the salt sodium ethanoate.



What type of chemical reaction is this?

..... [1]

(c) Sodium ethanoate is soluble in water. What do you understand by the term *soluble*?

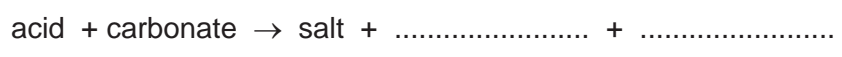
..... [1]

(d) Which **one** of the following is the most likely pH value of ethanoic acid? Put a ring around the correct answer.

- pH 3      pH 7      pH 9      pH 13

[1]

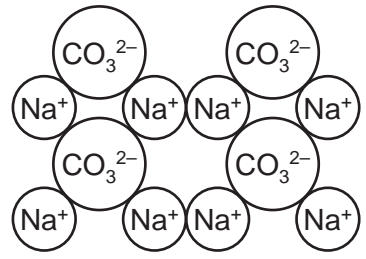
(e) All acids react with carbonates. Complete the general equation for this reaction.



.....

[2]

(f) The structure of sodium carbonate is shown below.

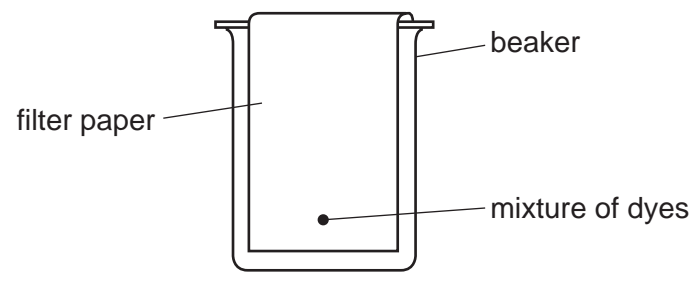


Write the simplest formula for sodium carbonate.

..... [1]

[Total: 8]

3 A student used the apparatus shown below to separate a mixture of coloured dyes. The solvent is not shown.



(a) On the diagram above, draw and label the position of the solvent at the start of the experiment. [1]

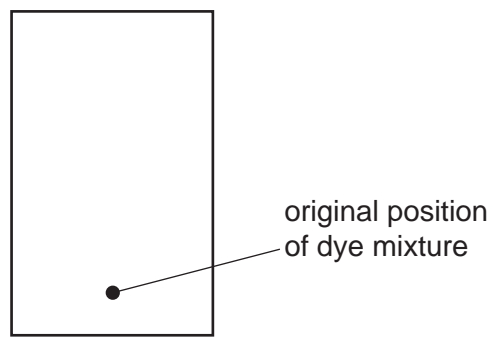
(b) The student let the solvent move up the filter paper to separate the dyes.

(i) State the name of this method of separation.

..... [1]

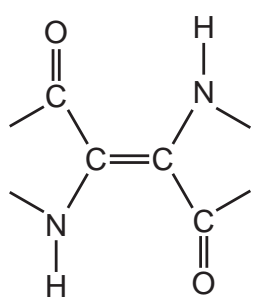
(ii) The student found that four different dyes had been separated by this method. On the diagram below draw

- the position of four separated dyes (show as spots)
- the solvent front (show as a line).



[3]

(c) Part of the structure of a dye called indigo is shown below.

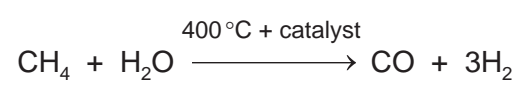


Is this a saturated or unsaturated compound?  
Give a reason for your answer.

..... [1]

[Total: 6]

4 Hydrogen can be manufactured by heating methane with steam.



(a) (i) Draw the structure of methane showing all atoms and bonds.

[1]

(ii) Methane is a greenhouse gas.  
What do you understand by the term *greenhouse gas*?

..... [1]

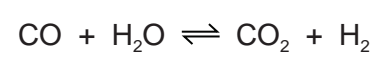
(iii) State **one** source of the methane in the atmosphere.

..... [1]

(iv) When 16 g of methane reacts completely with an excess of steam, 6 g of hydrogen are produced.  
Calculate the mass of methane required to produce 300 g of hydrogen.

Answer = ..... [1]

(b) More hydrogen can be formed by reacting the carbon monoxide with more steam at 500°C.



This reaction is reversible.

(i) How do you know from this equation that the reaction is reversible?

..... [1]

(ii) What do you understand by the term *reversible reaction*?

..... [1]

(iii) Carbon monoxide is a common atmospheric pollutant.  
State a source of the carbon monoxide in the atmosphere other than from  
manufacture of hydrogen.

..... [1]

(iv) Carbon dioxide is a product of the reaction between carbon monoxide and steam.  
Is carbon dioxide an acidic or a basic oxide?  
Give a reason for your answer.

..... [1]

[Total: 8]

5 Ethanol can be made by

- an addition reaction with ethene or
- by fermentation.

(a) (i) State the name of the substance that needs to be added to ethene to make ethanol.

..... [1]

(ii) What conditions are needed to make ethanol from ethene?

.....  
..... [2]

(b) (i) Complete the word equation for fermentation in the presence of yeast.

..... → ethanol + .....

[2]

(ii) The yeast contains enzymes.  
What do you understand by the term *enzyme*?

.....  
..... [2]

(c) The speed of ethanol formation during fermentation depends on the temperature.

(i) Use the information in the table below to describe how the speed of this reaction changes with temperature.

temperature /°C	speed of reaction /g ethanol formed per hr
10	1
20	3
30	7
40	11
50	6
60	2
70	0

.....  
.....  
..... [3]



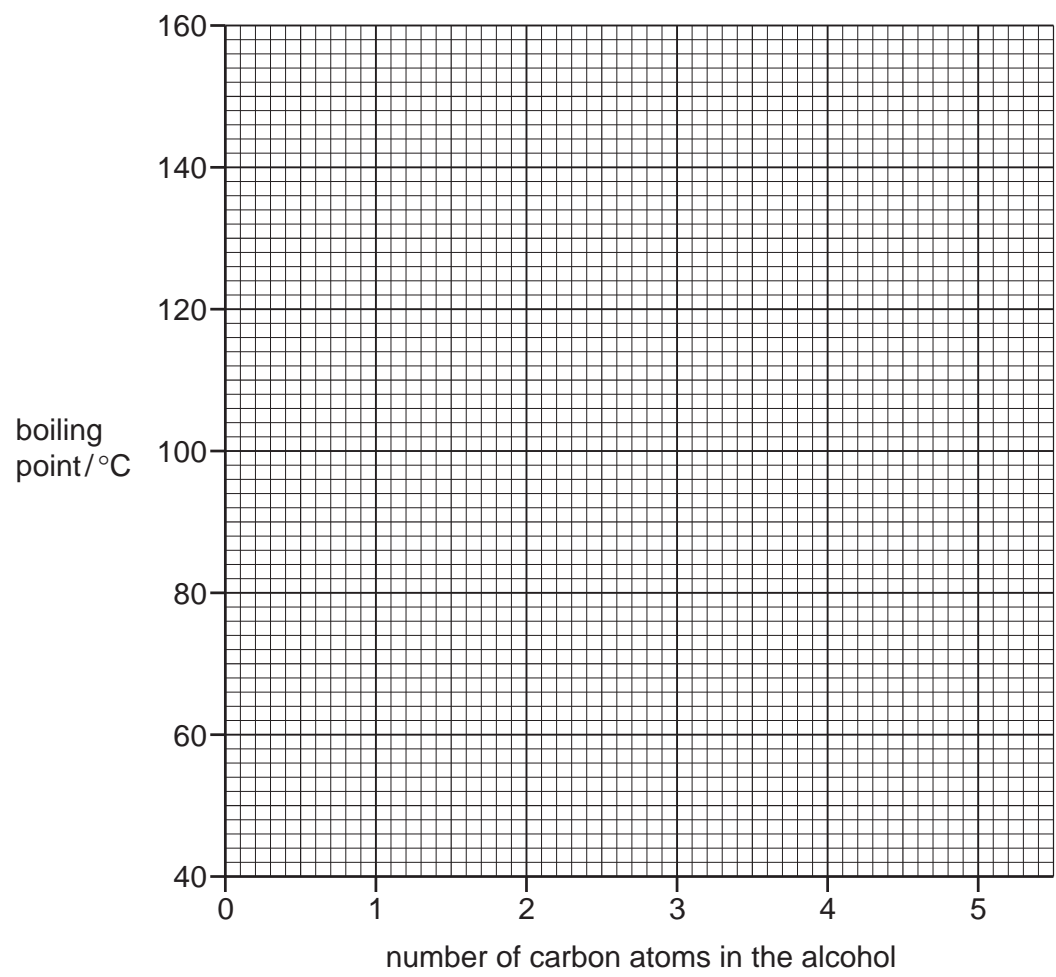
(ii) State **two** factors which should be kept constant during this experiment.

.....  
..... [2]

(d) Ethanol belongs to the alcohol homologous series.  
The boiling points of some alcohols are given in the table below.

alcohol	number of carbon atoms in the alcohol	boiling point / °C
methanol	1	65
ethanol	2	79
propanol	3	98
butanol	4	117

(i) On the grid below, plot a graph of boiling point against the number of carbon atoms. Join the points with a smooth line.



[3]

(ii) Use your graph to estimate the boiling point of the alcohol having five carbon atoms.

boiling point = .....°C [1]

[Total: 16]

6 Lead and lead compounds are common pollutants of the air.

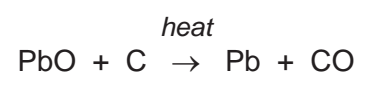
(a) (i) State **one** source of lead in the air.

..... [1]

(ii) State **one** effect of lead on human health.

..... [1]

(b) Lead(II) oxide can be reduced by heating with carbon.



(i) Write a word equation for this reaction.

..... [1]

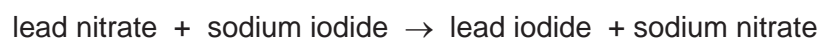
(ii) Explain how you know that lead(II) oxide is reduced in this reaction.

..... [1]

(iii) Explain why this reaction is described as endothermic.

..... [1]

(c) Lead nitrate solution reacts with sodium iodide solution.



Lead iodide is insoluble in water but the reactants and sodium nitrate are soluble. Draw a labelled diagram to explain how you can separate lead iodide from the rest of the reaction mixture.

[2]

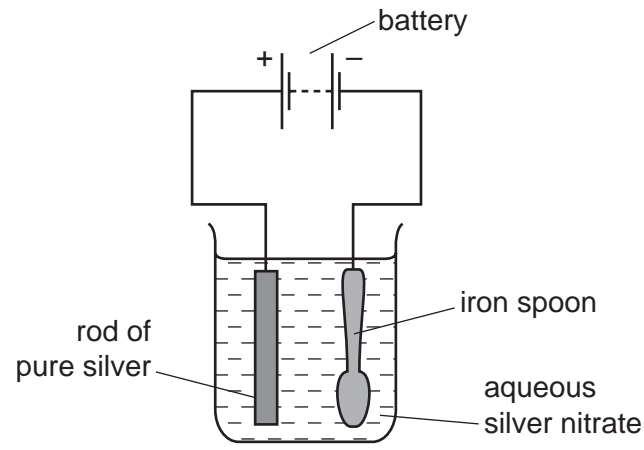
(d) Complete the table below to show the number of protons, electrons and neutrons in the isotope of lead  $^{204}_{82}\text{Pb}$ .

number of protons	
number of electrons	
number of neutrons	

[2]

[Total: 9]

7 The diagram below shows the apparatus used to electroplate a spoon with silver.



(a) Which is the anode?  
Put a ring around the correct answer in the list below.

- aqueous silver nitrate
- battery
- iron spoon
- rod of pure silver

[1]

(b) Describe what happens to the silver rod and the iron spoon during electroplating.

silver rod .....

iron spoon ..... [2]

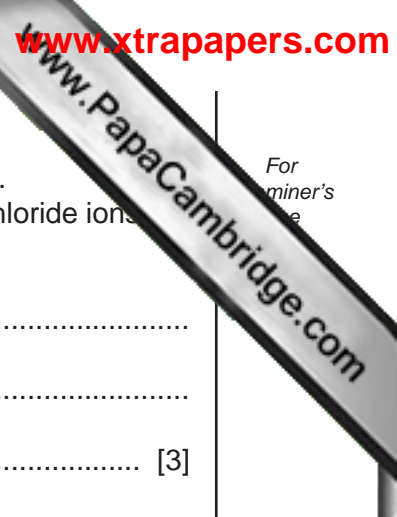
(c) Why are metal objects electroplated?

..... [1]

(d) During the electroplating, silver atoms are converted to silver ions.  
Which one of the following statements about this reaction is correct?  
Tick **one** box.

- Silver atoms gain electrons.
- Silver atoms lose neutrons.
- Silver atoms lose electrons.
- Silver atoms gain protons.

[1]



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(e) A student is given a slightly alkaline solution which contains chloride ions. Describe how the student could use aqueous silver nitrate to show that chloride ions are present in the solution.

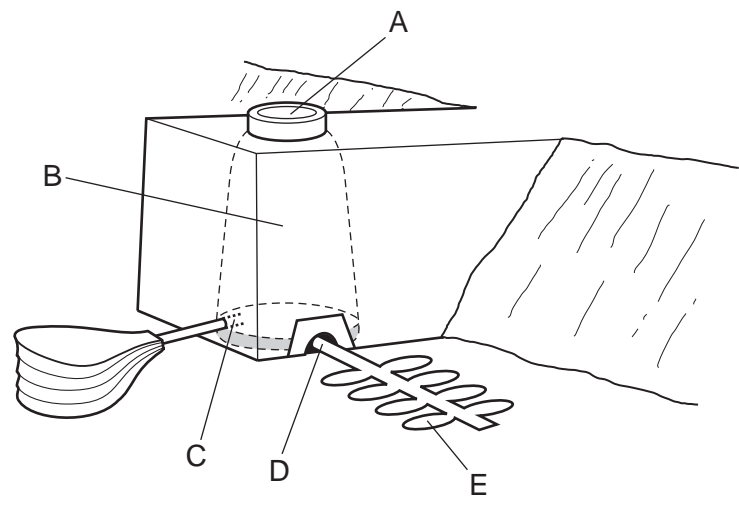
.....  
.....  
..... [3]

(f) Silver is a shiny metallic solid with a high melting point and boiling point. Describe two **other** physical properties of silver.

1 .....  
2 ..... [2]

[Total: 10]

8 The diagram shows a type of blast furnace built about 230 years ago. It was used to extract iron from iron ore.



(a) Which letter on the diagram shows

- (i) where the solid raw materials are put into the furnace, ..... [1]
- (ii) where air is blown into the furnace, ..... [1]
- (iii) where iron is removed from the furnace? ..... [1]

(b) Describe the main reactions occurring in a blast furnace for extracting iron from iron ore. In your answer, include

- the names of the raw materials used
- the main chemical reactions which occur
- relevant word equations.

.....

.....

.....

.....

.....

.....

.....

.....

.....

.....

.....

[5]

(c) Iron reacts with hydrochloric acid.

(i) Complete the word equation for this reaction.



[2]

(ii) Iron(II) ions are formed in this reaction.  
Describe a test for iron(II) ions.

test .....

result ..... [2]

(d) Steel is an alloy of iron.

Which one of the following statements about steel is correct?

Tick **one** box.

Steel is a mixture of iron with sulfur atoms.

Stainless steel is commonly used to make car bodies.

The physical properties of steel are exactly the same as those of iron.

Steel is made by blowing oxygen through the molten iron obtained from the blast furnace.

[1]

[Total: 13]



**DATA SHEET**  
**The Periodic Table of the Elements**

I		Group										VII		VIII		IX		X		XI		XII		XIII		XIV		XV		XVI		XVII		XVIII		IX																																																																																																																							
		II	III	IV	V	VI	VII	0																																																																																																																																																			
7 <b>Li</b> Lithium 3	9 <b>Be</b> Beryllium 4	1 <b>H</b> Hydrogen 1	11 <b>B</b> Boron 5	12 <b>C</b> Carbon 6	13 <b>Al</b> Aluminium 13	14 <b>Si</b> Silicon 14	15 <b>P</b> Phosphorus 15	16 <b>S</b> Sulfur 16	17 <b>Cl</b> Chlorine 17	18 <b>Ar</b> Argon 18	19 <b>F</b> Fluorine 9	20 <b>Ne</b> Neon 10	21 <b>Sc</b> Scandium 21	22 <b>Ti</b> Titanium 22	23 <b>V</b> Vanadium 23	24 <b>Cr</b> Chromium 24	25 <b>Mn</b> Manganese 25	26 <b>Fe</b> Iron 26	27 <b>Co</b> Cobalt 27	28 <b>Ni</b> Nickel 28	29 <b>Cu</b> Copper 29	30 <b>Zn</b> Zinc 30	31 <b>Ga</b> Gallium 31	32 <b>Ge</b> Germanium 32	33 <b>As</b> Arsenic 33	34 <b>Se</b> Selenium 34	35 <b>Br</b> Bromine 35	36 <b>Kr</b> Krypton 36	37 <b>Rb</b> Rubidium 37	38 <b>Sr</b> Strontium 38	39 <b>Y</b> Yttrium 39	40 <b>Zr</b> Zirconium 40	41 <b>Nb</b> Niobium 41	42 <b>Mo</b> Molybdenum 42	43 <b>Tc</b> Technetium 43	44 <b>Ru</b> Ruthenium 44	45 <b>Rh</b> Rhodium 45	46 <b>Pd</b> Palladium 46	47 <b>Ag</b> Silver 47	48 <b>Cd</b> Cadmium 48	49 <b>In</b> Indium 49	50 <b>Sn</b> Tin 50	51 <b>Sb</b> Antimony 51	52 <b>Te</b> Tellurium 52	53 <b>I</b> Iodine 53	54 <b>Xe</b> Xenon 54	55 <b>Cs</b> Caesium 55	56 <b>Ba</b> Barium 56	57 <b>La</b> Lanthanum 57	58-71 <b>Lanthanoid series</b>	72 <b>Hf</b> Hafnium 72	73 <b>Ta</b> Tantalum 73	74 <b>W</b> Tungsten 74	75 <b>Re</b> Rhenium 75	76 <b>Os</b> Osmium 76	77 <b>Ir</b> Iridium 77	78 <b>Pt</b> Platinum 78	79 <b>Au</b> Gold 79	80 <b>Hg</b> Mercury 80	81 <b>Tl</b> Thallium 81	82 <b>Pb</b> Lead 82	83 <b>Bi</b> Bismuth 83	84 <b>Po</b> Polonium 84	85 <b>At</b> Astatine 85	86 <b>Rn</b> Radon 86	87 <b>Fr</b> Francium 87	88 <b>Ra</b> Radium 88	89 <b>Ac</b> Actinium 89	†90-103 <b>Actinoid series</b>	90 <b>Th</b> Thorium 90	91 <b>Pa</b> Protactinium 91	92 <b>U</b> Uranium 92	93 <b>Np</b> Neptunium 93	94 <b>Pu</b> Plutonium 94	95 <b>Am</b> Americium 95	96 <b>Cm</b> Curium 96	97 <b>Bk</b> Berkelium 97	98 <b>Cf</b> Californium 98	99 <b>Es</b> Einsteinium 99	100 <b>Fm</b> Fermium 100	101 <b>Md</b> Mendelevium 101	102 <b>No</b> Nobelium 102	103 <b>Lr</b> Lawrencium 103	104 <b>Rf</b> Rutherfordium 104	105 <b>Db</b> Dubnium 105	106 <b>Sg</b> Seaborgium 106	107 <b>Bh</b> Bohrium 107	108 <b>Hs</b> Hassium 108	109 <b>Mt</b> Meitnerium 109	110 <b>Ds</b> Darmstadtium 110	111 <b>Rg</b> Roentgenium 111	112 <b>Cn</b> Copernicium 112	113 <b>Nh</b> Nihonium 113	114 <b>Fl</b> Flerovium 114	115 <b>Lv</b> Livermorium 115	116 <b>Ts</b> Tennessine 116	117 <b>Og</b> Oganesson 117	118 <b>Uu</b> Ununseptium 118	119 <b>Uub</b> Ununbium 119	120 <b>Uuq</b> Ununquadium 120	121 <b>Uubk</b> Ununbikium 121	122 <b>Uuqk</b> Ununquadium 122	123 <b>Uubk</b> Ununbikium 123	124 <b>Uuqk</b> Ununquadium 124	125 <b>Uubk</b> Ununbikium 125	126 <b>Uuqk</b> Ununquadium 126	127 <b>Uubk</b> Ununbikium 127	128 <b>Uuqk</b> Ununquadium 128	129 <b>Uubk</b> Ununbikium 129	130 <b>Uuqk</b> Ununquadium 130	131 <b>Uubk</b> Ununbikium 131	132 <b>Uuqk</b> Ununquadium 132	133 <b>Uubk</b> Ununbikium 133	134 <b>Uuqk</b> Ununquadium 134	135 <b>Uubk</b> Ununbikium 135	136 <b>Uuqk</b> Ununquadium 136	137 <b>Uubk</b> Ununbikium 137	138 <b>Uuqk</b> Ununquadium 138	139 <b>Uubk</b> Ununbikium 139	140 <b>Uuqk</b> Ununquadium 140	141 <b>Uubk</b> Ununbikium 141	142 <b>Uuqk</b> Ununquadium 142	143 <b>Uubk</b> Ununbikium 143	144 <b>Uuqk</b> Ununquadium 144	145 <b>Uubk</b> Ununbikium 145	146 <b>Uuqk</b> Ununquadium 146	147 <b>Uubk</b> Ununbikium 147	148 <b>Uuqk</b> Ununquadium 148	149 <b>Uubk</b> Ununbikium 149	150 <b>Uuqk</b> Ununquadium 150	151 <b>Uubk</b> Ununbikium 151	152 <b>Uuqk</b> Ununquadium 152	153 <b>Uubk</b> Ununbikium 153	154 <b>Uuqk</b> Ununquadium 154	155 <b>Uubk</b> Ununbikium 155	156 <b>Uuqk</b> Ununquadium 156	157 <b>Uubk</b> Ununbikium 157	158 <b>Uuqk</b> Ununquadium 158	159 <b>Uubk</b> Ununbikium 159	160 <b>Uuqk</b> Ununquadium 160	161 <b>Uubk</b> Ununbikium 161	162 <b>Uuqk</b> Ununquadium 162	163 <b>Uubk</b> Ununbikium 163	164 <b>Uuqk</b> Ununquadium 164	165 <b>Uubk</b> Ununbikium 165	166 <b>Uuqk</b> Ununquadium 166	167 <b>Uubk</b> Ununbikium 167	168 <b>Uuqk</b> Ununquadium 168	169 <b>Uubk</b> Ununbikium 169	170 <b>Uuqk</b> Ununquadium 170	171 <b>Uubk</b> Ununbikium 171	172 <b>Uuqk</b> Ununquadium 172	173 <b>Uubk</b> Ununbikium 173	174 <b>Uuqk</b> Ununquadium 174	175 <b>Uubk</b> Ununbikium 175

The volume of one mole of any gas is 24 dm<sup>3</sup> at room temperature and pressure (r.t.p.).

**Key**

a	<b>X</b>
b	
c	

a = relative atomic mass  
 X = atomic symbol  
 b = proton (atomic) number

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