



Cambridge International Examinations
Cambridge International General Certificate of Secondary Education

CANDIDATE
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CHEMISTRY

0620/23

Paper 2

May/June 2014

1 hour 15 minutes

Candidates answer on the Question Paper.

No Additional Materials are required.

READ THESE INSTRUCTIONS FIRST

Write your Centre number, candidate number and name in the spaces at the top of this page.

Write in dark blue or black pen.

You may use an HB pencil for any diagrams or graphs.

Do not use staples, paper clips, glue or correction fluid.

DO **NOT** WRITE IN ANY BARCODES.

Answer **all** questions.

Electronic calculators may be used.

A copy of the Periodic Table is printed on page 20.

You may lose marks if you do not show your working or if you do not use appropriate units.

At the end of the examination, fasten all your work securely together.

The number of marks is given in brackets [] at the end of each question or part question.

The syllabus is approved for use in England, Wales and Northern Ireland as a Cambridge International Level 1/Level 2 Certificate.

This document consists of **17** printed pages and **3** blank pages.



- 1 (a) Choose from the list of compounds below to answer the following questions.

aluminium oxide
calcium carbonate
calcium oxide
copper(II) sulfate
hydrogen chloride
potassium bromide
sodium chloride
sodium hydroxide

Each compound can be used once, more than once or not at all.

Which compound:

- (i) reacts with aqueous ammonia to form a light blue precipitate,
 [1]
- (ii) is formed by the decomposition of limestone,
 [1]
- (iii) forms an acidic solution when dissolved in water,
 [1]
- (iv) when electrolysed, gives a red-brown vapour at the anode,
 [1]
- (v) is an oxide of a metal in Group III of the Periodic Table,
 [1]
- (vi) is a transition element compound?
 [1]

- (b) Complete the following sentences about compounds using words from the list below.

| | | |
|-------------------|-------------------|----------------|
| chemically | different | fixed |
| mixed | physically | similar |

A compound is a substance which consists of two or more different elements
 combined together.

The properties of a compound are from those of the elements from which it
 is formed.

In a compound, the elements are combined in proportions. [3]

[Total: 9]

- 2 (a) Calcium chloride, CaCl_2 , is a salt.
Suggest the name of an acid and a base that would react together to make calcium chloride.

acid

base

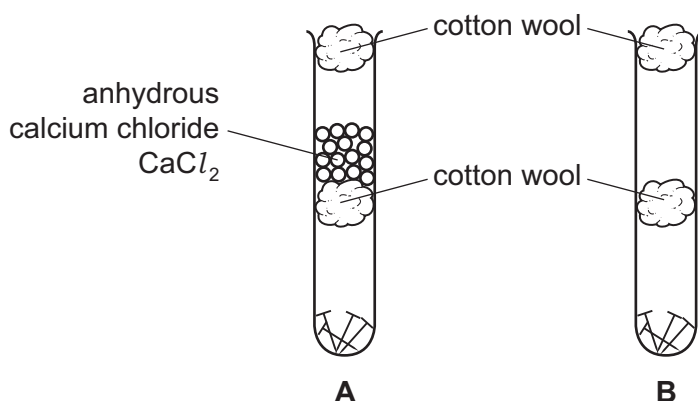
[2]

- (b) Calcium chloride absorbs water vapour.
When calcium chloride is heated, it loses its water of crystallisation.
Complete the symbol equation for this reaction. Include the sign for a reversible reaction.



[2]

- (c) A student put some clean iron nails in two test-tubes, as shown in the diagram. She then left the test-tubes for several weeks.



Explain why the nails in tube **A** did not rust but the nails in tube **B** rusted.

.....

[2]

- (d) Rust is hydrated iron(III) oxide.
What does the (III) in iron(III) oxide refer to?
Tick **one** box.

the oxidation state of the oxygen

the oxidation state of the iron

the number of atoms of oxygen in a formula unit of iron(III) oxide

the number of water molecules in the hydrated iron oxide

[1]

(e) (i) The table describes the ease of reduction of some metal oxides with carbon monoxide.

| | |
|-----------------|--|
| lead oxide | moderate heating to about 200 °C needed |
| iron oxide | high temperature furnace at 750 °C needed |
| magnesium oxide | temperatures above 1000 °C needed |
| zinc oxide | very high temperature furnace at 900 °C needed |

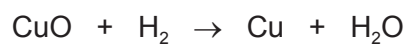
Put these metals in order of their reactivity with carbon monoxide.

least reactive $\xrightarrow{\hspace{15em}}$ most reactive

| | | | |
|--|--|--|--|
| | | | |
|--|--|--|--|

[2]

(ii) Some metal oxides can be reduced by heating with hydrogen gas.



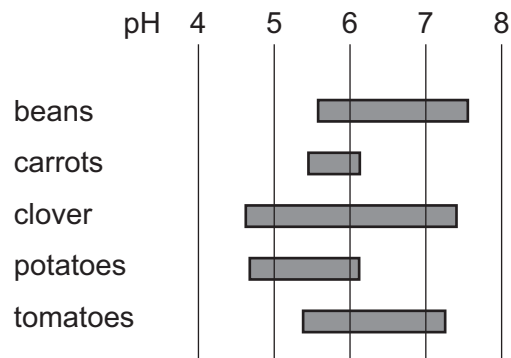
Explain how this equation shows that copper oxide is being reduced.

..... [1]

[Total: 10]

5

3 The diagram shows the best pH ranges for growing different plants.



(a) (i) Which **two** plants grow best in acidic conditions **only**?

..... and [1]

(ii) Which pH shown in the diagram above represents a neutral pH?

..... [1]

(b) (i) Explain why lime is added to acidic soils.

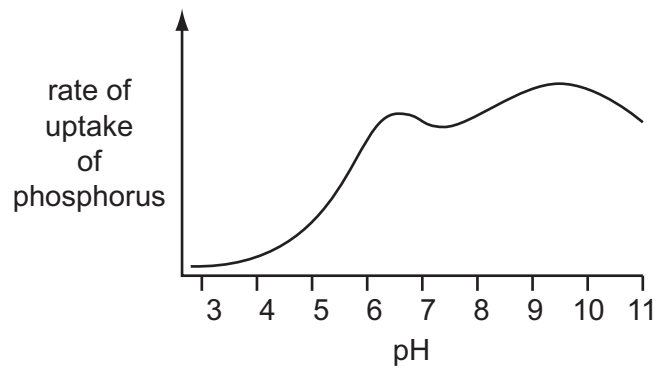
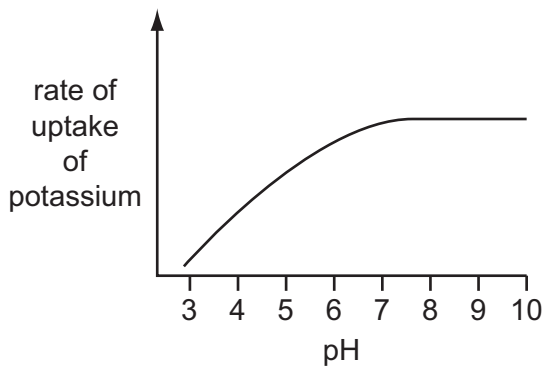
.....
 [2]

(ii) Farmers fertilise soil by adding compounds containing ammonium salts.
 Explain why adding lime to fertilised soil may cause a loss of nitrogen from the soil.

.....

 [3]

(c) The graphs below show the rate of uptake of potassium and phosphate ions by plant roots at different pH values.



(i) Describe the effect of pH on the rate of uptake of potassium by plant roots.

.....
 [2]

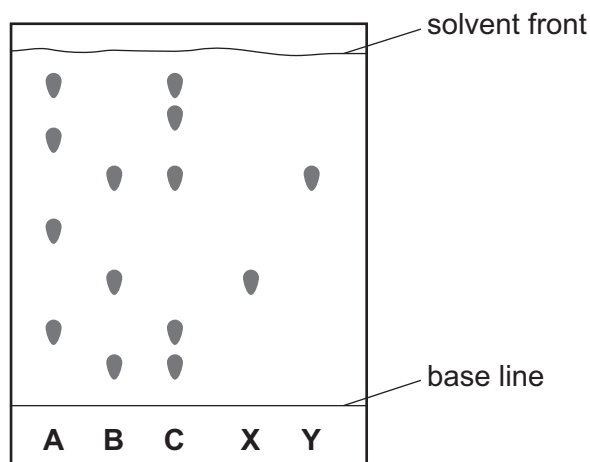
(ii) At which pH value is the rate of uptake of phosphorus by plant roots the highest?

..... [1]

[Total: 10]

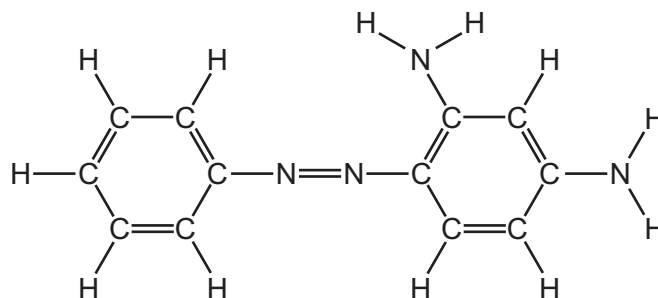
4 Chromatography is used to separate a mixture of coloured dyes.

- (a) Three different dye mixtures, **A**, **B** and **C**, were spotted onto a piece of chromatography paper. Two pure dyes, **X** and **Y**, were also spotted onto the same piece of paper. The diagram below shows the results of the chromatography.



- (i) State the name of a piece of apparatus that could be used to spot the dyes onto the paper.
 [1]
- (ii) Suggest why the base line was drawn in pencil and not in ink.
 [1]
- (iii) Which dye mixture contains **both** dye **X** and dye **Y**?
 [1]
- (iv) Which dye mixture does **not** contain dye **X** or dye **Y**?
 [1]
- (v) In which mixture, **A**, **B** or **C**, has the greatest number of dyes been separated?
 [1]

(b) The structure of the dye chrysoidine G is shown below.



(i) How many nitrogen atoms are there in a molecule of chrysoidine G?

..... [1]

(ii) Complete the table below to calculate the relative molecular mass of chrysoidine G.

| type of atom | number of atoms | atomic mass | |
|--------------|-----------------|-------------|----------------------|
| carbon | 12 | 12 | $12 \times 12 = 144$ |
| hydrogen | | | |
| nitrogen | | | |

relative molecular mass = [2]

(c) The fibres in the chromatography paper are polymers.

(i) What is meant by the term *polymer*?

.....
 [1]

(ii) State the chemical name of the polymer formed from ethene.

..... [1]

[Total: 10]

5 The table shows some properties of the first four carboxylic acids.

| acid | molecular formula | melting point /°C | boiling point /°C | density in g/cm ³ |
|----------------|--|-------------------|-------------------|------------------------------|
| methanoic acid | CH ₂ O ₂ | +10 | +101 | 1.22 |
| ethanoic acid | C ₂ H ₄ O ₂ | +17 | +118 | 1.05 |
| propanoic acid | C ₃ H ₆ O ₂ | -21 | | 0.99 |
| butanoic acid | C ₄ H ₈ O ₂ | -4 | +166 | |

(a) (i) How does the boiling point of these carboxylic acids vary with the number of carbon atoms?

..... [1]

(ii) Suggest a value for:

the boiling point of propanoic acid, °C

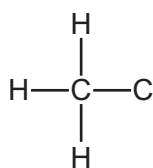
the density of butanoic acid. g/cm³
[2]

(iii) Is butanoic acid a solid, liquid or gas at room temperature?

Use the data in the table to explain your answer.

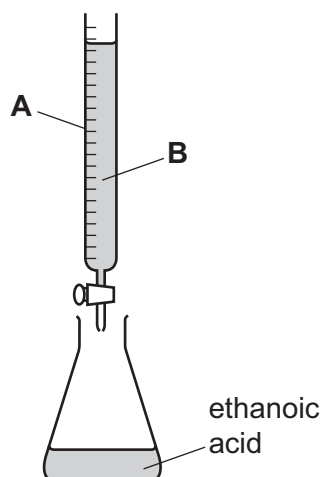
.....
..... [1]

(b) Complete the diagram below to show the structure of ethanoic acid.
Show all atoms and bonds.



[1]

- (c) The concentration of ethanoic acid can be determined by titration using the apparatus shown below.



- (i) State the name of the piece of glassware labelled **A**.

..... [1]

- (ii) Liquid **B** is an alkali.
Which **one** of the following compounds is also an alkali?
Put a ring around the correct answer.

calcium carbonate

calcium sulfate

sodium chloride

sodium hydroxide

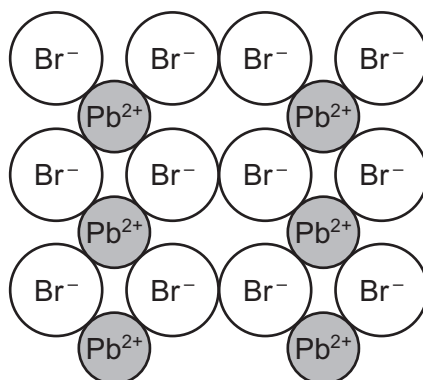
[1]

- (iii) Describe how you would carry out this titration.

.....
.....
.....
..... [2]

[Total: 9]

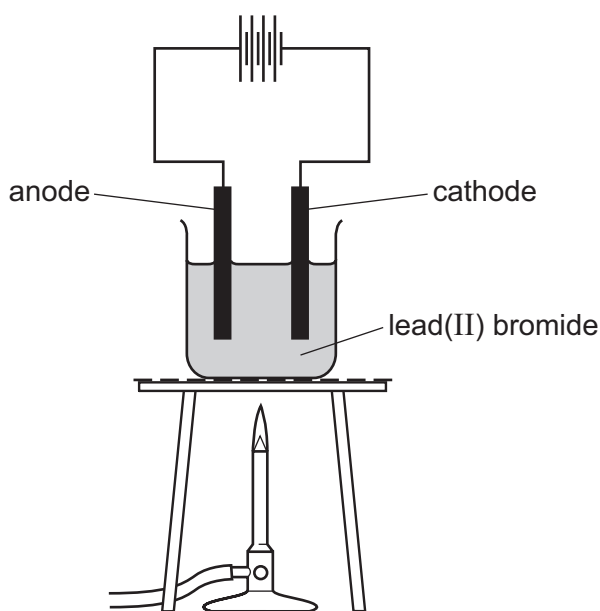
- 6 Lead(II) bromide is a white solid.
Part of the structure of lead(II) bromide is shown below.



- (a) Deduce the simplest formula for lead(II) bromide.

..... [1]

- (b) A student electrolysed lead(II) bromide in a fume cupboard using the apparatus shown below.



- (i) Why is heat needed for this electrolysis?

..... [1]

- (ii) Suggest the name of a substance that could be used for the electrodes.

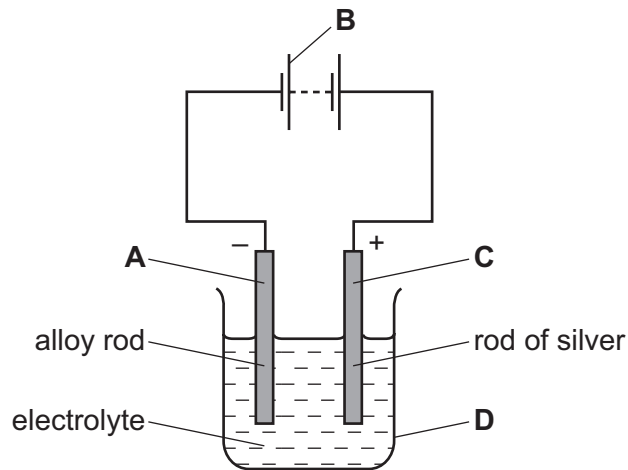
..... [1]

- (iii) State the name of the products of electrolysis at:

the anode,

the cathode. [1]

(c) Items can be electroplated with silver using the apparatus shown below.



(i) On the diagram, which letter, **A**, **B**, **C** or **D**, is the cathode?

..... [1]

(ii) What would you observe during the experiment at the:

positive electrode,

.....

negative electrode?

.....

[2]

(iii) The electrolyte used is aqueous silver cyanide, AgCN .
Calculate the relative formula mass of silver cyanide.
You must show all your working.

[2]

[Total: 9]

- 7 Dmitri Mendeleev published his first Periodic Table in 1869. Part of this table is shown below.

| | | | |
|--------|----------|-----------|-----------|
| | | | Ti = 50 |
| | | | V = 51 |
| | | | Cr = 52 |
| | | | Mn = 55 |
| | | | Fe = 56 |
| | | | Co = 59 |
| H = 1 | | | Cu = 63.4 |
| | Be = 9.4 | Mg = 24 | Zn = 65.2 |
| | B = 11 | Al = 27.4 | ? |
| | C = 12 | Si = 28 | ? |
| | N = 14 | P = 31 | As = 75 |
| | O = 16 | S = 32 | Se = 79.4 |
| | F = 19 | Cl = 35.5 | Br = 80 |
| Li = 7 | Na = 23 | K = 39 | Rb = 85.4 |

- (a) (i) What differences are there between Mendeleev's table and the Periodic Table we use today?

.....

.....

.....

.....

..... [4]

- (ii) State the names of any **two** elements in the table above which exist as diatomic molecules.

..... and [1]

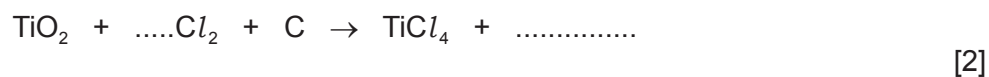
- (b) Titanium is a transition element. Sodium is a metal in Group I of the Periodic Table. State **three** differences in the physical properties of titanium and sodium.

1

2

3 [3]

- (c) Titanium(IV) oxide reacts with a mixture of chlorine and carbon. The products are titanium(IV) chloride, $TiCl_4$, and a gas which turns limewater milky. Complete the symbol equation for this reaction.



- (d) Titanium is extracted from titanium(IV) chloride by reduction with molten sodium in the presence of argon. Suggest why this reaction is carried out in the presence of argon.

.....
..... [2]

[Total: 12]

(d) What would you observe when a piece of blue cobalt chloride paper is dipped into water?

..... [1]

(e) Describe how impure water is treated so that it can be used for drinking.

.....
..... [2]

[Total: 11]

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