



Cambridge Assessment International Education
Cambridge International General Certificate of Secondary Education

COMBINED SCIENCE

0653/23

Paper 2 Multiple Choice (Extended)

October/November 2019

45 minutes

Additional Materials: Multiple Choice Answer Sheet
Soft clean eraser
Soft pencil (type B or HB is recommended)



READ THESE INSTRUCTIONS FIRST

Write in soft pencil.

Do not use staples, paper clips, glue or correction fluid.

Write your name, centre number and candidate number on the Answer Sheet in the spaces provided unless this has been done for you.

DO NOT WRITE IN ANY BARCODES.

There are **forty** questions on this paper. Answer **all** questions. For each question there are four possible answers **A, B, C** and **D**.

Choose the **one** you consider correct and record your choice in **soft pencil** on the separate Answer Sheet.

Read the instructions on the Answer Sheet very carefully.

Each correct answer will score one mark. A mark will not be deducted for a wrong answer.

Any rough working should be done in this booklet.

A copy of the Periodic Table is printed on page 20.

Electronic calculators may be used.

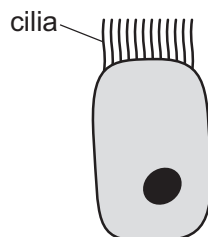
This document consists of **17** printed pages and **3** blank pages.

1 A biologist keeps a potted plant in a laboratory.

Which feature of the potted plant shows that it is a living organism?

- A It grows larger over time.
- B It has green leaves.
- C The compost in the pot dries after he waters it.
- D The stems contain xylem.

2 The diagram shows a ciliated cell.



Which row shows where ciliated cells are found in the human gas exchange system and their correct function?

	location of ciliated cells		function of ciliated cells	
	bronchi	trachea	move mucus away from lungs	move mucus towards lungs
A	✓	✓	✓	x
B	✓	✓	x	✓
C	✓	x	✓	x
D	x	✓	x	✓

3 What is the word equation for photosynthesis?

- A carbon dioxide + oxygen → glucose + water
- B carbon dioxide + water → glucose + oxygen
- C glucose + oxygen → carbon dioxide + water
- D glucose + water → carbon dioxide + oxygen

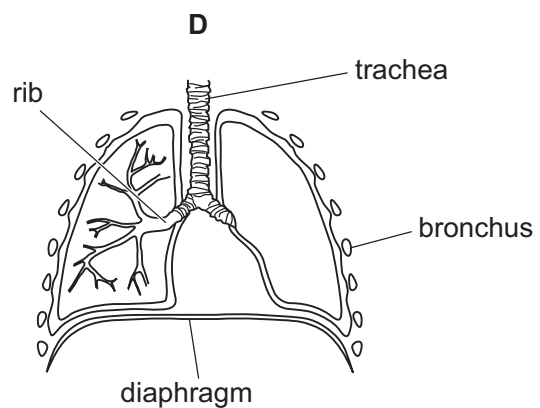
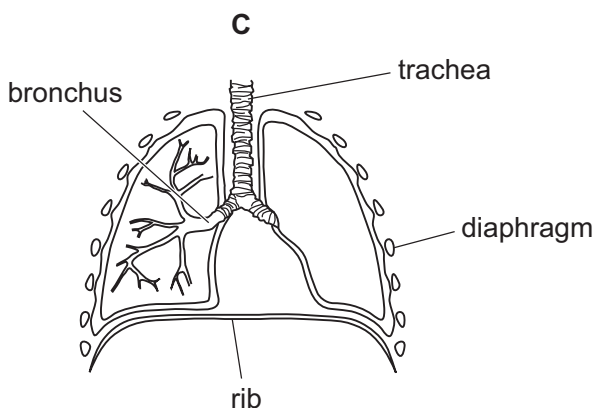
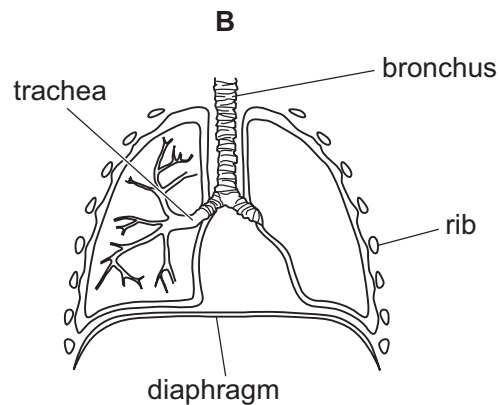
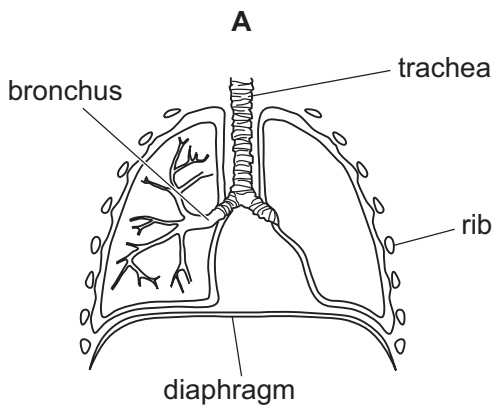
- 4 1 cm³ of substance **X** is added to 10 cm³ starch suspension and mixed. Food tests are carried out immediately after mixing and again after an hour.

The results of the tests are shown in the table.

test reagent	colour of solution after mixing	colour of solution after one hour
Benedict's solution	blue	orange
iodine solution	blue/black	brown

What is substance **X**?

- A amylase
 B protease
 C lipase
 D sugar
- 5 Which diagram is correctly labelled?



6 Which statement about aerobic respiration is correct?

- A It exchanges gases through the walls of the alveoli.
- B It expels carbon dioxide from the lungs.
- C It only produces carbon dioxide and energy.
- D It uses oxygen to release energy from glucose.

7 Which are absorbed from the alimentary canal into the blood?

- 1 fibre
- 2 glucose
- 3 vitamin C

- A 1 and 2 only B 1 and 3 only C 2 and 3 only D 1, 2 and 3

8 Shoots were grown in different light conditions.

Some shoots had their tips covered with foil.

	shoot tip	direction of light
1	covered	from all around
2	covered	from one direction
3	uncovered	from all around
4	uncovered	from one direction

Which shoots would grow straight upwards?

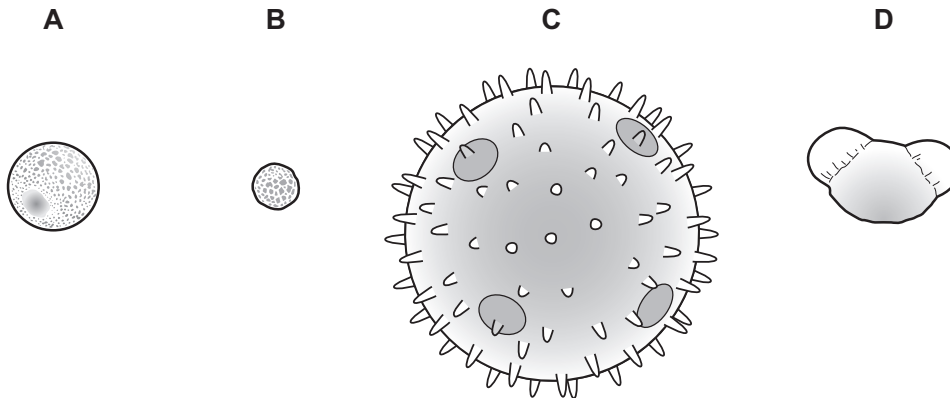
- A 1, 2 and 3 B 1 and 3 only C 3 and 4 only D 3 only

9 Which statement about sexual reproduction is **always** correct?

- A It involves only one parent.
- B It involves the fusion of nuclei.
- C It produces genetically identical offspring.
- D It takes place only in animals.

10 The diagram shows four pollen grains.

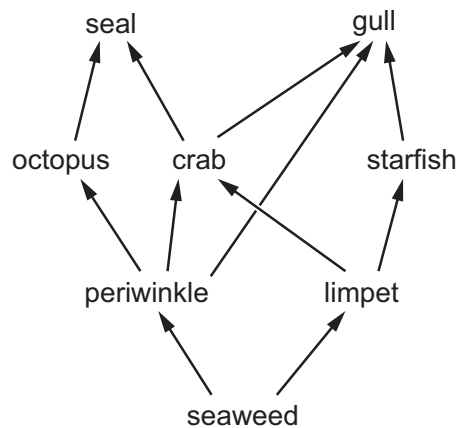
Which pollen grain is most likely to be distributed by an animal?



11 Which statement about human gametes is correct?

- A Sperm cells are much larger than egg cells.
- B Sperm cells are produced in smaller numbers than egg cells.
- C Sperm cells have a jelly coating that changes after fertilisation.
- D The flagellum is an adaptive feature of a sperm cell.

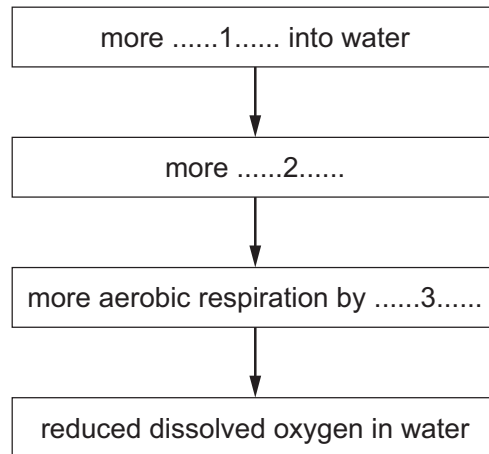
12 The diagram shows a food web.



Which organism is found in more than one trophic level?

- A crab
- B gull
- C octopus
- D starfish

13 The flow diagram shows some stages in the eutrophication of a pond.

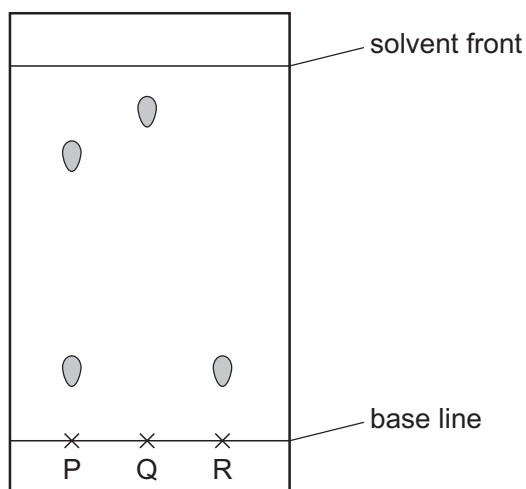


Which words complete gaps 1, 2 and 3?

	1	2	3
A	decomposers	nitrites	producers
B	decomposers	producers	nitrites
C	nitrites	producers	decomposers
D	nitrites	decomposers	producers

14 Chromatography is carried out on three solutions P, Q and R.

The chromatogram obtained is shown.



Which statement is **not** correct?

- A P contains at least two substances.
 - B Q contains the substance with the highest R_f value.
 - C R is insoluble in the solvent.
 - D P, Q and R together may contain only three substances.
- 15 Which substance is a single compound?
- A air
 - B oxygen
 - C petroleum
 - D water

16 The fertiliser ammonium sulfate has the formula $(\text{NH}_4)_2\text{SO}_4$.

How many atoms of each element are present in the formula?

	number of hydrogen atoms	number of nitrogen atoms	number of oxygen atoms	number of sulfur atoms
A	4	1	1	1
B	4	2	4	1
C	8	1	4	1
D	8	2	4	1

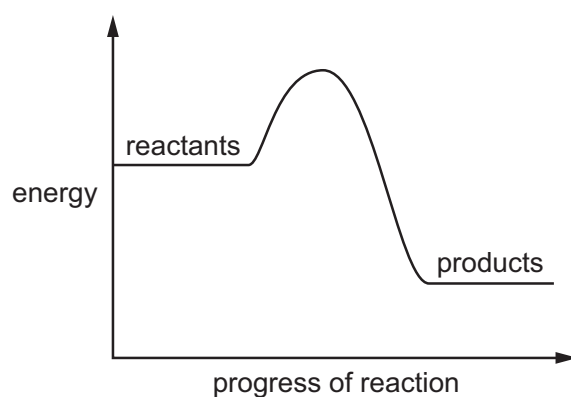
17 Element X is a non-metal used in the treatment of the water supply.

It is made during the electrolysis of a metal salt.

What is the colour of X and at which electrode is it made?

	colour	electrode
A	red	anode
B	red	cathode
C	yellow-green	anode
D	yellow-green	cathode

18 An energy level diagram for a reaction is shown.



Which row describes the energy transfer and the type of energy change for this reaction?

	energy transfer	energy change
A	energy is absorbed by reactants	endothermic
B	energy is absorbed by reactants	exothermic
C	energy is released to surroundings	endothermic
D	energy is released to surroundings	exothermic

- 19 Calcium carbonate reacts with 50 cm³ hydrochloric acid.

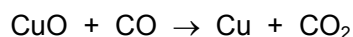
The carbon dioxide produced is collected in a gas syringe.

The experiment is done four times using concentrated or dilute hydrochloric acid and using 5 g calcium carbonate in powder or lump form.

Which experiment takes the longest time to collect 10 cm³ of gas?

	calcium carbonate	hydrochloric acid
A	lumps	concentrated
B	lumps	dilute
C	powder	concentrated
D	powder	dilute

- 20 The equation for a reaction is shown.



Which statement about this reaction is correct?

- A CO acts as a reducing agent.
 - B CO₂ is reduced.
 - C Cu is oxidised.
 - D CuO acts as a reducing agent.
- 21 Copper sulfate is a soluble salt which is prepared by reacting insoluble copper oxide with dilute sulfuric acid.

Which statement about the preparation of copper sulfate crystals is **not** correct?

- A After the reaction, the mixture is filtered and copper sulfate solution is collected.
 - B Excess copper oxide is used to ensure that all the acid is used up.
 - C The final solution is heated so that all the water boils off.
 - D The mixture of copper oxide and dilute sulfuric acid is heated to speed up the reaction.
- 22 Which statement about alloys is correct?
- A They are made from metals because metals are poor electrical conductors.
 - B They are mixtures of compounds that contain metals.
 - C They have all the same properties as the metals from which they are made.
 - D They have different properties to the metals from which they are made.

23 Which equation does **not** represent a reaction that takes place in the blast furnace?

- A $C + O_2 \rightarrow CO_2$
- B $C + CO_2 \rightarrow 2CO$
- C $2Fe + CO_2 \rightarrow 2FeO + C$
- D $Fe_2O_3 + 3CO \rightarrow 2Fe + 3CO_2$

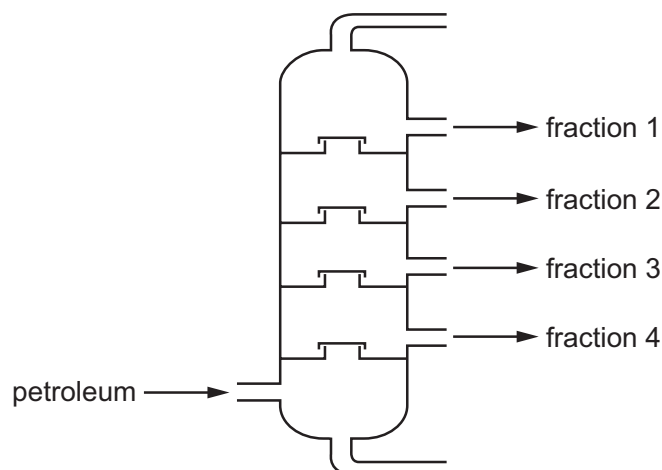
24 Which row describes the percentage composition of clean air?

	carbon dioxide	nitrogen	noble gases	oxygen
A	less than 1	78	less than 1	21
B	less than 1	78	21	less than 1
C	21	less than 1	less than 1	78
D	78	less than 1	less than 1	21

25 Which two gases cause an enhanced greenhouse effect when their concentrations in the atmosphere increase?

- A carbon monoxide and carbon dioxide
- B carbon dioxide and methane
- C methane and sulfur dioxide
- D sulfur dioxide and carbon monoxide

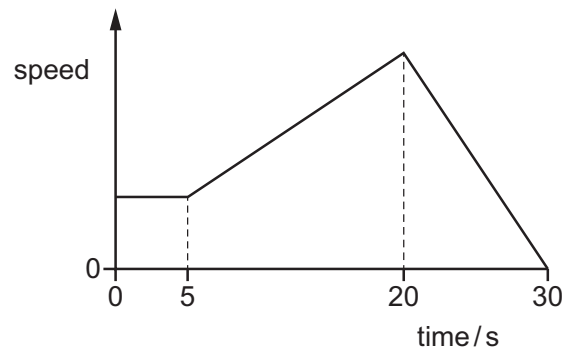
26 A simple fractionating column is shown.



Which statement about the fractions is correct?

- A Fraction 1 contains compounds with the highest boiling points.
 - B Fraction 2 contains larger hydrocarbon molecules than fraction 3.
 - C Fraction 3 is more viscous than fraction 4.
 - D Fraction 4 is the least flammable.
- 27 What is a typical property of alkanes?
- A They are catalysts.
 - B They burn in air.
 - C They can be neutralised.
 - D They react endothermically.

- 28 The graph shows how the speed of a car changes with time. The car travels at constant speed, then accelerates, and finally brakes to a stop.



The car travels 60 m while it brakes to a stop.

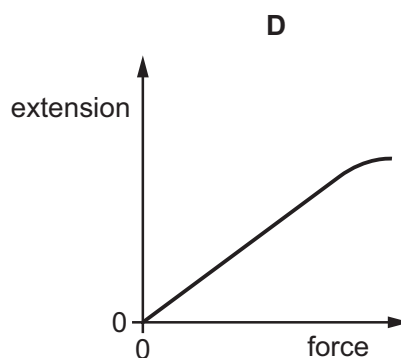
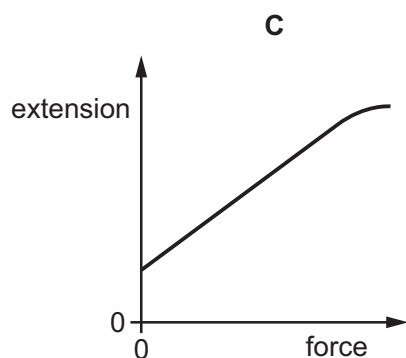
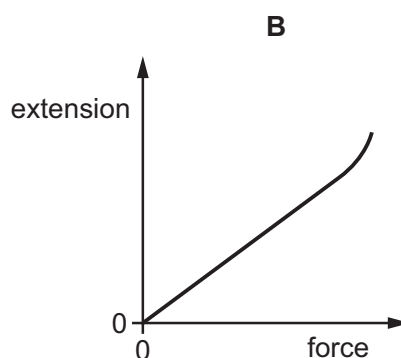
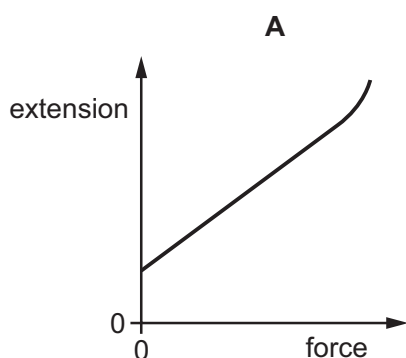
What is the average speed of the car while it is braking?

- A** 3.0 m/s **B** 4.0 m/s **C** 6.0 m/s **D** 12 m/s
- 29 Which of these bodies has a resultant force acting on it?
- A** a book at rest on a table
- B** a car travelling up a hill in a straight line at constant speed
- C** a football moving upwards freely after being kicked
- D** a parachutist descending vertically at constant speed

30 The force acting on a spring is gradually increased from 0 N.

The spring eventually passes its limit of proportionality.

Which graph shows how the extension of the spring changes as the force increases?



31 Some energy resources are less reliable than others.

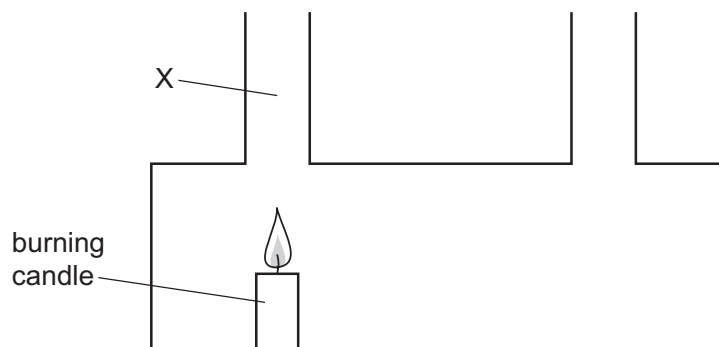
Which type of power station **cannot** produce electricity at all times?

- A coal-fired power station
- B geothermal power station
- C hydroelectric power station
- D solar power station

32 Which statement about the molecules in a gas is correct?

- A They are closer together than those in solids.
- B They are further apart than those in liquids.
- C They are **not** free to move around.
- D They are packed together in a regular pattern.

- 33 The equipment shown is used to demonstrate convection in air. Point X is labelled.



Which row describes and explains the movement of the air at X?

	movement of air at X	explanation
A	downwards	air becomes less dense when heated
B	downwards	air becomes more dense when heated
C	upwards	air becomes less dense when heated
D	upwards	air becomes more dense when heated

- 34 What type of wave is a sound wave and in which direction do air particles vibrate as the wave passes through the air?

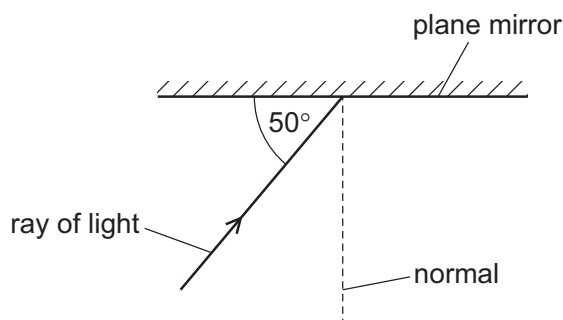
	type of wave	direction of vibration
A	longitudinal	parallel to wave direction
B	longitudinal	perpendicular to wave direction
C	transverse	parallel to wave direction
D	transverse	perpendicular to wave direction

- 35 A boy plays a series of musical notes of increasing frequency on a violin. As the frequency of the note increases, he plays the notes more loudly.

How do the amplitude and the wavelength of the sound waves change?

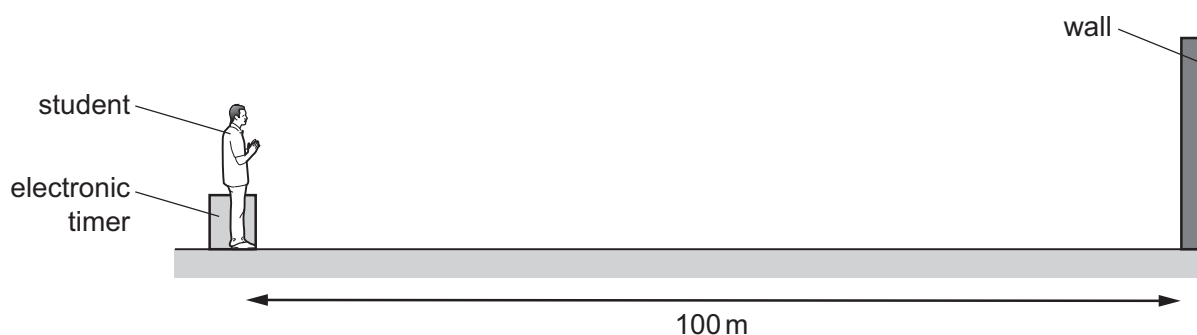
	amplitude	wavelength
A	decreases	decreases
B	decreases	increases
C	increases	decreases
D	increases	increases

- 36 The diagram shows light striking a plane mirror.



What is the angle of reflection of the ray when it is reflected from the mirror?

- A 40° B 50° C 80° D 100°
- 37 A student measures the speed of sound. He claps his hands and the sound reflects from a wall that is 100 m away from him.

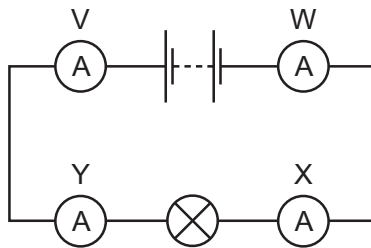


An electronic timer next to the student detects the echo of the sound 0.60 s after it is made.

Which calculation gives the speed of sound?

- A $\frac{200}{0.30}$ m/s B $\frac{200}{0.60}$ m/s C $\frac{100}{0.60}$ m/s D $\frac{100}{1.2}$ m/s
- 38 A piece of wire has a resistance of 8.0Ω .
- The length of the wire is doubled and the diameter of the wire is halved.
- What is the new resistance of the wire?
- A 2.0Ω B 4.0Ω C 8.0Ω D 64Ω

39 Four ammeters V, W, X and Y are connected in the circuit shown.



Which ammeters have the same reading as each other?

- A V and W only
 - B V and Y only
 - C X and Y only
 - D V, W, X and Y
- 40 There is a current I in a resistor and a potential difference V across it.

Which equation gives the energy E transferred by the resistor in a time t ?

- A $E = \frac{I}{Vt}$
- B $E = \frac{V}{It}$
- C $E = \frac{t}{VI}$
- D $E = IVt$

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The Periodic Table of Elements

		Group																
I	II	III	IV	V	VI	VII	VIII					VIII						
3 Li lithium 7	4 Be beryllium 9	1 H hydrogen 1	6 C carbon 12	7 N nitrogen 14	8 O oxygen 16	9 F fluorine 19	2 He helium 4					10 Ne neon 20						
11 Na sodium 23	12 Mg magnesium 24	13 Al aluminium 27	14 Si silicon 28	15 P phosphorus 31	16 S sulfur 32	17 Cl chlorine 35.5	18 Ar argon 40					36 Kr krypton 84						
19 K potassium 39	20 Ca calcium 40	21 Sc scandium 45	22 Ti titanium 48	23 V vanadium 51	24 Cr chromium 52	25 Mn manganese 55	26 Fe iron 56	27 Co cobalt 59	28 Ni nickel 59	29 Cu copper 64	30 Zn zinc 65	31 Ga gallium 70	32 Ge germanium 73	33 As arsenic 75	34 Se selenium 79	35 Br bromine 80	36 Kr krypton 84	
37 Rb rubidium 85	38 Sr strontium 88	39 Y yttrium 89	40 Zr zirconium 91	41 Nb niobium 93	42 Mo molybdenum 96	43 Tc technetium —	44 Ru ruthenium 101	45 Rh rhodium 103	46 Pd palladium 106	47 Ag silver 108	48 Cd cadmium 112	49 In indium 115	50 Sn tin 119	51 Sb antimony 122	52 Te tellurium 128	53 I iodine 127	54 Xe xenon 131	86 Rn radon —
55 Cs caesium 133	56 Ba barium 137	57–71 lanthanoids	72 Hf hafnium 178	73 Ta tantalum 181	74 W tungsten 184	75 Re rhenium 186	76 Os osmium 190	77 Ir iridium 192	78 Pt platinum 195	79 Au gold 197	80 Hg mercury 201	81 Tl thallium 204	82 Pb lead 207	83 Bi bismuth 209	84 Po polonium —	85 At astatine —	86 Rn radon —	
87 Fr francium —	88 Ra radium —	89–103 actinoids	104 Rf rutherfordium —	105 Db dubnium —	106 Sg seaborgium —	107 Bh bohrium —	108 Hs hassium —	109 Mt meitnerium —	110 Ds darmstadtium —	111 Rg roentgenium —	112 Cn copernicium —	114 Fl flerovium —	116 Lv livermorium —	—	—	—	—	

Key

atomic number
atomic symbol
name
relative atomic mass

lanthanoids	57 La lanthanum 139	58 Ce cerium 140	59 Pr praseodymium 141	60 Nd neodymium 144	61 Pm promethium —	62 Sm samarium 150	63 Eu europium 152	64 Gd gadolinium 157	65 Tb terbium 159	66 Dy dysprosium 163	67 Ho holmium 165	68 Er erbium 167	69 Tm thulium 169	70 Yb ytterbium 173	71 Lu lutetium 175
actinoids	89 Ac actinium —	90 Th thorium 232	91 Pa protactinium 231	92 U uranium 238	93 Np neptunium —	94 Pu plutonium —	95 Am americium —	96 Cm curium —	97 Bk berkelium —	98 Cf californium —	99 Es einsteinium —	100 Fm fermium —	101 Md mendelevium —	102 No nobelium —	103 Lr lawrencium —

The volume of one mole of any gas is 24 dm³ at room temperature and pressure (r.t.p.).