



Cambridge International Examinations
Cambridge International General Certificate of Secondary Education

CO-ORDINATED SCIENCES

0654/13

Paper 1 Multiple Choice

May/June 2014

45 minutes

Additional Materials: Multiple Choice Answer Sheet
 Soft clean eraser
 Soft pencil (type B or HB is recommended)

READ THESE INSTRUCTIONS FIRST

Write in soft pencil.

Do not use staples, paper clips, glue or correction fluid.

Write your name, Centre number and candidate number on the Answer Sheet in the spaces provided unless this has been done for you.

DO NOT WRITE IN ANY BARCODES.

There are **forty** questions on this paper. Answer **all** questions. For each question there are four possible answers **A, B, C** and **D**.

Choose the **one** you consider correct and record your choice in **soft pencil** on the separate Answer Sheet.

Read the instructions on the Answer Sheet very carefully.

Each correct answer will score one mark. A mark will not be deducted for a wrong answer.

Any rough working should be done in this booklet.

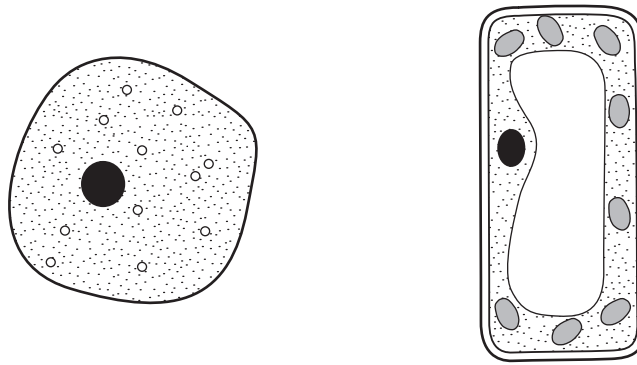
A copy of the Periodic Table is printed on page 16.

Electronic calculators may be used.

This document consists of **15** printed pages and **1** blank page.

2

1 The diagram shows two different cells.



Which feature do they both have?

- A cell membrane
- B cell wall
- C central vacuole
- D chloroplasts

2 Which rows correctly match characteristics of living things with their descriptions?

	characteristic	description
1	excretion	removing the waste products of metabolism
2	growth	making more living things of the same type
3	nutrition	taking in or producing food
4	respiration	obtaining energy from food

- A 1, 2 and 4 B 1, 3 and 4 C 1 and 3 only D 2 and 4 only

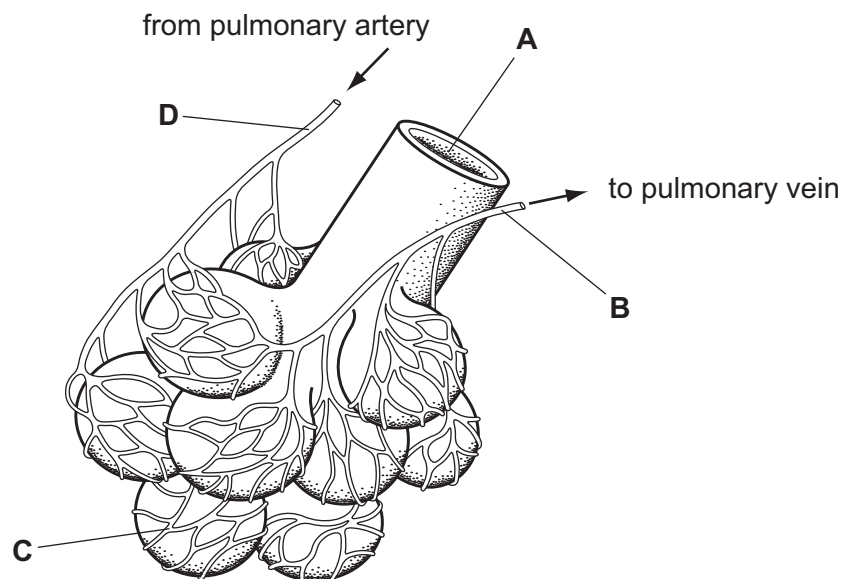
3 A species of bacterium lives in acidic, hot springs at a temperature of 90 °C.

Which conditions will best suit the enzymes of this bacterium?

- A 30 °C and pH 4
- B 30 °C and pH 9
- C 80 °C and pH 4
- D 80 °C and pH 9

- 4 Why are green plants called producers?
- A They can make oxygen from sunlight.
 - B They form organic nutrients from simple substances.
 - C They have cells containing chlorophyll.
 - D They produce starch from sugar.
- 5 In the maintenance of body temperature, which response does **not** need energy from respiration?
- A secretion of sweat
 - B shivering
 - C vasoconstriction
 - D vasodilation
- 6 How does oxygen pass from the alveoli to the blood capillaries in the lungs?
- A diffusion
 - B evaporation
 - C secretion
 - D transpiration
- 7 The diagram shows some of the structures in a human lung.

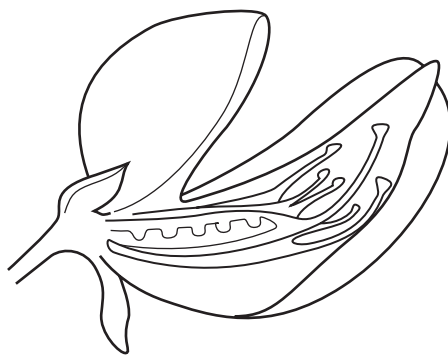
Where is the oxygen concentration lowest?



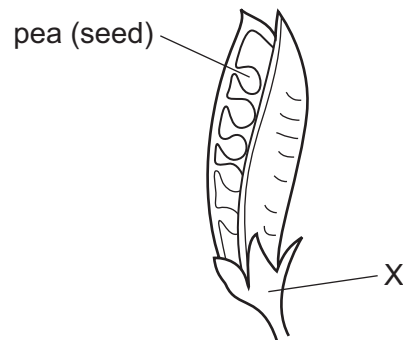
- 8 A plant shoot grows towards a light source.

This an example of what?

- A geotropism
 - B homeostasis
 - C photosynthesis
 - D phototropism
- 9 What is a function of adrenaline?
- A to increase the concentration of blood sugar
 - B to raise the level of oxygen in the blood
 - C to reduce the rate of heart beat
 - D to remove urea from the blood
- 10 The diagram shows the flower of a pea plant and the fruit that develops from the flower after fertilisation.



flower

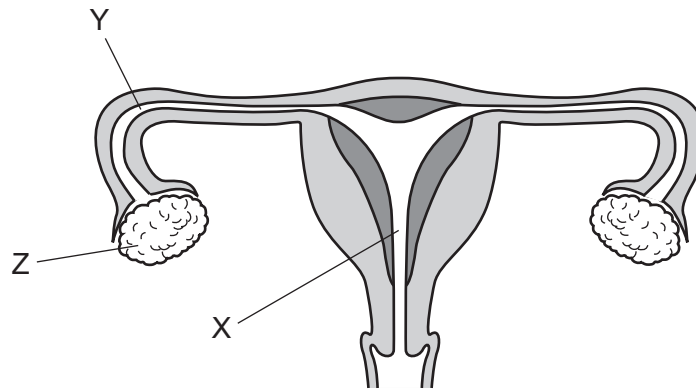


fruit

Which part of the flower becomes part X on the fruit?

- A ovary
 - B sepal
 - C stamen
 - D stigma
- 11 What is **not** produced by artificial selection?
- A bacteria with antibiotic resistance
 - B cows with high milk yield
 - C sheep with thick wool
 - D wheat with resistance to disease

12 The diagram shows the female reproductive system.



Which structures are the ovary and the oviduct?

	ovary	oviduct
A	X	Y
B	X	Z
C	Z	X
D	Z	Y

13 The diagram shows a food chain.

oak tree → insect → small bird → hawk

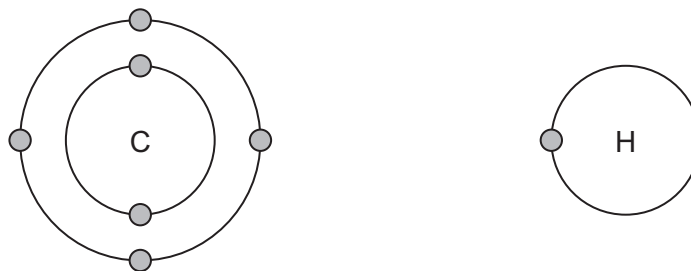
Which statement describes a member of this food chain?

- A** The oak tree is a consumer.
- B** The insect is a producer.
- C** The small bird is a consumer.
- D** The hawk is a producer.

14 Which process is used to separate the coloured compounds in chlorophyll?

- A** chromatography
- B** distillation
- C** evaporation
- D** filtration

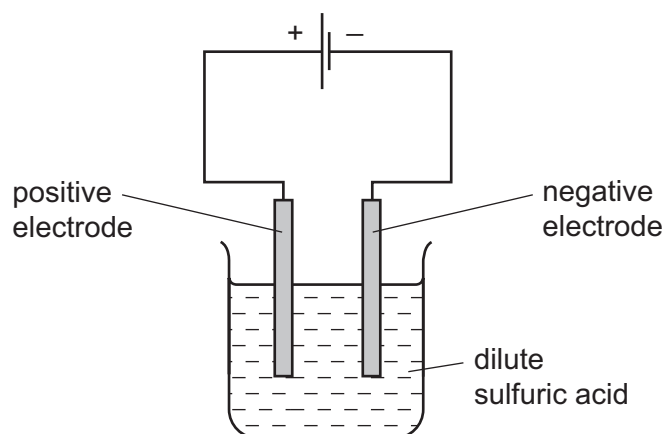
15 The diagram shows the electronic structures of carbon and hydrogen atoms.



What is the formula of the simplest compound formed between carbon and hydrogen?

- A** CH₂ **B** CH₄ **C** C₂H **D** C₄H

16 When dilute sulfuric acid is electrolysed each electrode gives off a different gas.



Which test identifies the gas given off at the positive electrode?

- A** Damp red litmus is bleached.
B Damp red litmus turns blue.
C A glowing splint relights.
D A lighted splint burns with a squeaky pop.

17 Magnesium forms an ionic compound with chlorine.

Which row describes how the magnesium ion is formed and the formula of the magnesium ion?

	formation of the ion	formula of the ion
A	electron gain	Mg ²⁺
B	electron gain	Mg ²⁻
C	electron loss	Mg ²⁺
D	electron loss	Mg ²⁻

18 Marble (calcium carbonate) reacts with dilute hydrochloric acid.

1 g of powdered marble reacts faster with the same volume and concentration of acid than a 1 g lump of marble.

What is the reason for this observation?

- A The powder has a larger mass.
- B The powder has a larger surface area.
- C The powder has a smaller mass.
- D The powder has a smaller surface area.

19 A pupil wants to find out if the reaction of 25cm^3 of an acid with 25cm^3 of an alkali is exothermic.

Which two pieces of apparatus are needed?

- A balance and measuring cylinder
- B Bunsen burner and measuring cylinder
- C Bunsen burner and thermometer
- D thermometer and measuring cylinder

20 Some white anhydrous copper(II) sulfate powder is put into a beaker of water and stirred.

Which observation shows that the process is exothermic?

- A A blue solution forms.
- B A colourless solution forms.
- C The beaker feels cooler.
- D The beaker feels warmer.

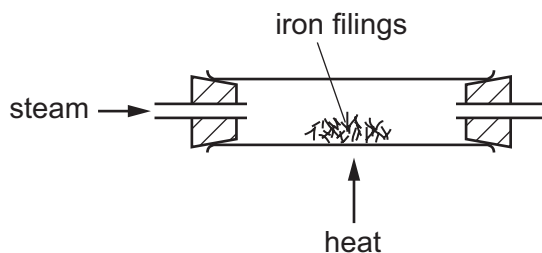
21 Hydrochloric acid is added to calcium carbonate.

Gas X, which turns limewater milky, is given off.

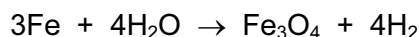
What is X?

- A carbon dioxide
- B chlorine
- C hydrogen
- D oxygen

22 When iron is heated with steam a black solid is formed.



The equation for the reaction is shown:



Which statement is correct for this reaction?

- A Iron has been oxidised because it has gained oxygen.
- B Iron has been reduced because it removed oxygen from water.
- C Iron oxide has been reduced because it contains oxygen.
- D Water has been oxidised because it contains oxygen.

23 Calcium carbonate, CaCO_3 , is decomposed by heating in an industrial process as shown:



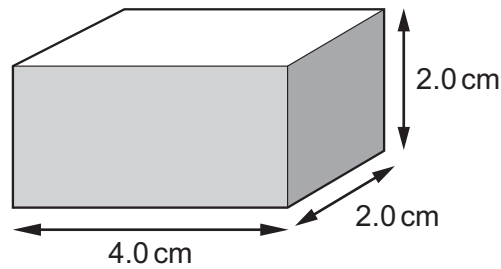
Which statement is **not** correct?

- A The common name for calcium carbonate is limestone.
- B The common name for CaO is lime.
- C CaO is used to neutralise alkaline soil.
- D CaO is used to neutralise industrial waste products.

24 Which row describes an element on the left of the Periodic Table and its oxide?

	type of oxide	type of element
A	acidic	metallic
B	acidic	non-metallic
C	basic	metallic
D	basic	non-metallic

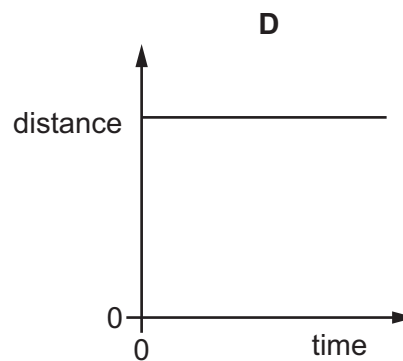
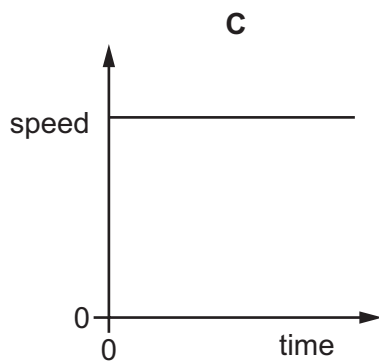
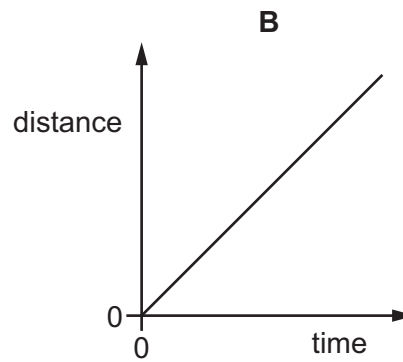
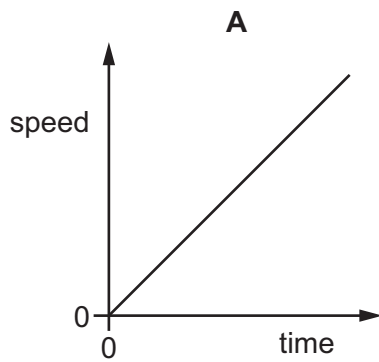
28 The rectangular block shown has a mass of 48 g.



What is the density of the block?

- A** 0.17 g/cm^3 **B** 0.33 g/cm^3 **C** 3.0 g/cm^3 **D** 6.0 g/cm^3

29 Which graph represents the motion of an object that is accelerating?



30 A person wearing wet clothes can feel cold even on a warm day.

Why does he feel cold?

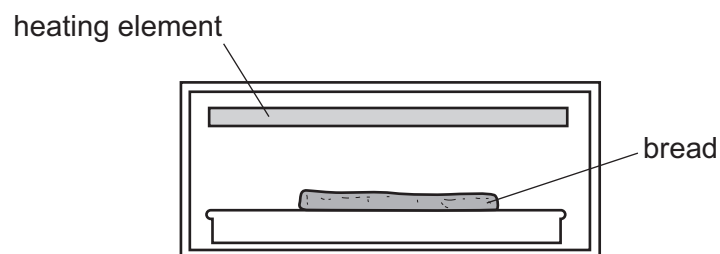
- A** Water gives out heat as it evaporates.
B Water takes in heat as it evaporates.
C Water vapour gives heat out as it condenses.
D Water vapour takes heat in as it condenses.

31 The table lists four energy resources. For each resource it states if the energy resource was originally derived from the Sun's energy.

Which row contains an **error**?

	energy resource	derived from the Sun's energy
A	geothermal	no
B	hydroelectric	no
C	oil	yes
D	waves	yes

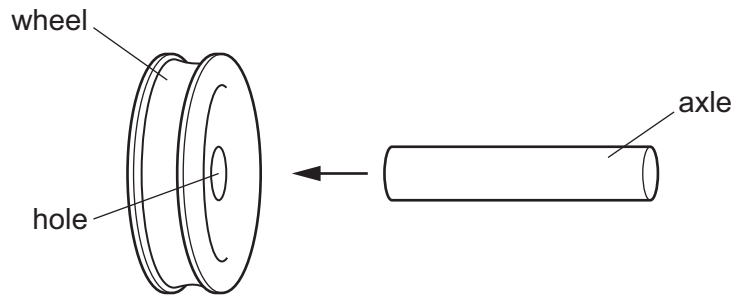
32 Bread can be cooked by placing it below a heating element.



Which process transfers thermal energy from the heating element to the bread?

- A** conduction
- B** convection
- C** evaporation
- D** radiation

- 33 A metal wheel has to be fitted to an axle made from the same metal. The axle is larger than the hole in the wheel.



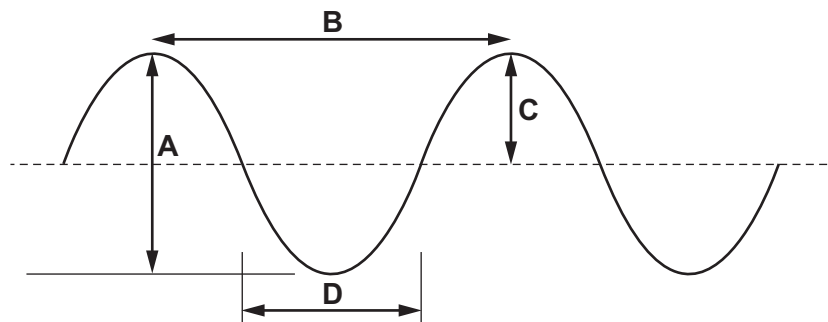
Which action could make it possible to fit the axle in the hole?

- A cooling the axle only
 B cooling the axle and cooling the wheel by the same temperature change
 C heating the axle only
 D heating the axle and heating the wheel by the same temperature change
- 34 A short, loud sound is made in front of a tall building. An echo returns to the source of the sound 0.6 s later.

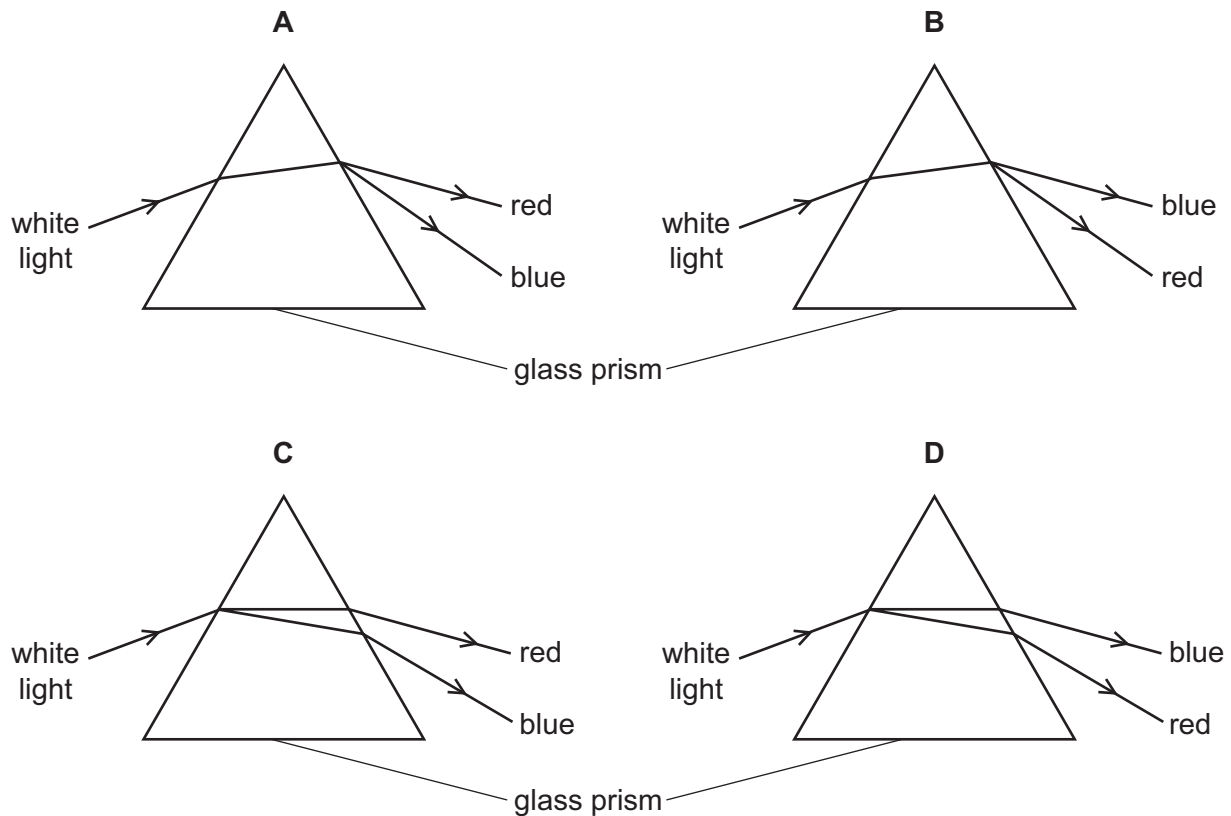
The speed of sound is 330 m/s.

How far away is the building from the source of the sound?

- A 99 m B 198 m C 550 m D 1100 m
- 35 Which distance on the diagram represents the amplitude of the wave?

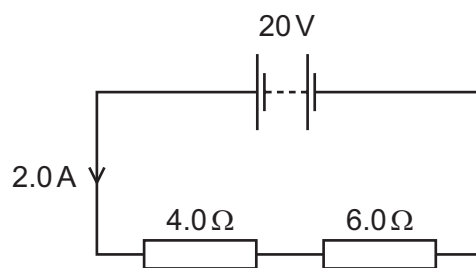


- 36 Which diagram shows the paths taken by the red light and by the blue light when a beam of white light enters a glass prism?



- 37 A 20V battery is connected in series with a 4.0Ω resistor and a 6.0Ω resistor.

The current in the circuit is 2.0A.

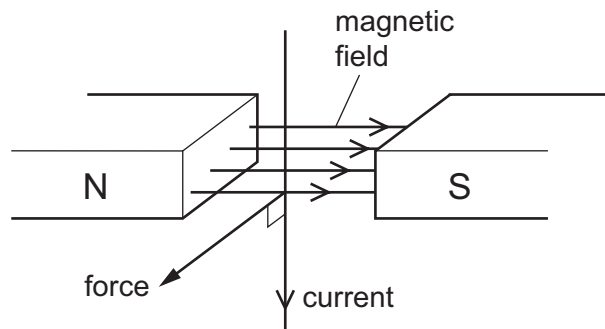


What is the potential difference across the 6.0Ω resistor?

- A 8.0V B 10V C 12V D 20V

- 38 A wire in a magnetic field carries a current. The wire experiences a force due to the magnetic field.

The diagram shows the directions of the magnetic field, the current and the force.



The direction of the current and the direction of the magnetic field are both reversed.

In which direction does the force act now?

- A** in the opposite direction from before the change
B in the same direction as before the change
C towards the north pole
D towards the south pole
- 39 A student believes that a certain steel bar is a magnet.
 What shows that the bar is a magnet?
- A** The bar attracts a copper rod.
B The bar is attracted by one end of another magnet.
C The bar is attracted by both ends of another magnet.
D The bar is repelled by one end of another magnet.
- 40 The table gives the nucleon number and the proton number of three atoms X, Y and Z.

	nucleon number	proton number
X	35	17
Y	37	17
Z	37	18

Which of these atoms are isotopes of the same element?

- A** X and Y only **B** X and Z only **C** Y and Z only **D** X, Y and Z

DATA SHEET
The Periodic Table of the Elements

		Group																																																																						
I	II	III	IV	V	VI	VII	0																																																																	
7 Li Lithium 3	9 Be Beryllium 4	1 H Hydrogen 1	11 B Boron 5	12 C Carbon 6	14 N Nitrogen 7	16 O Oxygen 8	19 F Fluorine 9	20 Ne Neon 10	23 Na Sodium 11	24 Mg Magnesium 12	27 Al Aluminium 13	28 Si Silicon 14	31 P Phosphorus 15	32 S Sulfur 16	35.5 Cl Chlorine 17	40 Ar Argon 18	39 K Potassium 19	40 Ca Calcium 20	45 Sc Scandium 21	48 Ti Titanium 22	51 V Vanadium 23	52 Cr Chromium 24	55 Mn Manganese 25	56 Fe Iron 26	59 Co Cobalt 27	59 Ni Nickel 28	64 Cu Copper 29	65 Zn Zinc 30	70 Ga Gallium 31	73 Ge Germanium 32	75 As Arsenic 33	79 Se Selenium 34	80 Br Bromine 35	84 Kr Krypton 36	85 Rb Rubidium 37	88 Sr Strontium 38	89 Y Yttrium 39	91 Zr Zirconium 40	93 Nb Niobium 41	96 Mo Molybdenum 42	101 Ru Ruthenium 44	103 Rh Rhodium 45	106 Pd Palladium 46	108 Ag Silver 47	112 Cd Cadmium 48	115 In Indium 49	119 Sn Tin 50	122 Sb Antimony 51	128 Te Tellurium 52	127 I Iodine 53	131 Xe Xenon 54	133 Cs Caesium 55	137 Ba Barium 56	139 La Lanthanum 57	178 Hf Hafnium 72	181 Ta Tantalum 73	184 W Tungsten 74	186 Re Rhenium 75	190 Os Osmium 76	192 Ir Iridium 77	195 Pt Platinum 78	197 Au Gold 79	201 Hg Mercury 80	204 Tl Thallium 81	207 Pb Lead 82	209 Bi Bismuth 83	210 Po Polonium 84	210 At Astatine 85	210 Rn Radon 86	226 Ra Radium 88	227 Ac Actinium 89	†
												140 Ce Cerium 58	141 Pr Praseodymium 59	144 Nd Neodymium 60	146 Pm Promethium 61	150 Sm Samarium 62	152 Eu Europium 63	157 Gd Gadolinium 64	159 Tb Terbium 65	162 Dy Dysprosium 66	165 Ho Holmium 67	167 Er Erbium 68	169 Tm Thulium 69	173 Yb Ytterbium 70	175 Lu Lutetium 71	232 Th Thorium 90	238 Pa Protactinium 91	238 U Uranium 92	238 Np Neptunium 93	238 Pu Plutonium 94	238 Am Americium 95	238 Cm Curium 96	238 Bk Berkelium 97	238 Cf Californium 98	238 Es Einsteinium 99	238 Fm Fermium 100	238 Md Mendelevium 101	238 No Nobelium 102	238 Lr Lawrencium 103																																	

*58-71 Lanthanoid series
†90-103 Actinoid series

a	X	a = relative atomic mass
b	X	b = proton (atomic) number

Key

The volume of one mole of any gas is 24 dm³ at room temperature and pressure (r.t.p.).

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