



Cambridge International Examinations
Cambridge International General Certificate of Secondary Education

CANDIDATE
NAME

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CO-ORDINATED SCIENCES

0654/33

Paper 3 (Extended)

May/June 2014

2 hours

Candidates answer on the Question Paper.

No Additional Materials are required.

READ THESE INSTRUCTIONS FIRST

Write your Centre number, candidate number and name on all the work you hand in.

Write in dark blue or black pen.

You may use an HB pencil for any diagrams or graphs.

Do not use staples, paper clips, glue or correction fluid.

DO NOT WRITE IN ANY BARCODES.

Answer **all** questions.

Electronic calculators may be used.

You may lose marks if you do not show your working or if you do not use appropriate units.

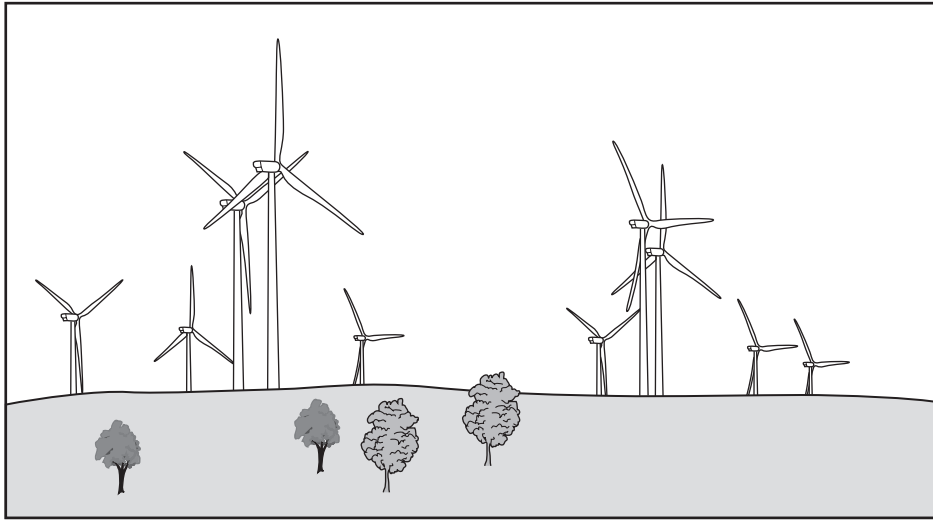
A copy of the Periodic Table is printed on page 32.

At the end of the examination, fasten all your work securely together.

The number of marks is given in brackets [] at the end of each question or part question.

This document consists of **32** printed pages.

- 1 (a) Wind farms are areas of land containing many wind turbines. Four thousand wind turbines can produce the same power as one coal-fired power station.



- (i) State **one** advantage and **one** disadvantage of using wind, rather than coal, to generate electrical power.

advantage

.....

disadvantage

.....

[1]

- (ii) On a particular day, the power input to a wind turbine is 1500 kW. The turbine produces 900 kW of electrical power.

Calculate the efficiency of the wind turbine.

State any formula that you use and show your working. State your answer as a percentage.

formula

working

..... % [2]

(b) Nuclear power stations generate electricity using energy released by the nuclear fission of atoms.

(i) Describe the process that transforms this energy into electrical energy.

.....
.....
.....
..... [3]

(ii) Energy is released in the Sun by a different nuclear process.

Name this process.

..... [1]

(c) A wind farm generates 33MW of electrical power. The wind farm is connected to a transmission line at a potential difference of 132kV.

Calculate the current produced by the wind farm.

State the formula that you use and show your working.

formula

working

..... A [2]

- (d) Fig. 1.1 shows how the electricity cables carrying electricity from a wind farm are attached to pylons.

The cables hang loosely in hot weather.

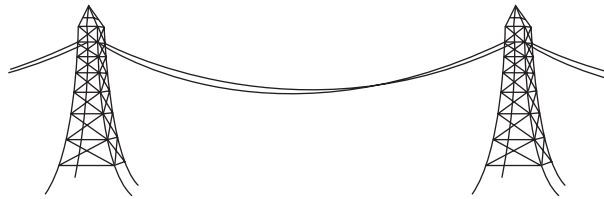


Fig. 1.1

Explain why the cables must hang loosely in hot weather.

.....

.....

..... [2]

- (e) A scientist investigates six different wires used in making these cables. He wants to determine the resistance of each piece of wire.

wire	metal composition	length / m	cross-sectional area / cm ²
A	copper	10	0.1
B	nichrome	10	0.1
C	copper	20	0.1
D	nichrome	20	0.1
E	copper	10	0.2
F	nichrome	20	0.2

- (i) Which wire, **A** or **E**, will have the greater resistance?

Explain your answer.

wire because

..... [1]

(ii) Wire **B** has a greater resistance than wire **A**.

Which wire, **B**, **C**, **D**, **E** or **F**, has the greatest resistance?

Explain your answer.

wire

explanation

..... [2]

(iii) The resistance of wire **B** is 0.15Ω .

Calculate the current passing through the wire when a voltage of 12V is applied across it.

State the formula that you use and show your working.

formula

working

..... A [2]

2 (a) Fig. 2.1 shows some of the cells that line the trachea.

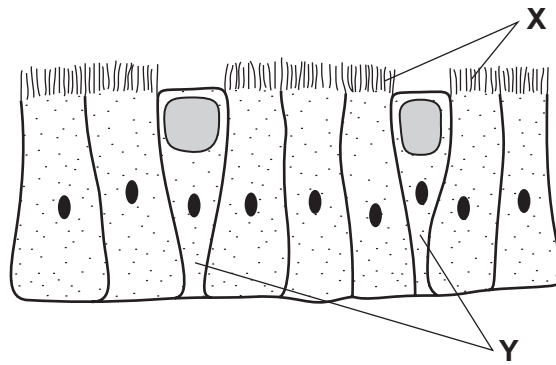


Fig. 2.1

(i) Name the structures labelled X.

..... [1]

(ii) Explain how these structures, and the cells labelled Y, protect the gas exchange system from pathogens.

.....

 [3]

(b) Tobacco smoke can have a damaging effect on the working of the cells in Fig. 2.1.

(i) Name a component of tobacco smoke that damages these cells.

..... [1]

(ii) Describe how this component of tobacco smoke affects the structures labelled X and the cells labelled Y.

structures labelled X

 cells labelled Y

[2]

Please turn over for Question 3.

- 3 (a) Dutch metal is an alloy of copper and zinc that has been formed into very thin sheets.

When a small piece of Dutch metal is dropped into a container filled with chlorine, it bursts into flame and two compounds are produced as shown in Fig. 3.1.

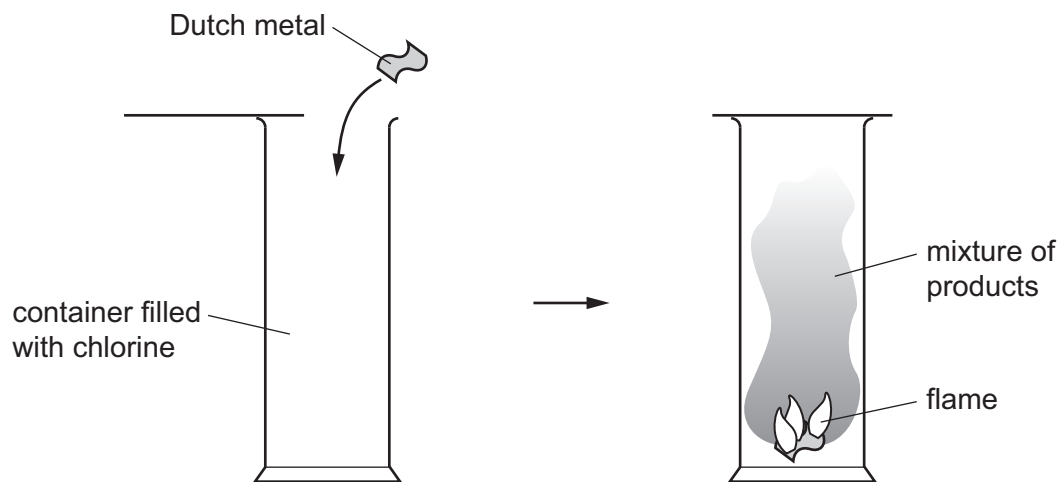


Fig. 3.1

- (i) State the meaning of the term *alloy*.

.....
 [1]

- (ii) State the physical property of metals that allows them to be formed into very thin sheets.

..... [1]

- (iii) Suggest the names of the **two** compounds formed when Dutch metal reacts with chlorine.

1

2 [1]

- (b) Sodium burns in oxygen gas to produce a white solid that contains the ionic compound, sodium oxide.

Fig. 3.2 shows a sodium atom and an oxygen atom.

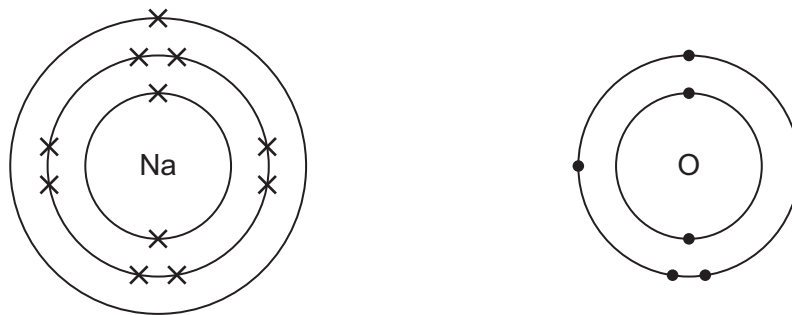


Fig. 3.2

Predict and explain, in terms of changes in electronic structure, the chemical formula of sodium oxide. You may wish to draw diagrams to help you to answer this question.

.....
 [3]

- (c) Phosphorus is a non-metallic element containing molecules that have the formula P_4 .

The chemical formula of phosphorus oxide shows four phosphorus atoms bonded with ten oxygen atoms.

Construct a balanced symbolic equation for the reaction between phosphorus and oxygen gas to form phosphorus oxide.

..... [3]

- 4 Fig. 4.1 shows a river with nearby agricultural land. Large amounts of artificial fertiliser have been sprayed onto the agricultural land.

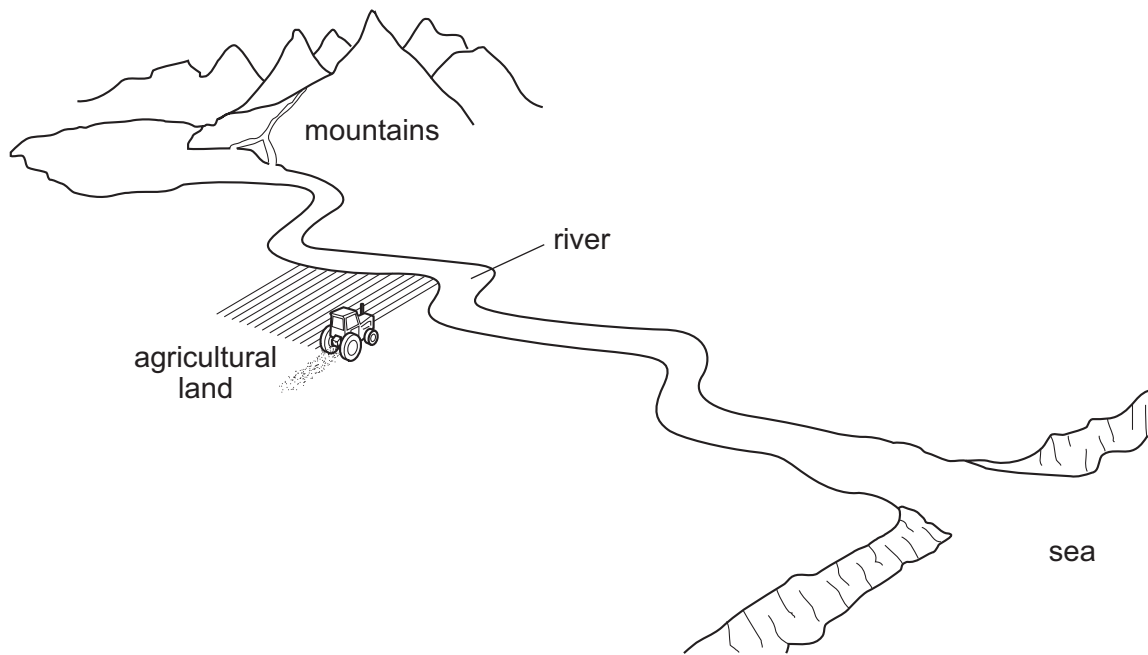


Fig. 4.1

- (a) Name a mineral ion that would be present in the fertiliser.

..... [1]

- (b) Describe how mineral ions in the fertiliser might reach the river.

.....
 [1]

- (c) When large amounts of mineral ions are added to a river a sequence of effects on the living organisms can take place.

Explain the effects on the following organisms

- (i) algae (photosynthesising microorganisms),

.....
 [1]

- (ii) submerged aquatic plants,

.....

 [2]

(iii) bacteria,

.....
.....
..... [2]

(iv) fish.

.....
..... [1]

(d) If the farmer uses artificial fertiliser, suggest **two** ways in which the effect of the fertiliser on the river could be reduced.

1
.....
2
..... [2]

5 (a) Two bar magnets **A** and **B** are shown in Fig. 5.1. Magnet **A** is moved towards magnet **B**.

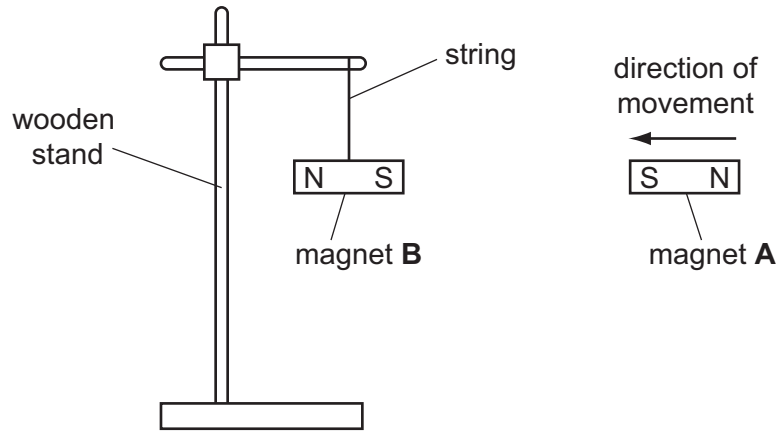


Fig. 5.1

(i) Describe and explain what happens to magnet **B** as magnet **A** is moved towards it.

.....
 [1]

(ii) Magnet **A** is replaced by a piece of unmagnetised iron **C**.

Predict what happens as the unmagnetised iron **C** is moved towards **B**.

Explain your prediction.

.....

 [2]

(b) Fig. 5.2 shows two plastic balls hanging from threads. Both balls are electrically charged.

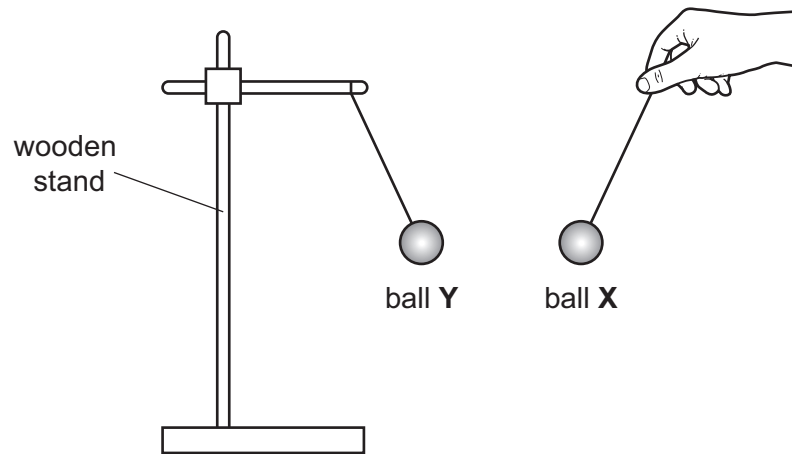


Fig. 5.2

Ball Y is negatively charged.

(i) State the charge on ball X. Give a reason for your answer.

.....
 [1]

(ii) Describe and explain how ball Y has been given a negative charge.

.....
 [2]

(iii) There is an electric field between ball X and ball Y.

State what happens to an electrical charge placed in this field.

.....
 [1]

14

(c) The mass of ball **X** is 3.97 g (3.97×10^{-3} kg). The volume of ball **X** is 4.17 cm³ (4.17×10^{-6} m³).

Calculate the density of the plastic used to make ball **X**.

State the formula that you use and show your working. State the units of your answer.

formula

working

density = unit = [3]

Please turn over for Question 6.

- 6 (a) Fig. 6.1 shows diagrams **P**, **Q** and **R**, of three molecules containing carbon atoms.

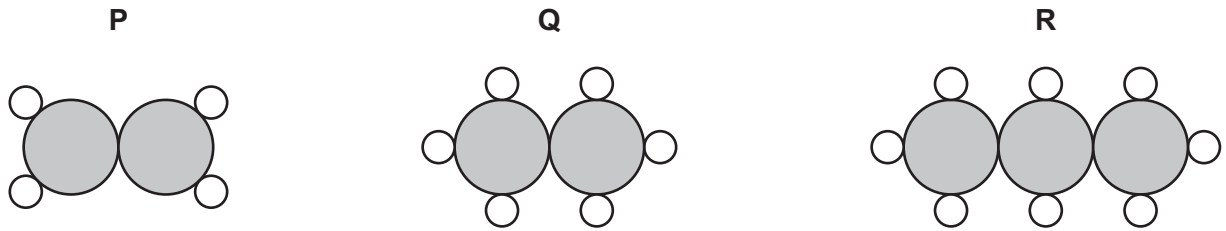


Fig. 6.1

- (i) Using the Periodic Table on page 32, state the number of electrons in one atom of carbon.

Explain how you obtained your answer.

number of electrons

explanation

..... [2]

- (ii) State and explain which diagram, **P**, **Q** or **R**, represents one molecule of ethane.

diagram

explanation

.....

..... [2]

- (iii) Name the type of chemical bonding found in all of the compounds shown in Fig. 6.1.

Give a reason for your answer.

type of bonding

reason

..... [2]

- (b) Methane hydrate is a solid mixture in which methane molecules are contained inside ice crystals.

Large amounts of methane hydrate exist under the oceans and in the cold polar regions of the Earth.

Table 6.1 shows the relative numbers of moles of methane and water in a typical sample of methane hydrate.

Table 6.1

substance	chemical formula	relative number of moles
methane	CH ₄	1.00
water (ice)	H ₂ O	5.75

- (i) The mass of 1.00 moles of methane is 16.0g.

Calculate the mass of 5.75 moles of water.

Show your working.

..... [2]

- (ii) Calculate the mass of methane hydrate that contains 1.00 moles of methane.

..... [1]

- (iii) When the temperature of methane hydrate increases, the ice melts and releases the methane.

Some scientists think that methane hydrate might have a serious effect on global warming.

Suggest how the breakdown of methane hydrate might affect global warming.

.....

 [2]

7 An electric motor inflates a car tyre by pumping air into it.

(a) Explain, in terms of particles, how the air causes the tyre to inflate.

.....

.....

.....

..... [3]

(b) Fig. 7.1 shows a simple electric motor.

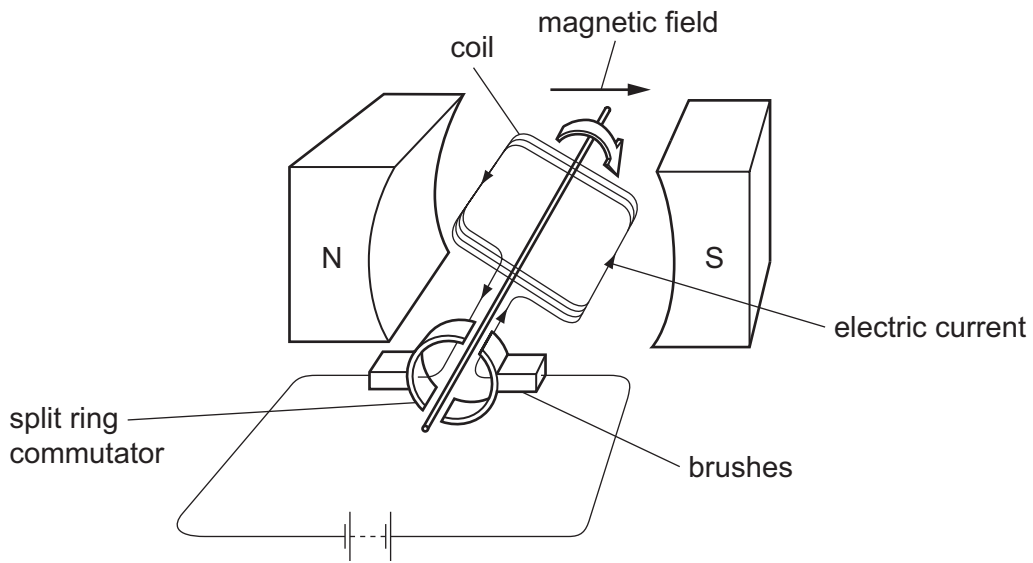


Fig. 7.1

Explain why the coil turns when an electric current passes through it.

.....

.....

.....

.....

..... [4]

Please turn over for Question 8.

- 8 After its flowers have been pollinated, a sweetcorn (maize) plant produces a corncob as shown in Fig. 8.1.

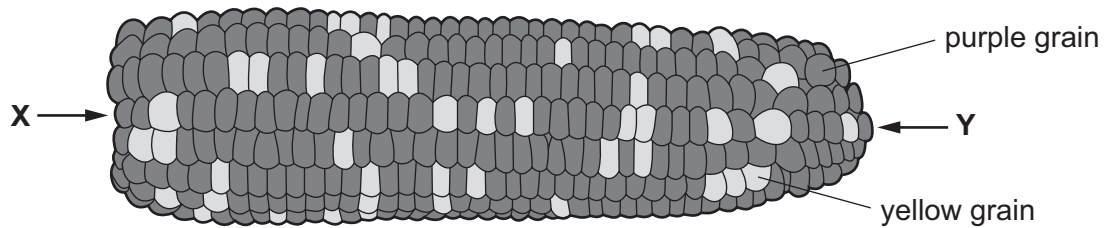


Fig. 8.1

Each of the individual grains on the corncob results from the fertilisation of a different egg cell in the female parent. The pollen all came from the same (male) parent.

Some of the grains are purple (dark) in colour and others yellow (light) in colour.

- (a) The variation in grain colour is an example of discontinuous variation.

Explain why this variation is described as *discontinuous*.

.....

.....

..... [2]

- (b) (i) In the row of grains labelled X to Y, count the number of purple (dark) grains and the number of yellow (light) grains.

number of purple (dark) grains

number of yellow (light) grains [1]

- (ii) State, to the nearest whole number, the ratio of purple grains to yellow grains.

..... [1]

- (c) The allele for purple colour (**G**) is dominant and the allele for yellow colour (**g**) is recessive.

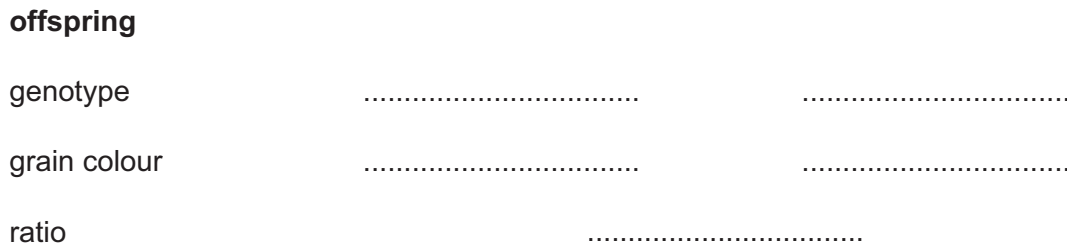
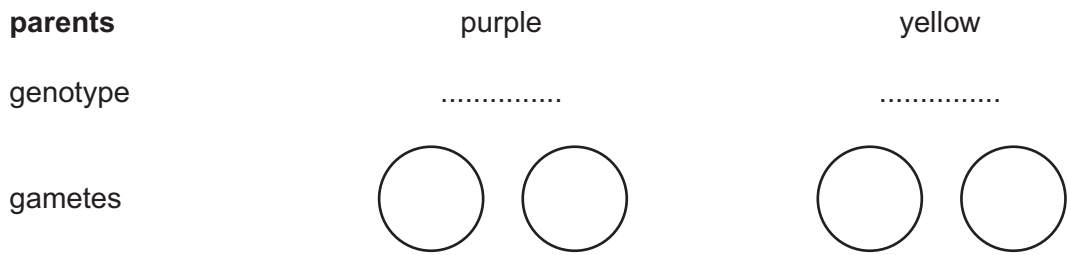
- (i) What would be the colour of a sweetcorn grain with the genotype **Gg**?

..... [1]

- (ii) Use the ratio of purple grains and yellow grains in (b)(ii) to state the genotypes of the parents.

genotypes and [2]

(d) Complete the genetic diagram below to show the result of crossing a heterozygous sweetcorn plant with a yellow-grained sweetcorn plant.



[5]

- 9 (a) Fig. 9.1 shows air passing into the engine of a car, and a mixture of exhaust (waste) gases being released.

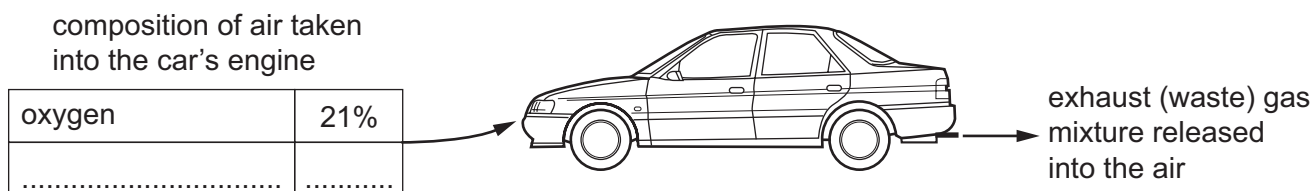


Fig. 9.1

- (i) Complete the table in Fig. 9.1 to show the name and percentage of the main gas in air. [2]

- (ii) Name **one** gas, other than carbon dioxide, in the mixture of exhaust gases which causes air pollution.

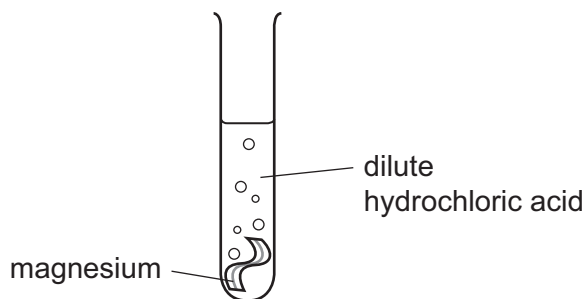
State **one** harmful effect that this gas has in the environment.

gas

harmful effect

..... [2]

- (b) Hydrogen gas is released when magnesium reacts with dilute hydrochloric acid.



- (i) Describe the test for hydrogen gas.

..... [2]

- (ii) State the **word** equation for the reaction between magnesium and dilute hydrochloric acid.

..... [1]

- (c) Fig. 9.2 shows the apparatus a student used to measure the temperature change when magnesium powder reacted with dilute hydrochloric acid.

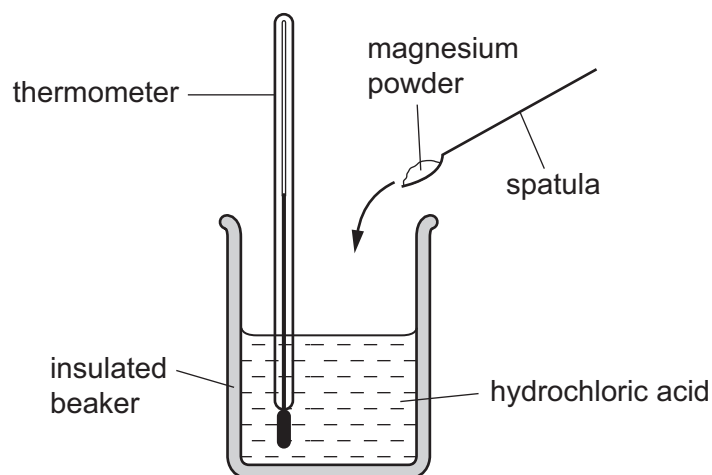


Fig. 9.2

The student repeated the experiment using different masses of magnesium powder.

After each experiment he rinsed out the insulated beaker and then refilled it using the same volume of 1.0 mol/dm^3 hydrochloric acid. His results are shown in Fig. 9.3.

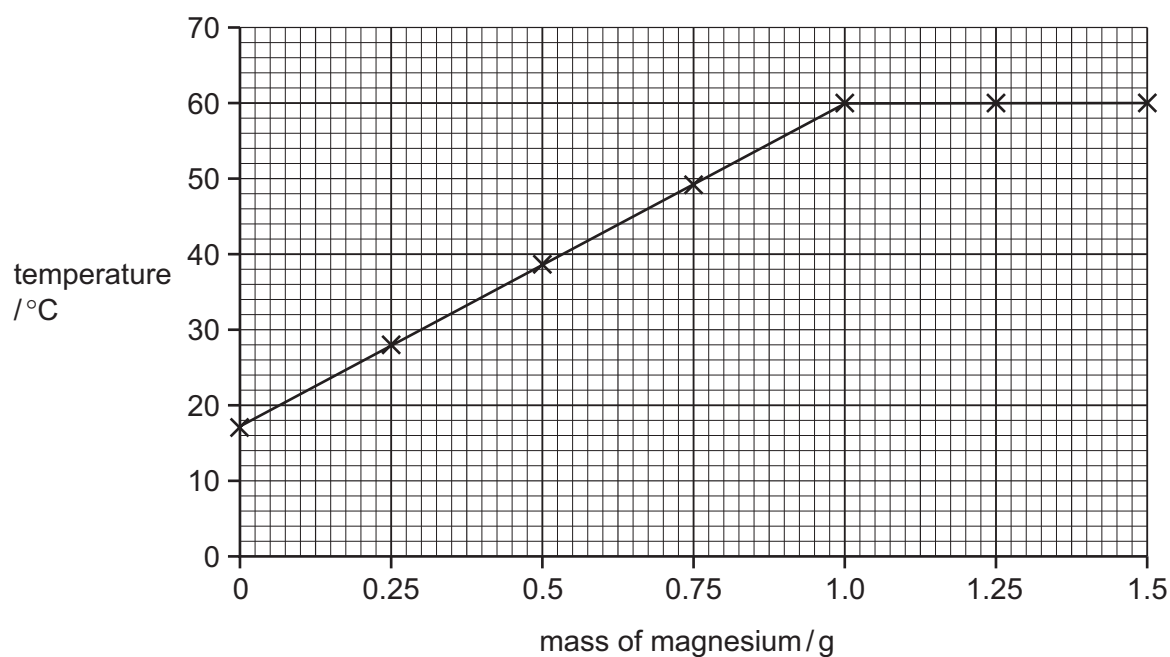


Fig. 9.3

- (i) Explain, in terms of energy, why the temperature of the reaction mixture increases when magnesium powder is added to dilute hydrochloric acid.

.....

.....

.....

..... [2]

(ii) Suggest why in this experiment the graph eventually became horizontal.

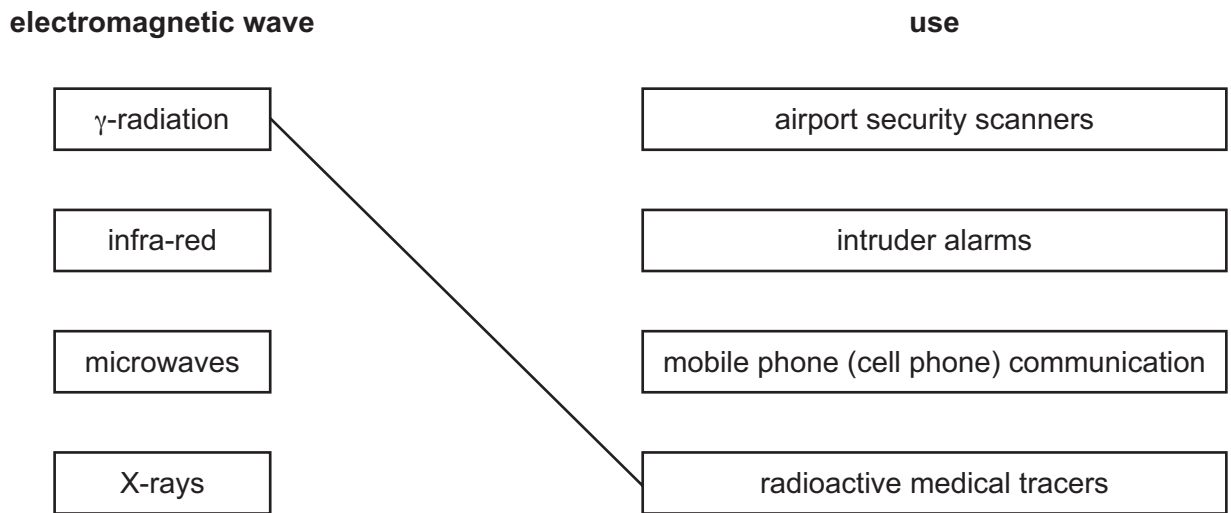
.....

.....

..... [2]

Please turn over for Question 10.

10 (a) Draw lines to link the waves in the electromagnetic spectrum to their uses. One line has been drawn for you.



[1]

(b) Different waves in the electromagnetic spectrum have different wavelengths and frequencies.

State the meaning of the terms *frequency* and *wavelength*.

You may use diagrams to help your explanation.

frequency

.....

.....

.....

wavelength

.....

.....

.....

[2]

(c) α -radiation, β -radiation and γ -radiation are three radioactive emissions.

(i) Place the three radiations in order of their ionising ability, placing the most ionising first.

most ionising

.....

least ionising

[1]

(ii) Fig. 10.1 shows α , β , and γ radiations passing through a magnetic field.

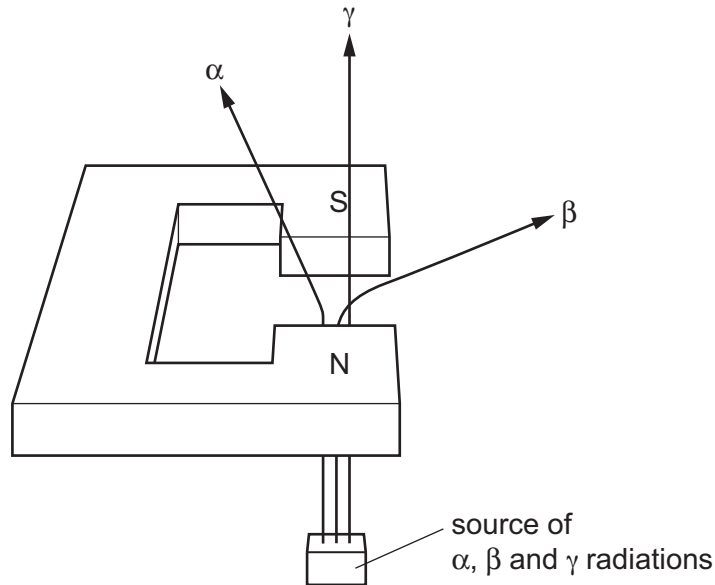


Fig. 10.1

Explain the results.

.....

 [3]

11 (a) Define *osmosis*.

.....

.....

.....

..... [3]

(b) A piece of plant tissue was placed in a concentrated sugar solution on a microscope slide. Fig. 11.1 shows the appearance of three of the cells from this tissue after they had been in the sugar solution for one hour.

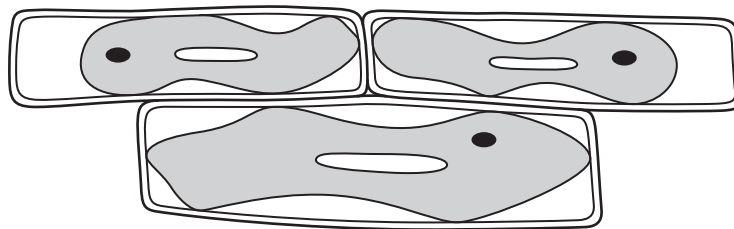


Fig. 11.1

(i) Describe the effect, as shown in Fig. 11.1, that the sugar solution has had on the cells.

.....

..... [1]

(ii) Explain this effect in terms of osmosis.

.....

.....

..... [2]

(iii) Complete Fig. 11.2, to show how the cells would appear if they had been placed in water, instead of in a concentrated sugar solution.

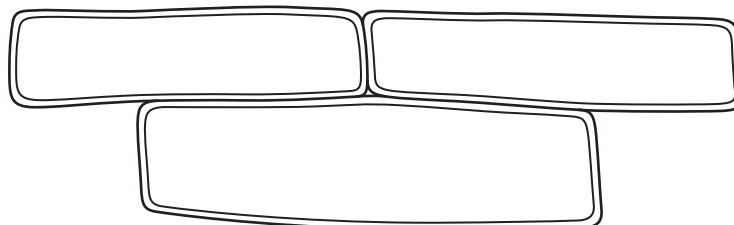


Fig. 11.2

[2]

(c) Plants absorb water by osmosis into their root hair cells.

(i) Explain how the structure of the root hair cells is related to this function.

.....
.....
..... [2]

(ii) State **one** other function of root hair cells.

..... [1]

12 (a) Fig. 12.1 shows some of the particles present in a mixture of gases.

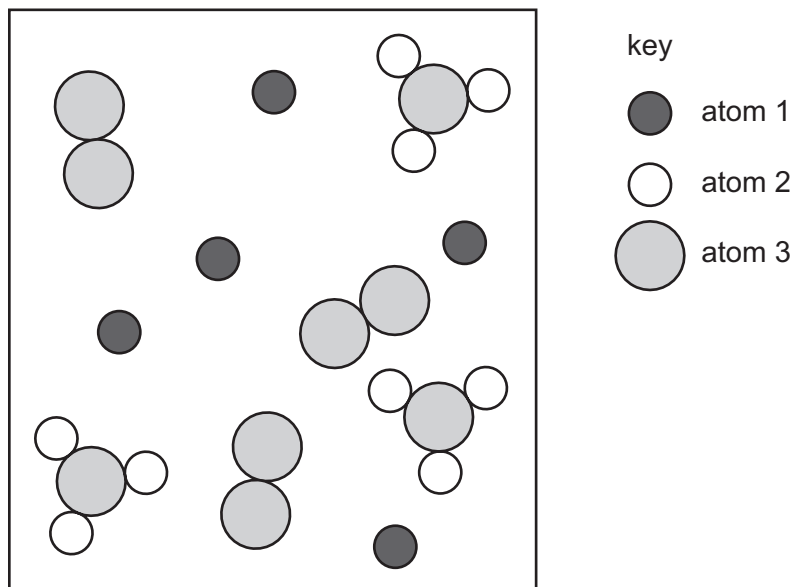


Fig. 12.1

(i) State the number of different gases that are contained in the mixture shown in Fig. 12.1.

..... [1]

(ii) On Fig. 12.1 draw a label line to a molecule of a **compound**. Label this molecule **C**. [1]

(iii) Explain your answer to (ii).

.....
 [1]

(b) Name the family of metals that includes cobalt (proton number 27) and nickel (proton number 28).

..... [1]

(c) Fig. 12.2 shows a simplified diagram of the industrial process used to produce aluminium.

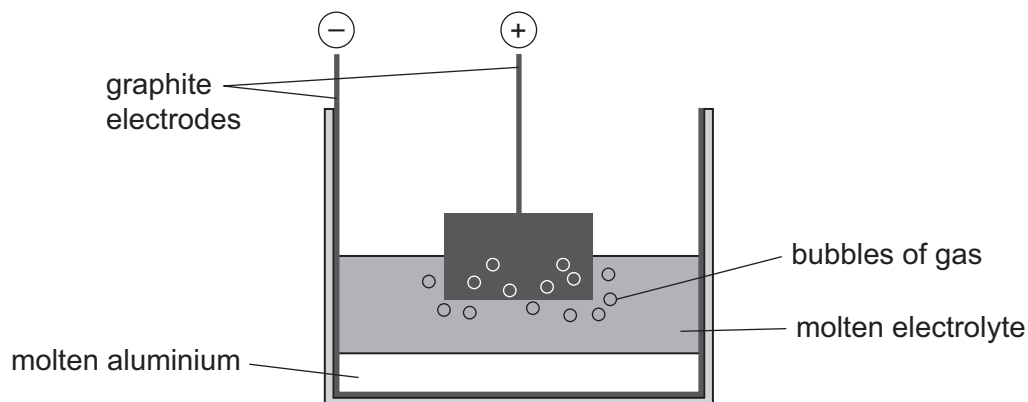


Fig. 12.2

(i) Name the **two** substances that are melted together to form the electrolyte.

1

2

[2]

(ii) Name **one** gas that bubbles from the surface of the anode.

..... [1]

(iii) Describe what happens on the surface of the cathode to convert aluminium ions, Al^{3+} , to aluminium atoms.

.....

.....

..... [2]

DATA SHEET
The Periodic Table of the Elements

		Group											
I	II	III	IV	V	VI	VII	0					0	
7 Li Lithium 3	9 Be Beryllium 4	1 H Hydrogen 1	11 B Boron 5	12 C Carbon 6	13 Al Aluminium 13	14 Si Silicon 14	15 P Phosphorus 15	16 S Sulfur 16	17 Cl Chlorine 17	18 Ar Argon 18	19 F Fluorine 9	20 Ne Neon 10	
23 Na Sodium 11	24 Mg Magnesium 12		27 Al Aluminium 13	28 Si Silicon 14	29 Co Cobalt 27	30 Zn Zinc 30	31 Ga Gallium 31	32 Ge Germanium 32	33 As Arsenic 33	34 Se Selenium 34	35 Br Bromine 35	36 Kr Krypton 36	
39 K Potassium 19	40 Ca Calcium 20		51 V Vanadium 23	52 Cr Chromium 24	55 Mn Manganese 25	56 Fe Iron 26	59 Ni Nickel 28	64 Cu Copper 29	70 Ga Gallium 31	73 Ge Germanium 32	75 As Arsenic 33	79 Se Selenium 34	84 Kr Krypton 36
85 Rb Rubidium 37	88 Sr Strontium 38		91 Ti Titanium 22	96 Mo Molybdenum 42	101 Ru Ruthenium 44	106 Pd Palladium 46	108 Ag Silver 47	112 Cd Cadmium 48	115 In Indium 49	119 Sn Tin 50	122 Sb Antimony 51	128 Te Tellurium 52	131 Xe Xenon 54
133 Cs Caesium 55	137 Ba Barium 56		48 Ti Titanium 22	74 W Tungsten 74	75 Re Rhenium 75	78 Pt Platinum 78	79 Au Gold 79	80 Hg Mercury 80	204 Tl Thallium 81	207 Pb Lead 82	209 Bi Bismuth 83	84 Po Polonium 84	86 Rn Radon 86
87 Fr Francium	226 Ra Radium		45 Sc Scandium 21	73 Ta Tantalum 73	76 Os Osmium 76	81 Ir Iridium 77	82 Pt Platinum 78	83 Au Gold 79	84 Hg Mercury 80	85 Tl Thallium 81	86 Pb Lead 82	87 Bi Bismuth 83	88 Rn Radon 86
			93 Nb Niobium 41	94 Zr Zirconium 40	103 Rh Rhodium 45	104 Pd Palladium 46	105 Ag Silver 47	106 Cd Cadmium 48	107 In Indium 49	108 Sn Tin 50	109 Sb Antimony 51	110 Te Tellurium 52	111 Xe Xenon 54
			140 Ce Cerium 58	141 Pr Praseodymium 59	144 Nd Neodymium 60	150 Sm Samarium 62	152 Eu Europium 63	157 Gd Gadolinium 64	162 Dy Dysprosium 66	165 Ho Holmium 67	167 Er Erbium 68	169 Tm Thulium 69	175 Lu Lutetium 71
			232 Th Thorium 90	91 Pa Protactinium 91	92 U Uranium 92	94 Pu Plutonium 94	95 Am Americium 95	96 Cm Curium 96	98 Cf Californium 98	99 Es Einsteinium 99	100 Fm Fermium 100	101 Md Mendelevium 101	103 Lr Lawrencium 103

*58-71 Lanthanoid series
†90-103 Actinoid series

a	X	a = relative atomic mass
b	X	X = atomic symbol
		b = proton (atomic) number

The volume of one mole of any gas is 24 dm³ at room temperature and pressure (r.t.p.).

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