



Cambridge IGCSE™ (9–1)

CO-ORDINATED SCIENCES**0973/11**

Paper 1 Multiple Choice (Core)

October/November 2021**45 minutes**

You must answer on the multiple choice answer sheet.

You will need: Multiple choice answer sheet
Soft clean eraser
Soft pencil (type B or HB is recommended)

INSTRUCTIONS

- There are **forty** questions on this paper. Answer **all** questions.
- For each question there are four possible answers **A, B, C** and **D**. Choose the **one** you consider correct and record your choice in soft pencil on the multiple choice answer sheet.
- Follow the instructions on the multiple choice answer sheet.
- Write in soft pencil.
- Write your name, centre number and candidate number on the multiple choice answer sheet in the spaces provided unless this has been done for you.
- Do **not** use correction fluid.
- Do **not** write on any bar codes.
- You may use a calculator.

INFORMATION

- The total mark for this paper is 40.
- Each correct answer will score one mark.
- Any rough working should be done on this question paper.
- The Periodic Table is printed in the question paper.

This document has **16** pages.



- 1 All living organisms can break down nutrient molecules to release energy.

What is this process?

- A excretion
 - B growth
 - C nutrition
 - D respiration
- 2 Which structure is only found in plant cells?
- A cell membrane
 - B cytoplasm
 - C nucleus
 - D vacuole
- 3 Which row shows a large molecule and a basic unit from which it is made?

	large molecule	basic unit
A	glycogen	amino acid
B	glycogen	glucose
C	oil	amino acid
D	oil	glucose

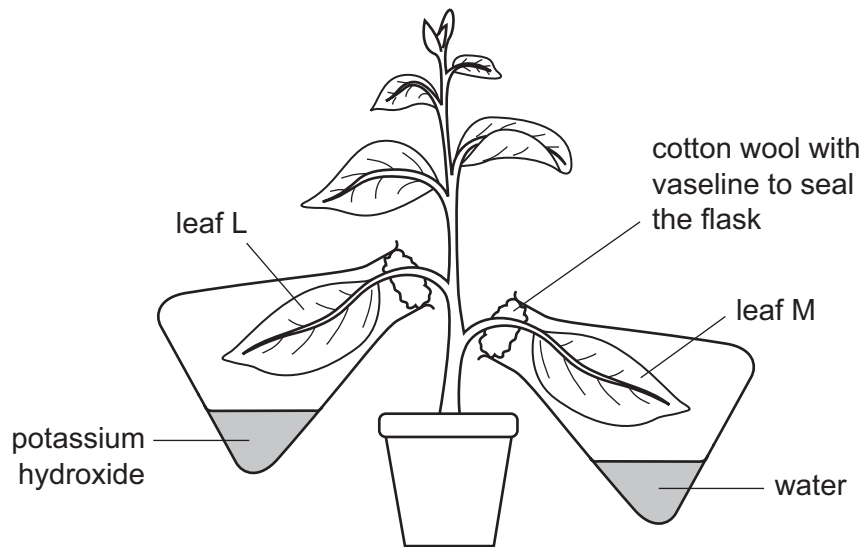
- 4 A student investigated the effect of pH on an enzyme that digests starch.

Which chemical will be needed to determine if any starch has been digested?

- A Benedict's solution
- B biuret solution
- C iodine solution
- D ethanol

- 5 The diagram shows an experiment to investigate photosynthesis. When leaves photosynthesise, they store some carbohydrates as starch.

Potassium hydroxide absorbs carbon dioxide.

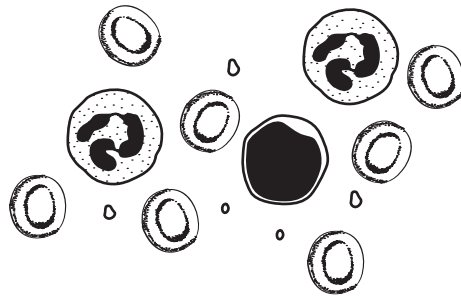


After standing in sunlight for 10 hours, leaf L contained no starch but leaf M contained a lot of starch.

What does this show?

- A A leaf cannot make starch in a sealed flask.
 - B A leaf cannot make starch without carbon dioxide.
 - C A leaf cannot make starch without light.
 - D A leaf cannot make starch without oxygen.
- 6 Which nutrient is well provided by citrus fruits such as oranges and lemons?
- A carbohydrate
 - B protein
 - C vitamin C
 - D vitamin D

7 The diagram shows some blood viewed under a light microscope.



How many red blood cells are shown?

- A** 1 **B** 3 **C** 5 **D** 7

8 What is the word equation for aerobic respiration?

- A** carbon dioxide + water → glucose + oxygen
B carbon dioxide + oxygen → glucose + water
C glucose + oxygen → carbon dioxide + water
D glucose + water → carbon dioxide + oxygen

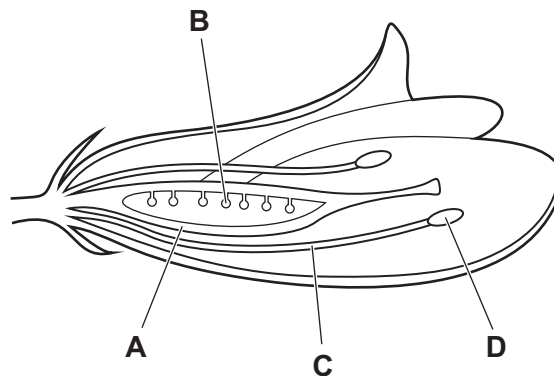
9 A person touches a hot object with their hand. They quickly pull their hand away.

Which statement is correct?

- A** The effector is their hand.
B The effector is the hot object.
C The receptor is in the muscles of their arm.
D The receptor is in the skin of their hand.

10 The diagram shows a section of a pea flower.

Which part is the ovary?



11 Natural selection involves several stages.

- 1 passing on alleles to the next generation
- 2 struggle for survival
- 3 competition for resources
- 4 production of many offspring

What is the correct order of these stages?

- A** 4 → 1 → 2 → 3
B 4 → 3 → 2 → 1
C 1 → 2 → 3 → 4
D 3 → 2 → 1 → 4

12 The diagram shows a food chain.

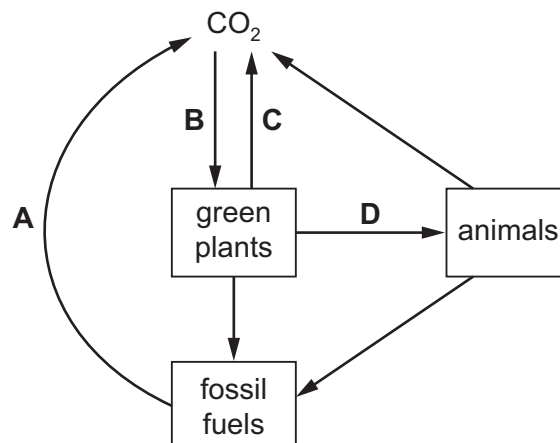
grass → gazelle → lion

Which position does the lion occupy in the food chain?

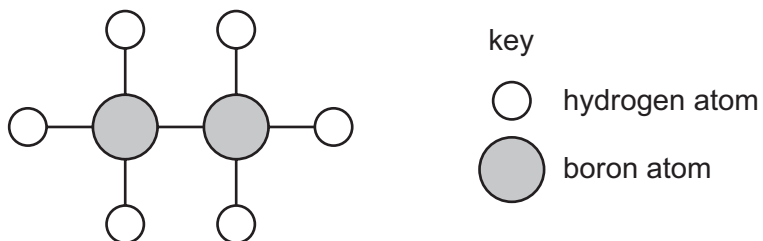
- A** primary consumer
B producer
C secondary consumer
D tertiary consumer

13 The diagram shows a simplified carbon cycle.

Which labelled arrow represents respiration?



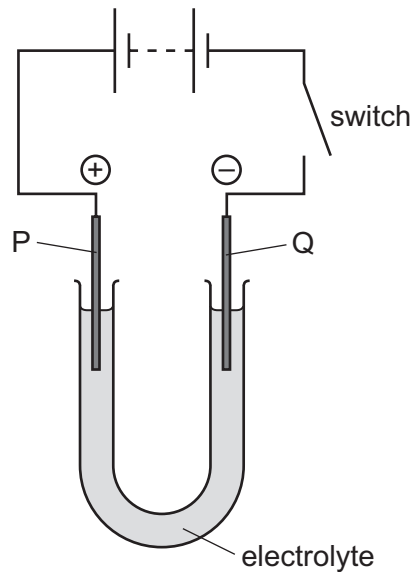
16 A model of a molecule is shown.



Which row shows the formula of this molecule and describes the type of bonding between the atoms?

	formula	bonding
A	2BH_3	covalent
B	2BH_3	ionic
C	B_2H_6	covalent
D	B_2H_6	ionic

17 The diagram shows the electrolysis of a compound.



When the switch is closed, the solution around electrode P turns orange because a halogen is formed.

The positive electrode P is called the1....., and the halogen is2..... .

Which words complete gaps 1 and 2?

	1	2
A	anode	bromine
B	anode	chlorine
C	cathode	bromine
D	cathode	chlorine

18 The initial and final temperatures of four different experiments are measured.

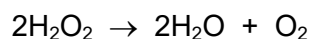
The results are shown.

Which experiment is the most endothermic?

	initial temperature /°C	final temperature /°C
A	22	17
B	21	25
C	20	27
D	20	18

- 19 Aqueous hydrogen peroxide decomposes slowly and produces water and oxygen gas.

The equation for this decomposition is shown.



The time taken to produce the first 20 cm³ of gas and the total volume of gas produced in the reaction are measured.

The experiment is repeated using a catalyst.

Which row describes the results for the second experiment?

	time taken to produce first 20 cm ³ of gas	total volume of gas produced
A	less than experiment 1	the same as experiment 1
B	less than experiment 1	more than experiment 1
C	more than experiment 1	the same as experiment 1
D	more than experiment 1	more than experiment 1

- 20 The colour of universal indicator in solutions S, T and U is shown.

solution	S	T	U
colour of universal indicator	orange	green	purple

Which row shows the pH values of the solutions?

	S	T	U
A	1	5	9
B	1	7	14
C	4	5	9
D	4	7	14

- 21 When a small piece of potassium is placed in water, hydrogen gas is given off very quickly.

Which element reacts in a similar way?

- A** copper
- B** iron
- C** magnesium
- D** sodium

22 Why does the steel used to make a drill contain manganese?

- A to increase the density of the steel
- B to increase the hardness of the steel
- C to increase the malleability of the steel
- D to increase the melting point of the steel

23 What is the colour of cobalt(II) chloride after water is added to it?

- A blue
- B pink
- C white
- D green

24 Which process does **not** produce carbon dioxide?

- A complete combustion of fossil fuels
- B reaction of an acid with a carbonate
- C respiration in plants
- D rusting iron

25 Calcium carbonate (limestone) is a base.

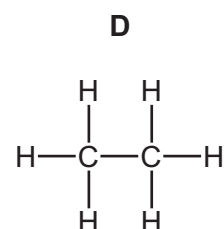
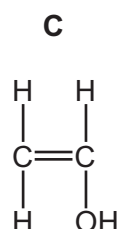
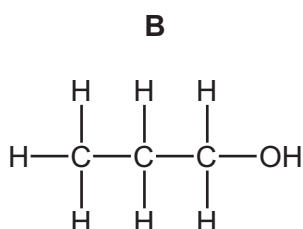
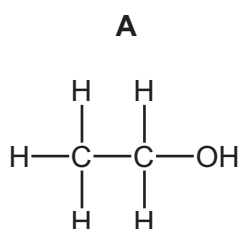
Which uses of limestone depend on it acting as a base?

- 1 making lime
- 2 neutralising acid waste
- 3 stone buildings
- 4 treatment of soil

- A** 1 and 2 **B** 2 and 3 **C** 2 and 4 **D** 3 and 4

26 Four molecules are shown.

Which structure represents ethanol?



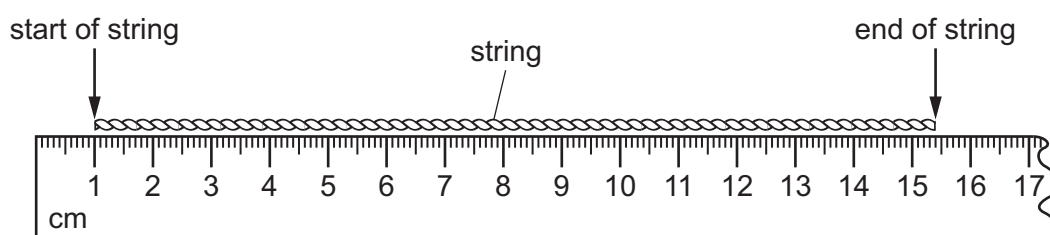
27 Which substance rapidly decolourises aqueous bromine?

- A ethane
- B ethanol
- C ethene
- D poly(ethene)

28 A student uses a piece of string to measure the circumference of a pencil.

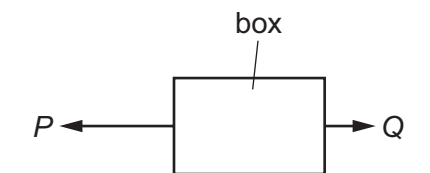
He wraps the string around the outside of the pencil. The string wraps round exactly six times.

He measures the length of string used with a rule.



What is the circumference of the pencil?

- A 2.4 cm
 - B 2.6 cm
 - C 14.4 cm
 - D 15.4 cm
- 29 The diagram shows a large force of magnitude P and a small force of magnitude Q acting on a box.



Which expression gives the magnitude of the resultant force on the box?

- A $P + Q$
- B $P - Q$
- C $P \times Q$
- D $\frac{P}{Q}$

30 A ball falls vertically downwards.

Which energy transfer takes place as the ball accelerates downwards?

- A gravitational potential to elastic potential (strain)
- B gravitational potential to kinetic
- C elastic potential (strain) to kinetic
- D kinetic to gravitational potential

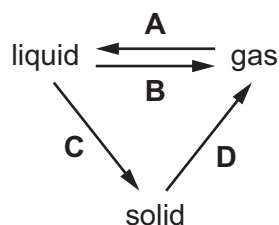
31 Four cars travel the same distance directly up the same steep hill.

The weights of the cars and the times for their journeys are shown in the table.

Which car develops the greatest power?

	weight of car / N	time taken / s
A	15 000	10
B	15 000	15
C	20 000	10
D	20 000	15

32 Which labelled arrow on the diagram represents condensation?

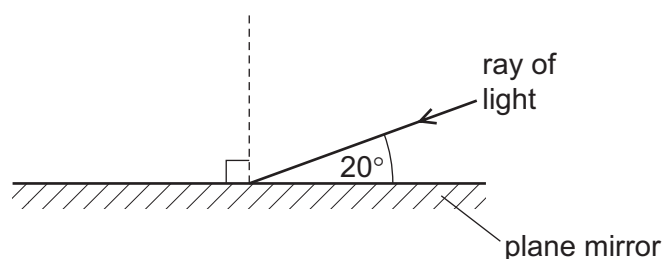


33 Water in a metal pan is heated on a gas burner.

What are the main methods by which heat is transferred through the metal pan to the water and throughout the water?

- A** conduction through the metal pan and convection in the water
- B** convection through the metal pan and conduction in the water
- C** convection through the metal pan and radiation in the water
- D** radiation through the metal pan and conduction in the water

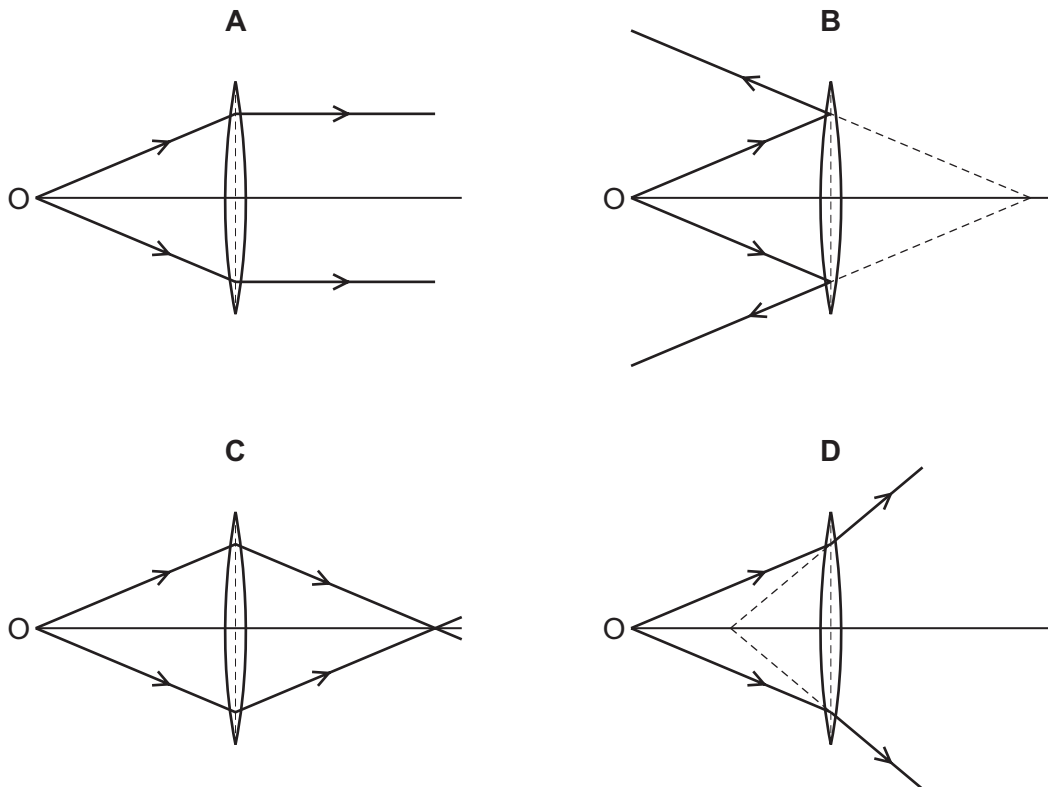
34 The diagram shows a ray of light striking a plane mirror.



What is the angle of reflection?

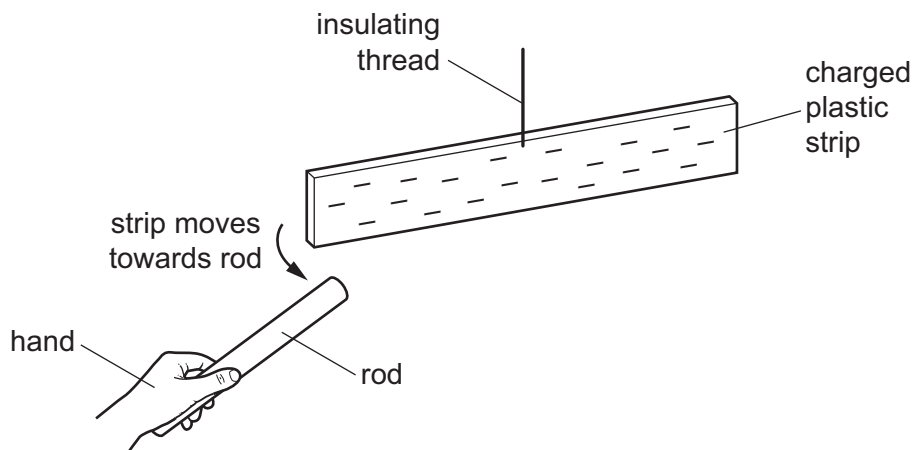
- A** 20°
- B** 40°
- C** 70°
- D** 90°

35 Which diagram shows a converging lens forming a real image of an object O?



36 A rod is rubbed with a dry piece of cloth. A scientist holds the rod in her hand and brings it close to a negatively charged plastic strip. The strip is suspended by an insulating thread.

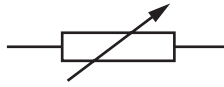
As the rod approaches the plastic strip, the strip moves towards the rod.



Which statement is correct?

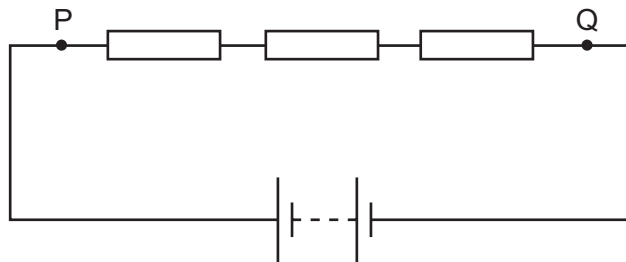
- A The rod is a negatively charged electrical conductor.
- B The rod is a negatively charged electrical insulator.
- C The rod is a positively charged electrical conductor.
- D The rod is a positively charged electrical insulator.

37 What is represented by the circuit symbol shown?



- A fixed resistor
- B fuse
- C switch
- D variable resistor

38 Three resistors are connected in series with a battery, as shown.



The current at point P is 6.0 A.

What is the current at point Q?

- A 0 A
- B 2.0 A
- C 3.0 A
- D 6.0 A

39 The table shows the usual current in each of four household appliances and the fuse used to protect each of them.

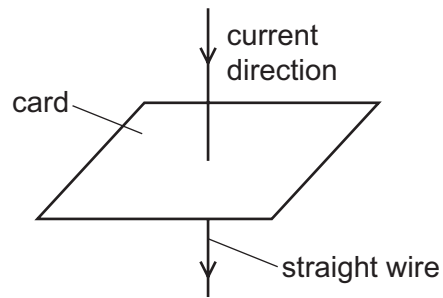
The only fuses available are rated at 3 A, 5 A or 13 A.

Which row shows an appliance that has been fitted with the most appropriate of the fuses available?

	appliance	current / A	fuse rating / A
A	hairdryer	5.5	5
B	kettle	7.5	13
C	lawnmower	5.0	3
D	slow cooker	1.0	5

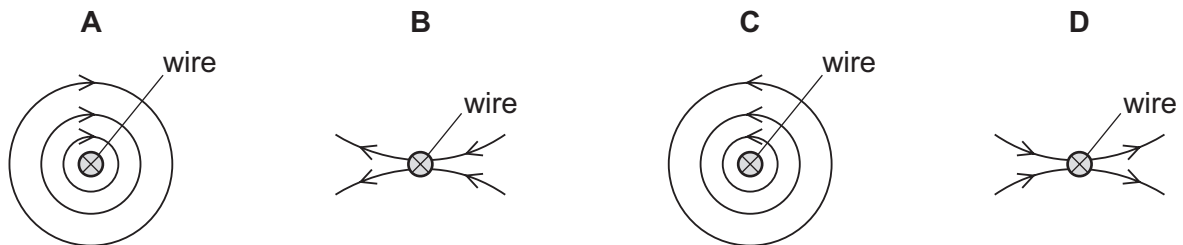
40 A straight vertical wire passes through the centre of a card.

The wire carries a current in the direction shown.



The current produces a magnetic field around the wire.

Which diagram shows the pattern of the magnetic field lines and their direction when seen from above?



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The Periodic Table of Elements

Group																		
I	II	III										IV	V	VI	VII	VIII		
3 Li lithium 7	4 Be beryllium 9	<div style="border: 1px solid black; padding: 5px; text-align: center;"> Key atomic number atomic symbol name relative atomic mass </div>										5 B boron 11	6 C carbon 12	7 N nitrogen 14	8 O oxygen 16	9 F fluorine 19	10 Ne neon 20	
11 Na sodium 23	12 Mg magnesium 24											13 Al aluminium 27	14 Si silicon 28	15 P phosphorus 31	16 S sulfur 32	17 Cl chlorine 35.5	18 Ar argon 40	
19 K potassium 39	20 Ca calcium 40	21 Sc scandium 45	22 Ti titanium 48	23 V vanadium 51	24 Cr chromium 52	25 Mn manganese 55	26 Fe iron 56	27 Co cobalt 59	28 Ni nickel 59	29 Cu copper 64	30 Zn zinc 65	31 Ga gallium 70	32 Ge germanium 73	33 As arsenic 75	34 Se selenium 79	35 Br bromine 80	36 Kr krypton 84	
37 Rb rubidium 85	38 Sr strontium 88	39 Y yttrium 89	40 Zr zirconium 91	41 Nb niobium 93	42 Mo molybdenum 96	43 Tc technetium —	44 Ru ruthenium 101	45 Rh rhodium 103	46 Pd palladium 106	47 Ag silver 108	48 Cd cadmium 112	49 In indium 115	50 Sn tin 119	51 Sb antimony 122	52 Te tellurium 128	53 I iodine 127	54 Xe xenon 131	
55 Cs caesium 133	56 Ba barium 137	57–71 lanthanoids	72 Hf hafnium 178	73 Ta tantalum 181	74 W tungsten 184	75 Re rhenium 186	76 Os osmium 190	77 Ir iridium 192	78 Pt platinum 195	79 Au gold 197	80 Hg mercury 201	81 Tl thallium 204	82 Pb lead 207	83 Bi bismuth 209	84 Po polonium —	85 At astatine —	86 Rn radon —	
87 Fr francium —	88 Ra radium —	89–103 actinoids	104 Rf rutherfordium —	105 Db dubnium —	106 Sg seaborgium —	107 Bh bohrium —	108 Hs hassium —	109 Mt meitnerium —	110 Ds darmstadtium —	111 Rg roentgenium —	112 Cn copernicium —	114 Fl flerovium —	116 Lv livermorium —	—	—	—	—	—

lanthanoids	57 La lanthanum 139	58 Ce cerium 140	59 Pr praseodymium 141	60 Nd neodymium 144	61 Pm promethium —	62 Sm samarium 150	63 Eu europium 152	64 Gd gadolinium 157	65 Tb terbium 159	66 Dy dysprosium 163	67 Ho holmium 165	68 Er erbium 167	69 Tm thulium 169	70 Yb ytterbium 173	71 Lu lutetium 175
actinoids	89 Ac actinium —	90 Th thorium 232	91 Pa protactinium 231	92 U uranium 238	93 Np neptunium —	94 Pu plutonium —	95 Am americium —	96 Cm curium —	97 Bk berkelium —	98 Cf californium —	99 Es einsteinium —	100 Fm fermium —	101 Md mendelevium —	102 No nobelium —	103 Lr lawrencium —

The volume of one mole of any gas is 24 dm³ at room temperature and pressure (r.t.p.).