



Cambridge IGCSE™ (9–1)

CO-ORDINATED SCIENCES

0973/31

Paper 3 Theory (Core)

October/November 2022

MARK SCHEME

Maximum Mark: 120

Published

This mark scheme is published as an aid to teachers and candidates, to indicate the requirements of the examination. It shows the basis on which Examiners were instructed to award marks. It does not indicate the details of the discussions that took place at an Examiners' meeting before marking began, which would have considered the acceptability of alternative answers.

Mark schemes should be read in conjunction with the question paper and the Principal Examiner Report for Teachers.

Cambridge International will not enter into discussions about these mark schemes.

Cambridge International is publishing the mark schemes for the October/November 2022 series for most Cambridge IGCSE™, Cambridge International A and AS Level components and some Cambridge O Level components.

This document consists of **13** printed pages.

PUBLISHED**Generic Marking Principles**

These general marking principles must be applied by all examiners when marking candidate answers. They should be applied alongside the specific content of the mark scheme or generic level descriptors for a question. Each question paper and mark scheme will also comply with these marking principles.

GENERIC MARKING PRINCIPLE 1:

Marks must be awarded in line with:

- the specific content of the mark scheme or the generic level descriptors for the question
- the specific skills defined in the mark scheme or in the generic level descriptors for the question
- the standard of response required by a candidate as exemplified by the standardisation scripts.

GENERIC MARKING PRINCIPLE 2:

Marks awarded are always **whole marks** (not half marks, or other fractions).

GENERIC MARKING PRINCIPLE 3:

Marks must be awarded **positively**:

- marks are awarded for correct/valid answers, as defined in the mark scheme. However, credit is given for valid answers which go beyond the scope of the syllabus and mark scheme, referring to your Team Leader as appropriate
- marks are awarded when candidates clearly demonstrate what they know and can do
- marks are not deducted for errors
- marks are not deducted for omissions
- answers should only be judged on the quality of spelling, punctuation and grammar when these features are specifically assessed by the question as indicated by the mark scheme. The meaning, however, should be unambiguous.

GENERIC MARKING PRINCIPLE 4:

Rules must be applied consistently, e.g. in situations where candidates have not followed instructions or in the application of generic level descriptors.

GENERIC MARKING PRINCIPLE 5:

Marks should be awarded using the full range of marks defined in the mark scheme for the question (however; the use of the full mark range may be limited according to the quality of the candidate responses seen).

GENERIC MARKING PRINCIPLE 6:

Marks awarded are based solely on the requirements as defined in the mark scheme. Marks should not be awarded with grade thresholds or grade descriptors in mind.

Science-Specific Marking Principles

1	Examiners should consider the context and scientific use of any keywords when awarding marks. Although keywords may be present, marks should not be awarded if the keywords are used incorrectly.
2	The examiner should not choose between contradictory statements given in the same question part, and credit should not be awarded for any correct statement that is contradicted within the same question part. Wrong science that is irrelevant to the question should be ignored.
3	Although spellings do not have to be correct, spellings of syllabus terms must allow for clear and unambiguous separation from other syllabus terms with which they may be confused (e.g. ethane / ethene, glucagon / glycogen, refraction / reflection).
4	The error carried forward (ecf) principle should be applied, where appropriate. If an incorrect answer is subsequently used in a scientifically correct way, the candidate should be awarded these subsequent marking points. Further guidance will be included in the mark scheme where necessary and any exceptions to this general principle will be noted.
5	<p><u>'List rule' guidance</u></p> <p>For questions that require <i>n</i> responses (e.g. State two reasons ...):</p> <ul style="list-style-type: none"> • The response should be read as continuous prose, even when numbered answer spaces are provided. • Any response marked <i>ignore</i> in the mark scheme should not count towards <i>n</i>. • Incorrect responses should not be awarded credit but will still count towards <i>n</i>. • Read the entire response to check for any responses that contradict those that would otherwise be credited. Credit should not be awarded for any responses that are contradicted within the rest of the response. Where two responses contradict one another, this should be treated as a single incorrect response. • Non-contradictory responses after the first <i>n</i> responses may be ignored even if they include incorrect science.

6 Calculation specific guidance

Correct answers to calculations should be given full credit even if there is no working or incorrect working, **unless** the question states 'show your working'.

For questions in which the number of significant figures required is not stated, credit should be awarded for correct answers when rounded by the examiner to the number of significant figures given in the mark scheme. This may not apply to measured values.

For answers given in standard form (e.g. $a \times 10^n$) in which the convention of restricting the value of the coefficient (a) to a value between 1 and 10 is not followed, credit may still be awarded if the answer can be converted to the answer given in the mark scheme.

Unless a separate mark is given for a unit, a missing or incorrect unit will normally mean that the final calculation mark is not awarded. Exceptions to this general principle will be noted in the mark scheme.

7 Guidance for chemical equations

Multiples / fractions of coefficients used in chemical equations are acceptable unless stated otherwise in the mark scheme.

State symbols given in an equation should be ignored unless asked for in the question or stated otherwise in the mark scheme.

Question	Answer	Marks
1(a)(i)	rice ; chicken ; (green) beans / cucumber / onions / salad ;	3
1(a)(ii)	as a solvent / <i>any valid point</i> ;	1
1(b)(i)	blue-black ;	1
1(b)(ii)	glucose circled ;	1
1(b)(iii)	prevents constipation / keeps food moving through the gut ;	1
1(c)	carbon, hydrogen and oxygen ticked ;	1

Question	Answer	Marks
2(a)(i)	chlorination ;	1
2(a)(ii)	chromatography ;	1
2(a)(iii)	electrolysis ;	1
2(a)(iv)	polymerisation ;	1
2(b)(i)	burette ;	1
2(b)(ii)	7 ;	1
2(b)(iii)	sodium chloride ; water ;	2
2(c)	pale blue precipitate ;	1

Question	Answer	Marks
3(a)(i)	<i>any two from:</i> coal gas petroleum ;;	2
3(a)(ii)	only existing as a finite quantity / being used up at a faster rate than it can be replaced ;	1
3(a)(iii)	nuclear accidents / nuclear / dangerous / harmful waste produced ;	1
3(b)(i)	(nucleus) splits ;	1
3(b)(ii)	in a lead lined container ;	1
3(c)	4 half-lives / division by 16 ; 6.25% ;	2

Question	Answer	Marks
4(a)	gland ; organs ;	2
4(b)	plasma ;	1
4(c)(i)	4 (minutes) ;	1
4(c)(ii)	109 (bpm) ;	1
4(c)(iii)	adrenaline ;	1
4(c)(iv)	widened pupils ; increased breathing rate ; <i>any valid point ;;</i>	2

Question	Answer	Marks
4(d)	peripheral ; brain ; spinal cord ;	3
5(a)(i)	1 shared pair ; all else correct ;	2
5(a)(ii)	water vapour ; carbon dioxide ;	1
5(b)	(thermal) energy released ;	1
5(c)	B – no mark and has formula C ₂ H ₆ ;	1
5(d)(i)	A – no mark and has two carbons but only four hydrogens / has general formula C _n H _{2n} ;	1
5(d)(ii)	(aqueous) bromine ;	1
5(d)(iii)	decolourises ;	1

Question	Answer	Marks
6(a)(i)	evaporation	1
6(a)(ii)	fastest moving / most energetic molecules / particles ; escape from surface ;	2
6(b)	Weight ;	1
6(c)(i)	4.4 (m / s) ;	1
6(c)(ii)	distance = speed × time (in any form symbols or words) or 4.4 × 200 or area under graph ; 880 (m) ;	2

Question	Answer	Marks
6(d)(i)	8 (m) ;	1
6(d)(ii)	amplitude correctly indicated ;	1
6(d)(iii)	2.0 ;	1

Question	Answer	Marks
7(a)(i)	light ; chlorophyll ;	2
7(a)(ii)	carbon dioxide is required for photosynthesis ;	1
7(a)(iii)	carbon dioxide + water → glucose and oxygen ;;	2
7(b)	chloroplast ;	1

Question	Answer	Marks
7(c)	<div style="display: flex; align-items: center; justify-content: center;"> <div style="border: 1px solid black; padding: 5px; margin-right: 20px;">Magnesium ions</div> <div style="display: flex; flex-direction: column; gap: 10px;"> <div style="border: 1px solid black; padding: 5px; width: 200px;">are a component of starch in a plant</div> <div style="border: 1px solid black; padding: 5px; width: 200px;">are absorbed by the root hairs.</div> <div style="border: 1px solid black; padding: 5px; width: 200px;">are transported in the xylem.</div> <div style="border: 1px solid black; padding: 5px; width: 200px;">are required for synthesis of chlorophyll.</div> <div style="border: 1px solid black; padding: 5px; width: 200px;">enter a plant through the stomata.</div> <div style="border: 1px solid black; padding: 5px; width: 200px;">provide the energy for photosynthesis.</div> </div> </div> <div style="text-align: right; margin-top: 20px;">⋮</div>	3
7(d)	nitrate ;	1

Question	Answer	Marks
8(a)(i)	29 ;	1
8(a)(ii)	34 ;	1
8(b)(i)	copper oxide + carbon → copper + carbon dioxide ;	1
8(b)(ii)	loss of oxygen ;	1

Question	Answer	Marks
8(b)(iii)	metal oxide / reacts with an acid to make a salt / neutralises an acid ;	1
8(c)	<i>any one from:</i> forms coloured compounds ; acts as catalyst ; variable valency ;	1
8(d)(i)	alloy ;	1
8(d)(ii)	can be beaten into shape;	1
8(d)(iii)	Cu and Zn ;	1
8(e)(i)	cathode / negative ;	1
8(e)(ii)	copper (II) sulfate ;	1

Question	Answer	Marks
9(a)(i)	all symbols correct ; two lamps in parallel with battery ; switch to control both lamps ;	3
9(a)(ii)	$R = V / I$ (in any form symbols or words) or $12 / 4$; $= 3$; ohms / Ω ;	3
9(b)(i)	gap is smaller / closes ;	1
9(b)(ii)	road/bridge needs to expand on hot day / road/bridge could be damaged by expansion ;	1
9(c)(i)	amplitude increases ;	1
9(c)(ii)	frequency increases ;	1

Question	Answer	Marks
9(c)(iii)	kinetic (energy); gravitational potential (energy) ;	2

Question	Answer	Marks																									
10(a)(i)	cactus → desert rats → snake → desert fox ;;	2																									
10(a)(ii)	<table border="1"> <thead> <tr> <th>organism</th> <th>carnivore</th> <th>herbivore</th> <th>producer</th> <th>secondary consumer</th> </tr> </thead> <tbody> <tr> <td>cactus</td> <td></td> <td></td> <td>✓</td> <td></td> </tr> <tr> <td>desert fox</td> <td>✓</td> <td></td> <td></td> <td></td> </tr> <tr> <td>desert rat</td> <td></td> <td>✓</td> <td></td> <td></td> </tr> <tr> <td>snake</td> <td>✓</td> <td></td> <td></td> <td>✓</td> </tr> </tbody> </table> <p>1 mark for each correct row ;;;</p>	organism	carnivore	herbivore	producer	secondary consumer	cactus			✓		desert fox	✓				desert rat		✓			snake	✓			✓	3
organism	carnivore	herbivore	producer	secondary consumer																							
cactus			✓																								
desert fox	✓																										
desert rat		✓																									
snake	✓			✓																							
10(b)	decomposer ;	1																									
10(c)	Sun ;	1																									
10(d)	<i>any three from:</i> movement ; respiration ; sensitivity ; growth ; reproduction ; excretion ;	3																									
10(e)	cells ;	1																									

Question	Answer	Marks
11(a)	element – contains only one type of atom ; compound – contains two or more elements (chemically combined) ;	2
11(b)	$2\text{Na} + \text{Cl}_2 \rightarrow 2\text{NaCl}$;	1
11(c)	sodium ion 2.8 ; chloride ion 2.8.8 ;	2
11(d)	solubility of ionic compound is greater ;	1
11(e)	add acidified aqueous silver nitrate ; white precipitate ;	2
11(f)	molten ; electricity ; hydrogen ; chlorine ;	4

Question	Answer	Marks							
12(a)	<table border="1" style="width: 100%; border-collapse: collapse;"> <tr> <td style="width: 15%; text-align: center;">γ-rays</td> <td style="width: 15%; text-align: center;">X-rays</td> <td style="width: 15%; text-align: center;">ultraviolet</td> <td style="width: 15%; text-align: center;">visible light</td> <td style="width: 15%; text-align: center;">infrared</td> <td style="width: 15%; text-align: center;">microwaves</td> <td style="width: 15%; text-align: center;">radio waves</td> </tr> </table> microwaves in correct place ; X-rays in correct place ;	γ -rays	X-rays	ultraviolet	visible light	infrared	microwaves	radio waves	2
γ -rays	X-rays	ultraviolet	visible light	infrared	microwaves	radio waves			
12(b)	area decreases ; so pressure increases ;	2							
12(c)(i)	volume = mass / density (in any form symbols or words) or $8000 / 1600$; = 5 m^3 ;	2							
12(c)(ii)	8000×10 ; (= $80\,000 \text{ N}$)	1							
12(d)(i)	rays meet at focus (F) ;	1							

Question	Answer	Marks
12(d)(ii)	principal focus ;	1
12(d)(iii)	focal length correctly identified ;	1